

Supporting Materials

Novel sulfonamide-based carbamates as selective inhibitors of BChE

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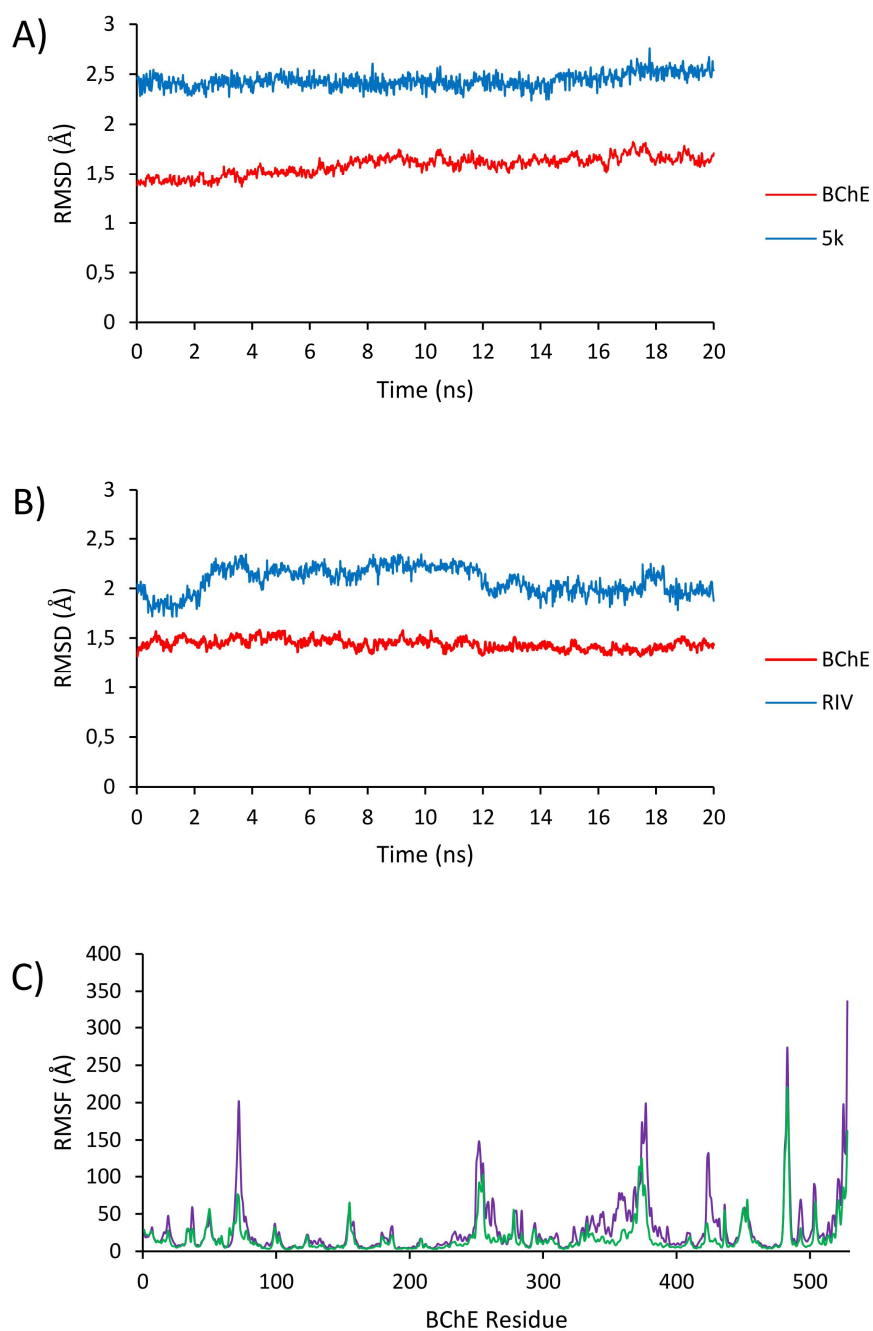


