

Supplementary Material for

“Binding and Degradation Reaction of Hydroxide Ions with Several Quaternary Ammonium Head Groups of Anion Exchange Membranes Investigated by the DFT Method”

Mirat Karibayev ¹, Bauyrzhan Myrzakhmetov ², Sandugash Kalybekkyzy ², Yanwei Wang ^{1,3,*} and Almagul Mentbayeva ^{1,2,*}

¹Department of Chemical and Materials Engineering, School of Engineering and Digital Sciences, Nazarbayev University, Nur-Sultan 010000, Kazakhstan

²Laboratory of Advanced Materials and Systems for Energy Storage, Center for Energy and Advanced

Materials Science, National Laboratory Astana, Nazarbayev University, Nur-Sultan 010000, Kazakhstan

³Laboratory of Computational Materials Science for Energy Applications, Center for Energy and Advanced Materials Science, National Laboratory Astana, Nazarbayev University, Nur-Sultan 010000, Kazakhstan

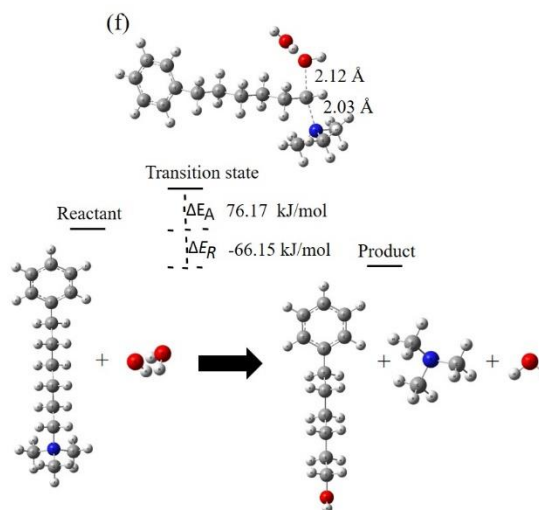
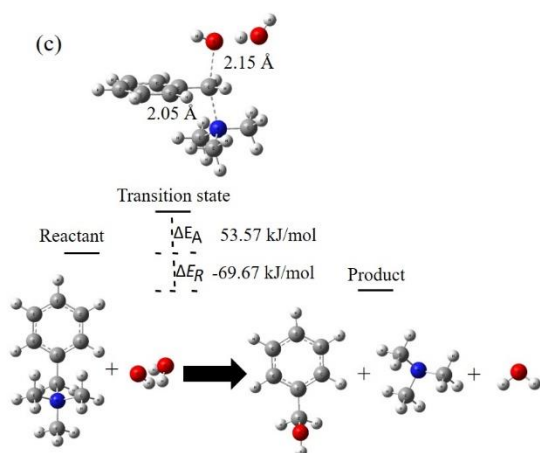
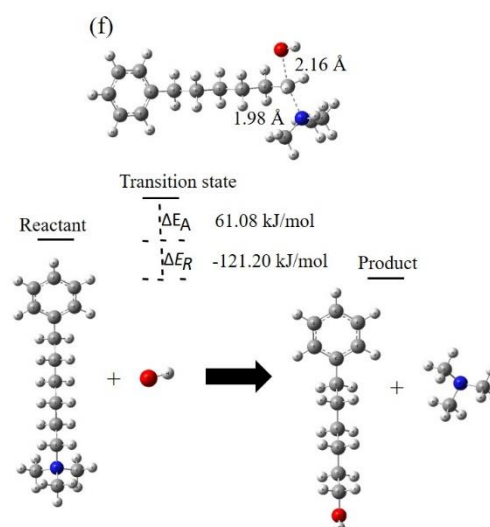
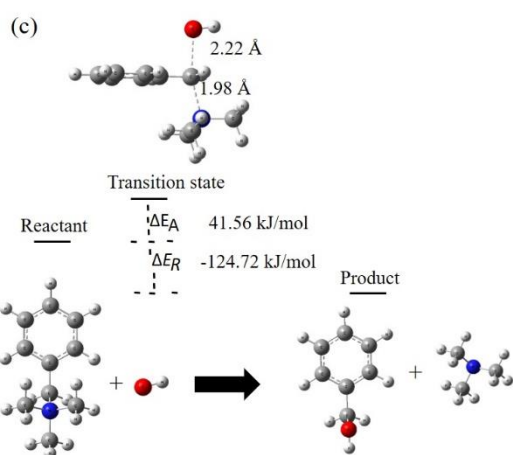
*Correspondence: yanwei.wang@nu.edu.kz (Y.W.); almagul.mentbayeva@nu.edu.kz (A.M.)