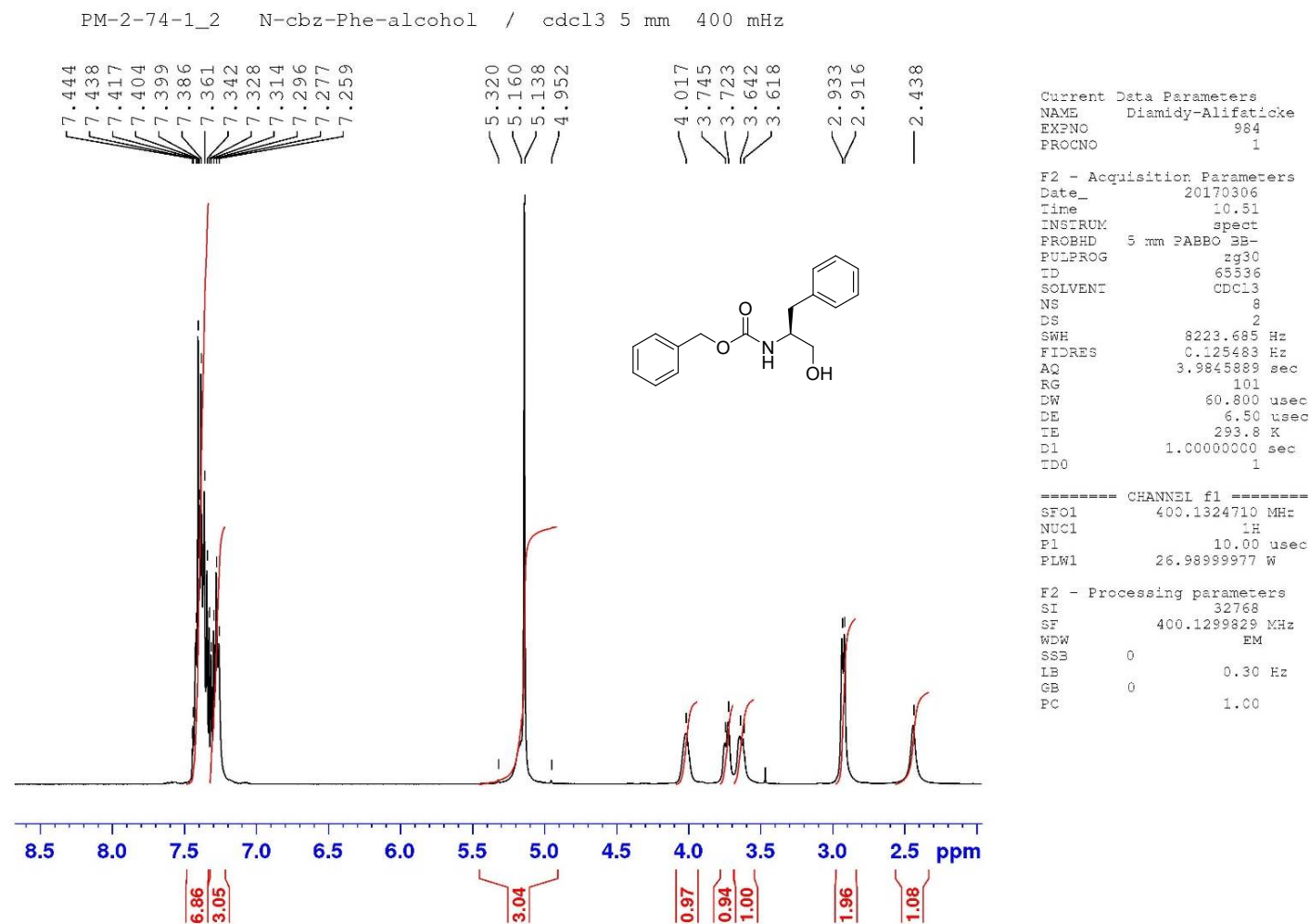
Figure S1 Root-mean-square deviations (RMSD) of (A) BChE-5k and (B) BChE-RIV complexes. (C) Comparative root-mean-square fluctuations (RMSF) of the C α atoms of BChE when complexed with 5k (violet) and RIV (green).

Experimental part

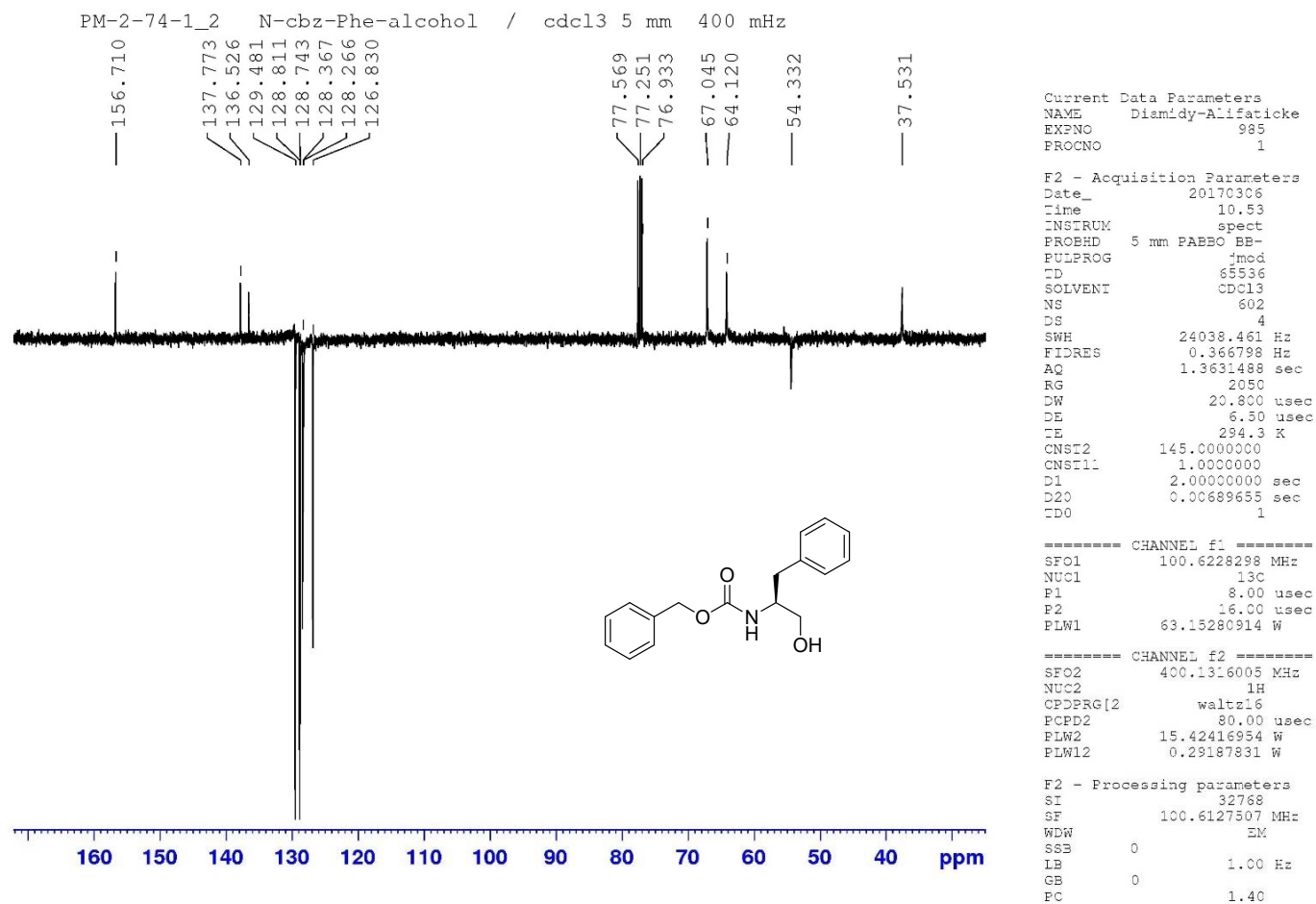
All reagents and solvents were purchased from commercial sources (TCI Europe, Sigma-Aldrich, Acros organics, fluorochem, Merck, Lach-Ner). Commercial grade reagents were used without further purification. Reactions were monitored by thin layer chromatography plates coated with 0.2 mm silica gel 60 F₂₅₄ (Merck). TLC plates were visualized by UV irradiation (254 nm) or in a 5% solution of phosphomolybdic acid in ethanol. All melting points were determined on a Melting Point B-540 apparatus (Büchi, Switzerland) and are uncorrected. IR spectra were recorded on a Nicolet 6700 FT-IR spectrometer (Thermo Fisher Scientific, Waltham, MA, USA) over the range of 400–4000 cm⁻¹ using the ATR technique. NMR spectra were measured in CDCl₃ solutions at ambient temperature on a Bruker Avance™ III 400 spectrometer at frequencies ¹H (400 MHz), ¹³C (100.26 MHz), and ¹⁹F (376.46) or a Bruker Ascend™ 500 spectrometer at frequencies ¹H (500.13 MHz), ¹³C {¹H} (125.76 MHz), and ¹⁹F (470.61 MHz). Chemical shifts, δ , are given in ppm, related to the residual solvent peak CDCl₃ 7.27 or to tetramethylsilane (TMS) as an internal standard. The coupling constants (*J*) are reported in [Hz]. Elemental analysis (C, H, N, S) was performed on an automatic microanalyzer Flash 2000 organic elemental analyzer. Mass spectrometry with high resolution was performed by the dried droplet method using a MALDI mass spectrometer LTQ Orbitrap XL (Thermo Fisher Scientific) equipped with a nitrogen UV laser (337 nm, 60 Hz). Spectra were measured in the positive ion mode and in regular mass extent with resolution 100,000 at *m/z* = 400. 2,5-Dihydroxybenzoic acid (DBH) was used as the matrix.

Characterization data of prepared compounds:

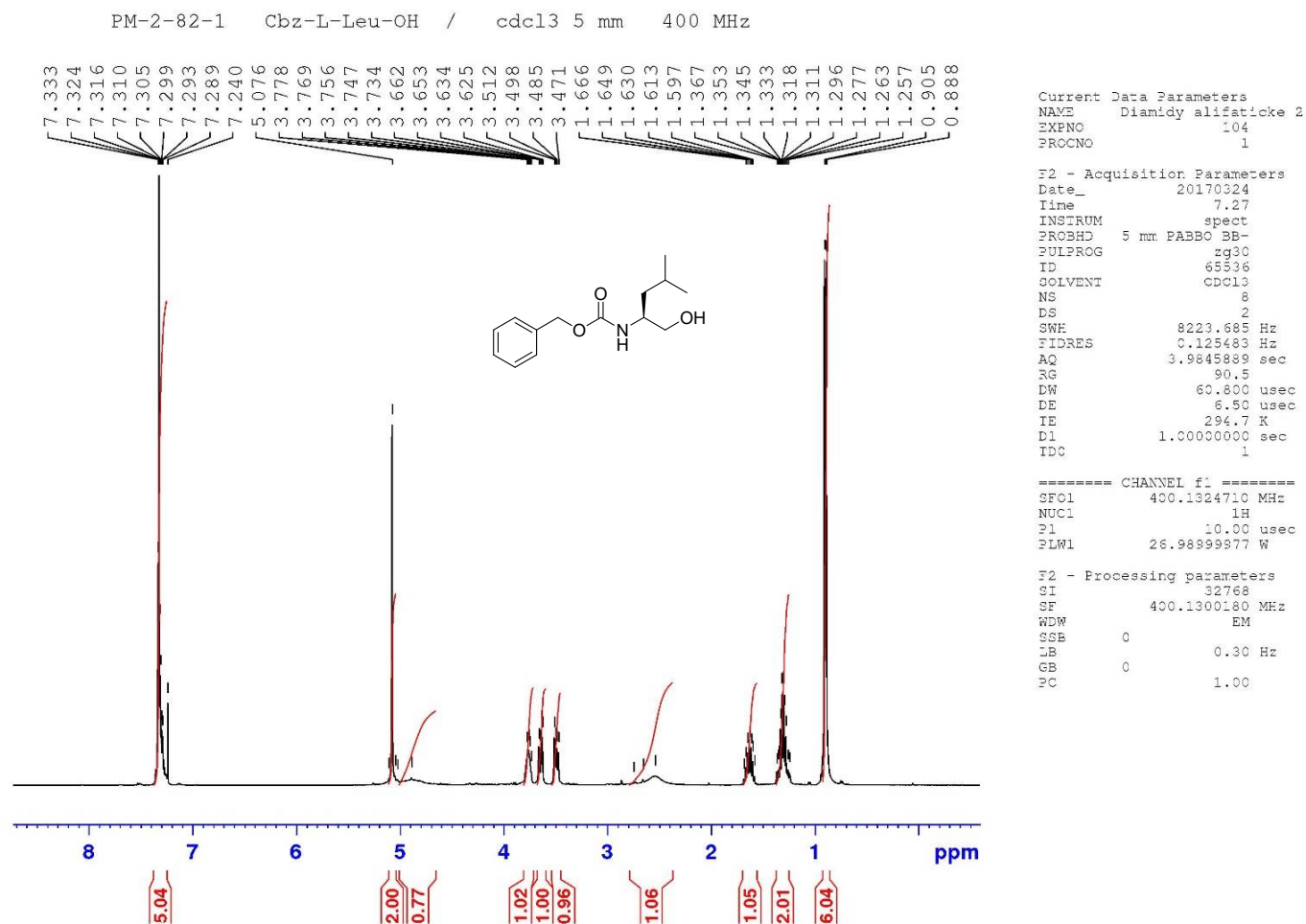
1.1. Copy of ^1H NMR of **2a**



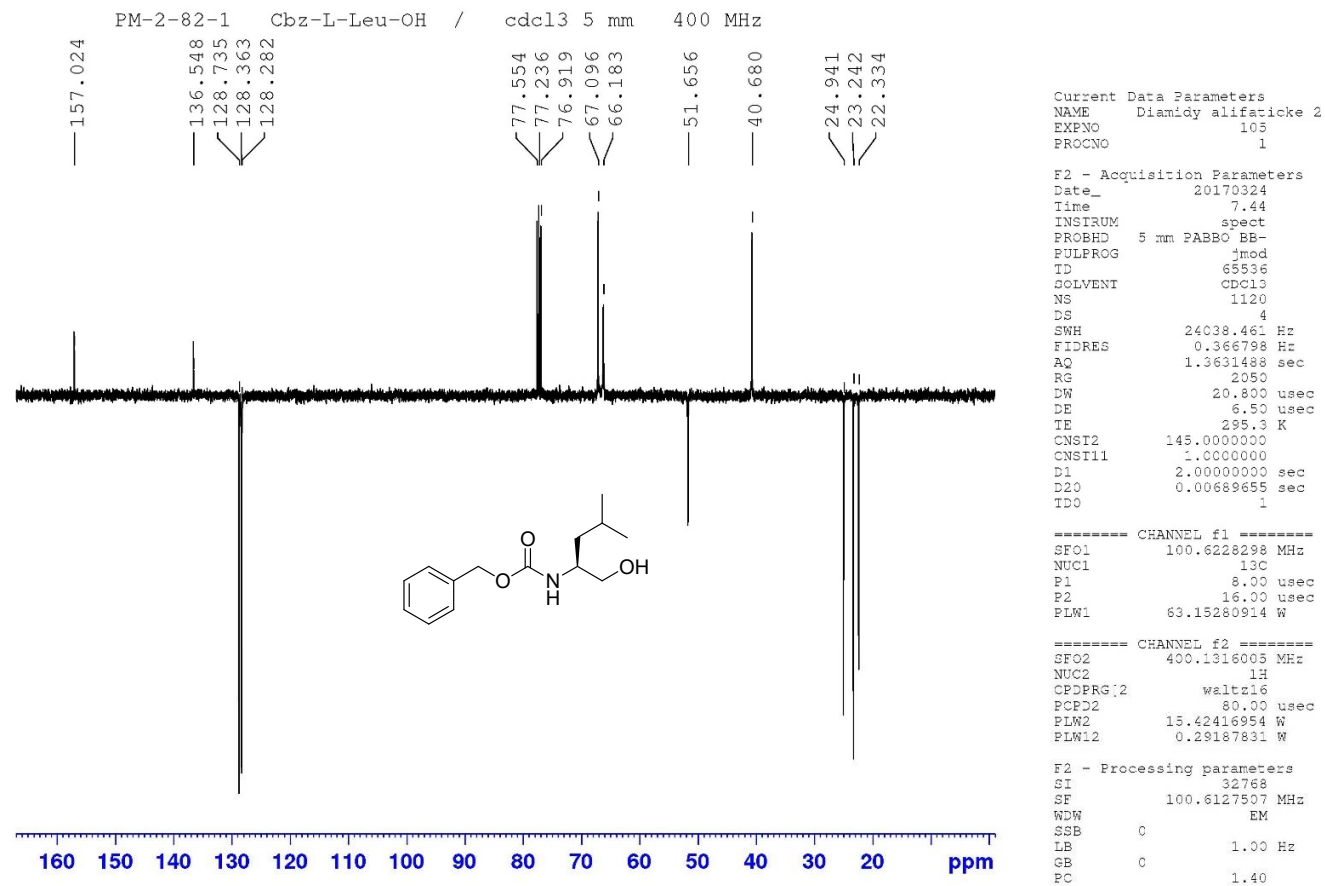