Table S1: The absolute energies of the calculated species and binding energies $\Delta E_{\text{Binding}}$ of QA head groups with OH^- ion, $\Delta E_{\text{binding}} = E_{\text{QA with OH}^-} - (E_{\text{QA}} + E_{\text{OH}^-})$.

QA	E (OH^-) (kJ/mol)	E (QA) (kJ/mol)	E (QA with OH^-) (kJ/mol)	$\Delta E_{\text{binding}}$ (kJ/mol)
(a)	−199431.16	−1466500.19	−1665964.27	−32.91
(c)	−199431.16	−1169420.80	−1368875.34	−23.37
(f)	−199431.16	−1685662.20	−1885115.30	−21.91
(d)	−199431.16	−1475970.88	−1675419.33	−17.28
(e)	−199431.16	−1766722.23	−1966162.29	−8.89
(b)	−199431.16	−1618116.55	−1817554.88	−7.16

Table S2: Values of the computed ΔE_R , and ΔE_A for S_N2 degradation reaction mechanism of QA head group (c)/(f). Unit: kJ/mol; Water molecules are treated as water cluster on the product side.

QA	HL	E(R1)	E(R2)	E(TS ⁻)	E(P1)	E(P2)	E(P3)	ΔE_R	BSSE	ΔE_A
(c)	0	-1169446.34	-199429.14	-1368822.03	-910761.97	-458238.23	—	-124.72	11.89	41.56
	1	-1169446.34	-400248.63	-1569632.52	-910761.97	-458238.23	-200764.45	-69.67	8.88	53.57
	2	-1169446.34	-601059.62	-1770433.28	-910761.97	-458238.23	-401545.78	-40.02	7.76	64.91
	3	-1169446.34	-801864.08	-1971217.59	-910761.97	-458238.23	-602329.96	-19.74	6.14	86.69
(f)	0	-1685699.31	-199429.14	-1885054.44	-1427011.42	-458238.23	—	-121.20	12.93	61.08
	1	-1685699.31	-400248.63	-2085861.60	-1427011.42	-458238.23	-200764.45	-66.15	10.16	76.17
	2	-1685699.31	-601059.62	-2286662.47	-1427011.42	-458238.23	-401545.78	-36.49	7.68	88.78
	3	-1685699.31	-801864.08	-2487450.75	-1427011.42	-458238.23	-602329.96	-16.21	6.39	106.25