1.2. Copy of ^{13}C NMR of **2a**



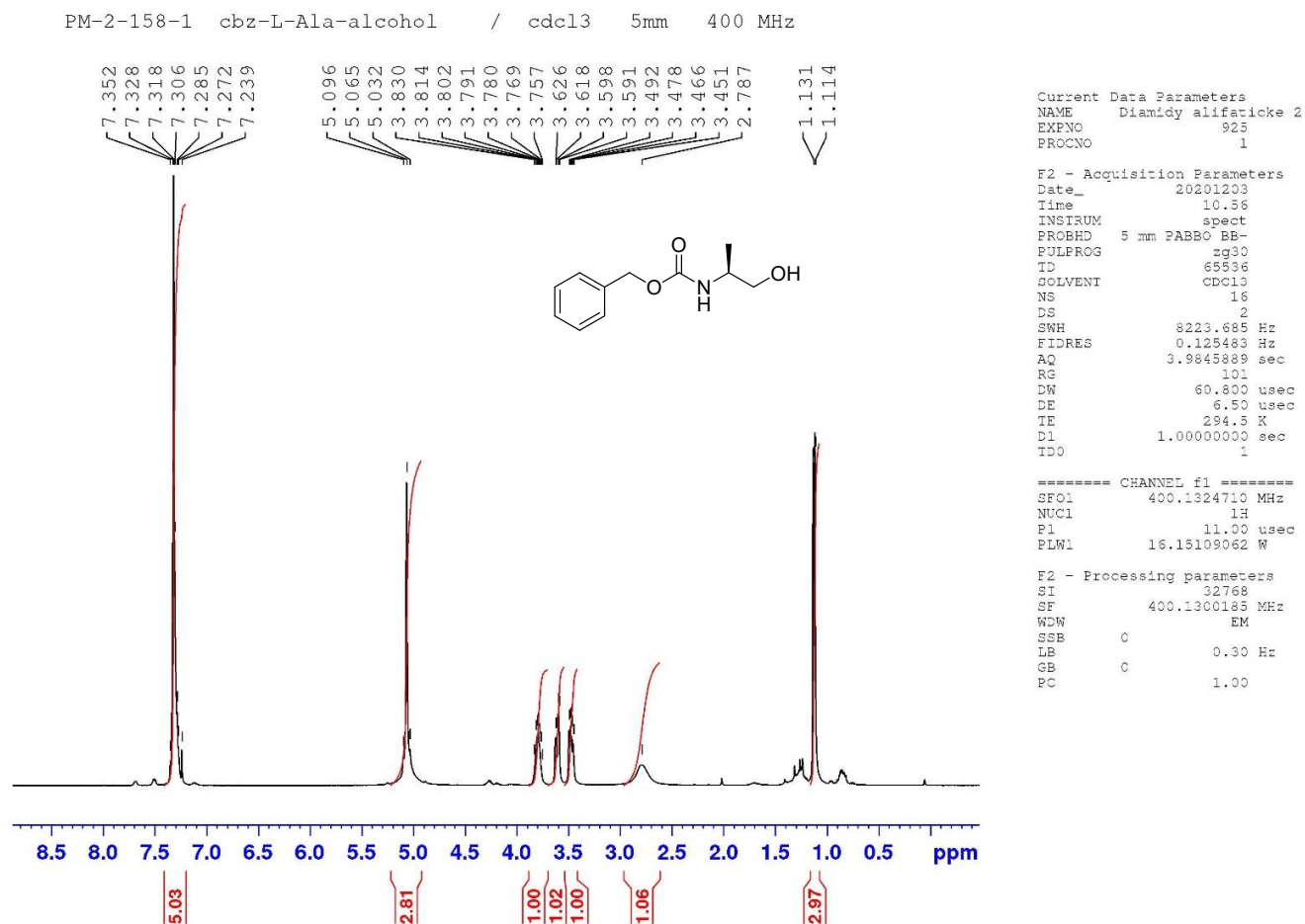
1.3. Copy of ^1H NMR of **2b**



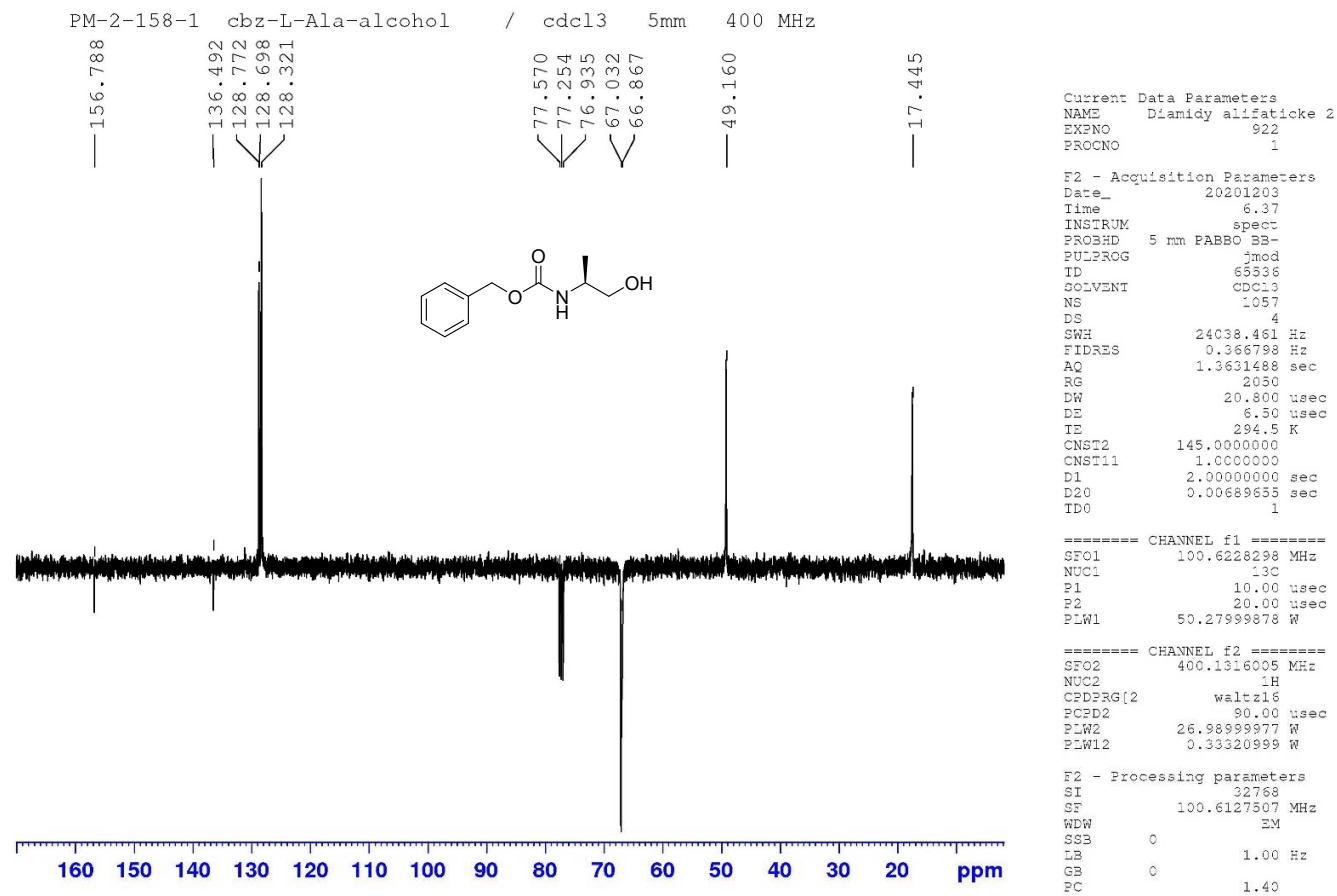
1.4. Copy of ^{13}C NMR of **2b**



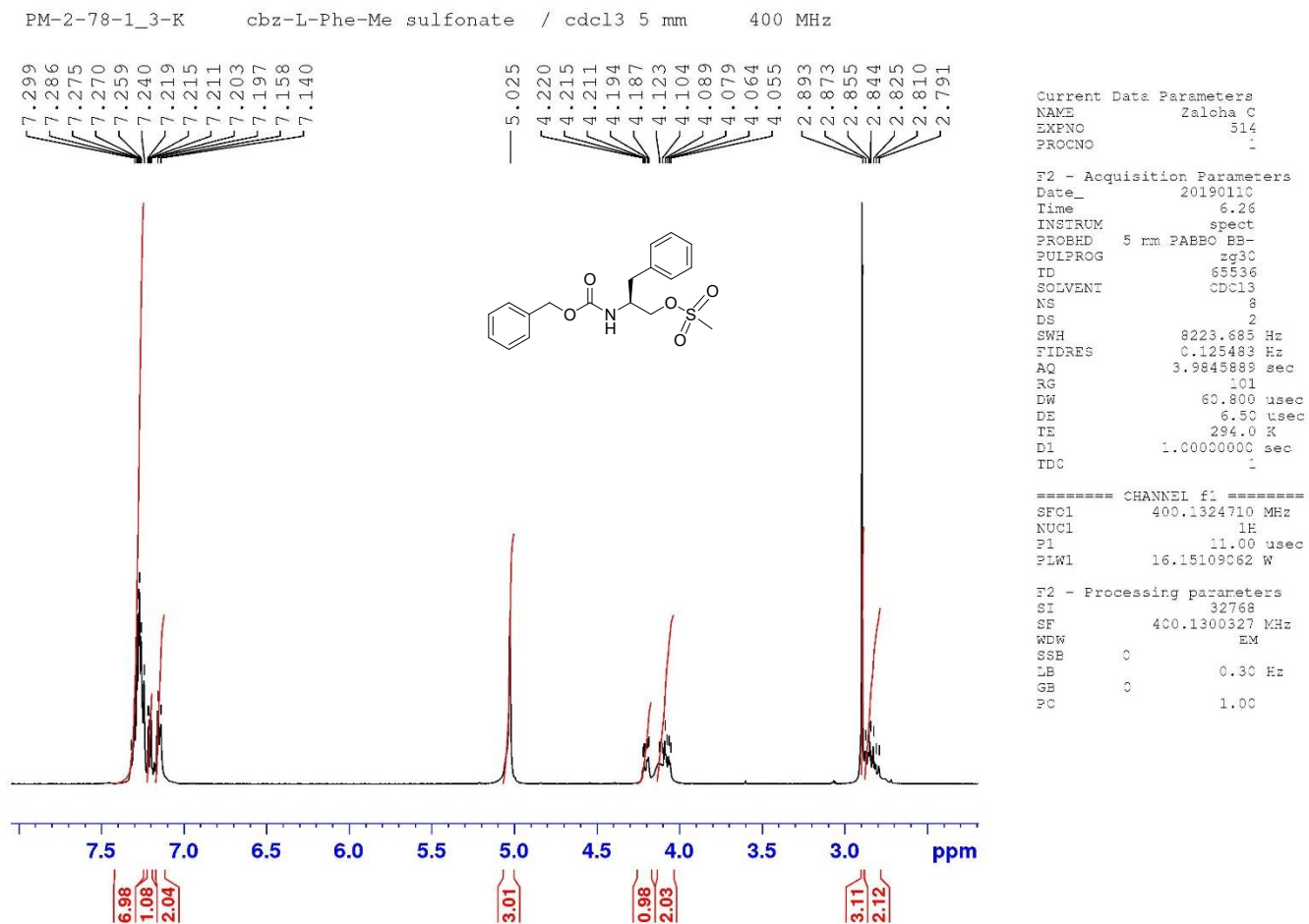