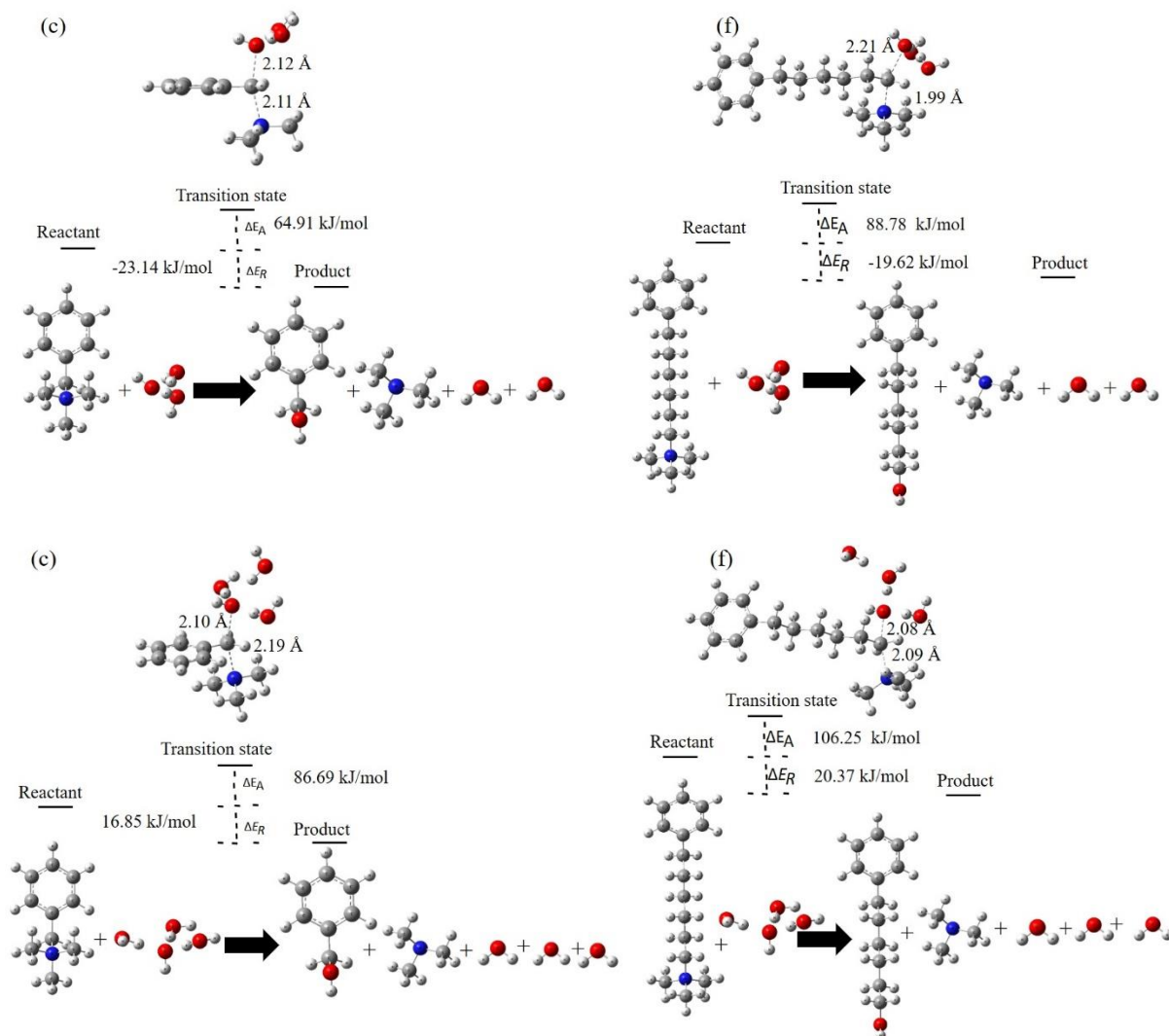


Figure S1: S_N2 degradation reaction for QA head group (c), and QA head group (f) at the different HL; Water molecules are treated as isolated monomers on the product side.

Table S3: Values of the computed ΔE_R , and ΔE_A for S_N2 degradation reaction mechanism of QA head group (c)/(f). Unit: kJ/mol; Water are treated as isolated monomers on the product side.

QA	HL	E(R1)	E(R2)	E(TS ⁻)	E(P1)	E(P2)	E(P3)	ΔE_R	BSSE	ΔE_A
(c)	0	-1169446.34	-199429.14	-1368822.03	-910761.97	-458238.23	—	-124.72	11.89	41.56
	1	-1169446.34	-400248.63	-1569632.52	-910761.97	-458238.23	-200764.45	-69.67	8.88	53.57
	2	-1169446.34	-601059.62	-1770433.28	-910761.97	-458238.23	-401528.91	-23.14	7.76	64.91
	3	-1169446.34	-801864.08	-1971217.59	-910761.97	-458238.23	-602293.37	16.85	6.14	86.69
(f)	0	-1685699.31	-199429.14	-1885054.44	-1427011.42	-458238.23	—	-121.20	12.93	61.08
	1	-1685699.31	-400248.63	-2085861.60	-1427011.42	-458238.23	-200764.45	-66.15	10.16	76.17
	2	-1685699.31	-601059.62	-2286662.47	-1427011.42	-458238.23	-401528.91	-19.62	7.68	88.78
	3	-1685699.31	-801864.08	-2487450.75	-1427011.42	-458238.23	-602293.37	20.37	6.39	106.25