1.5. Copy of ^1H NMR of **2c**



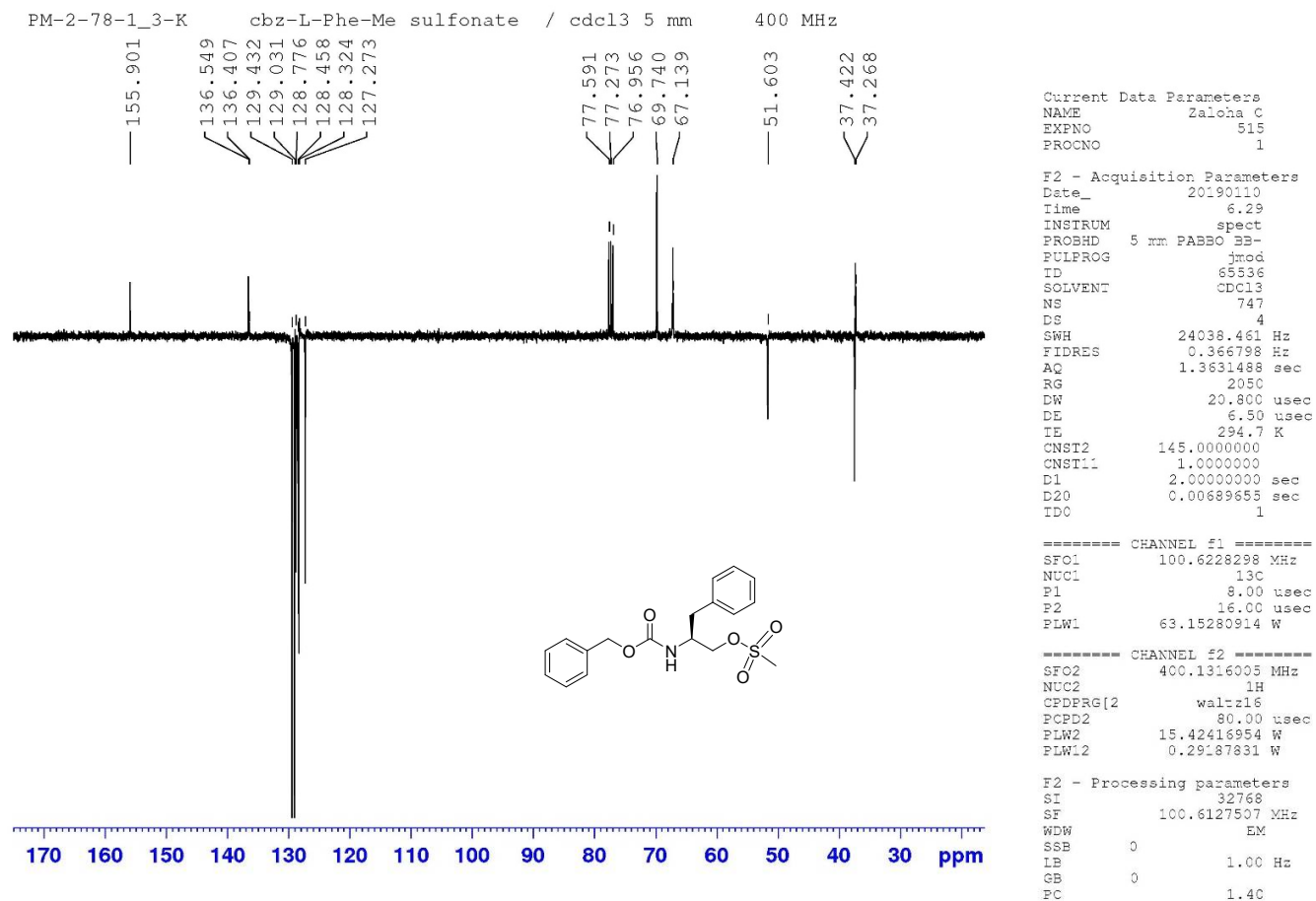
1.6. Copy of ^{13}C NMR of **2c**



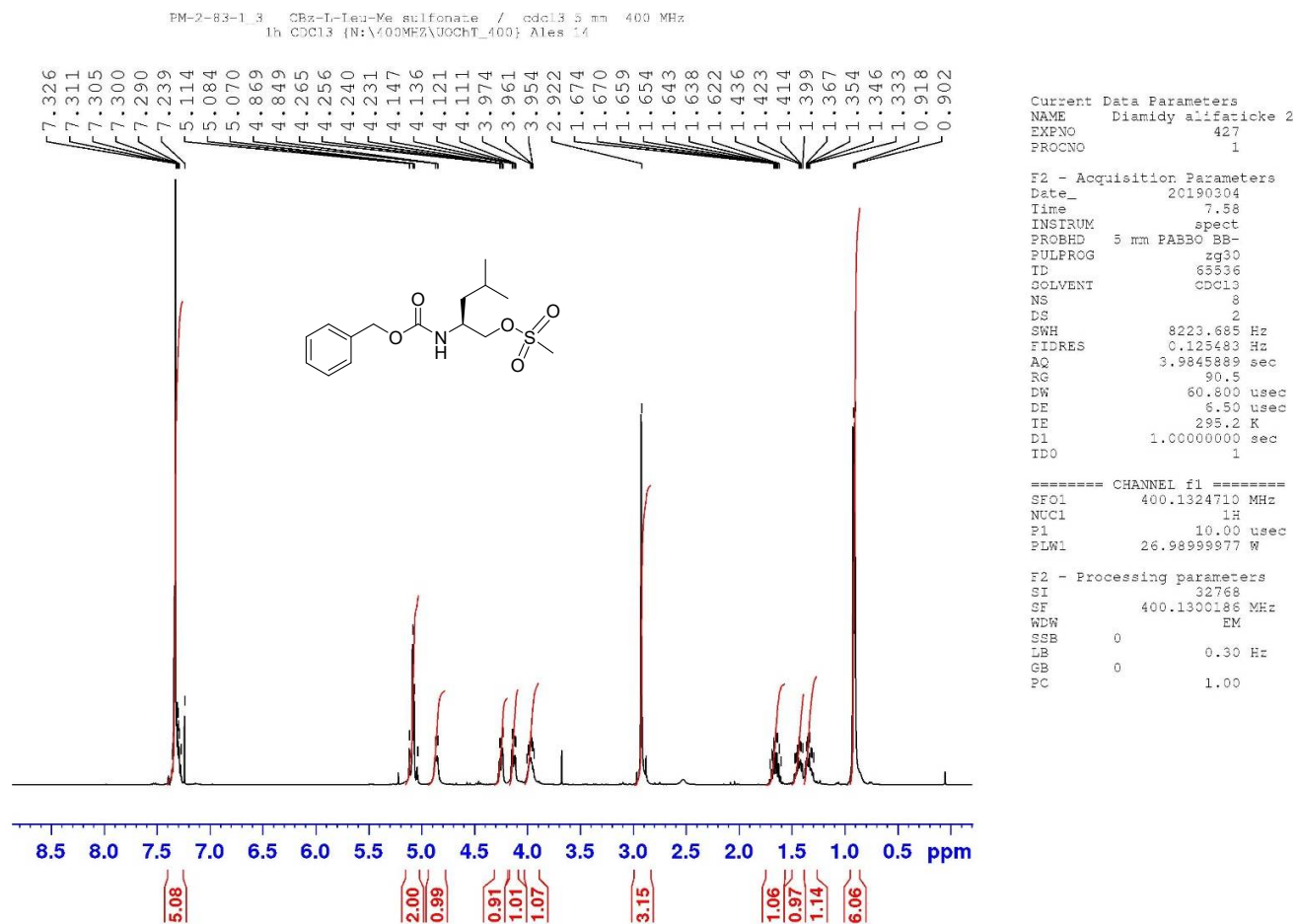
1.7. Copy of ^1H NMR of **3a**



1.8. Copy of ^{13}C NMR of **3a**



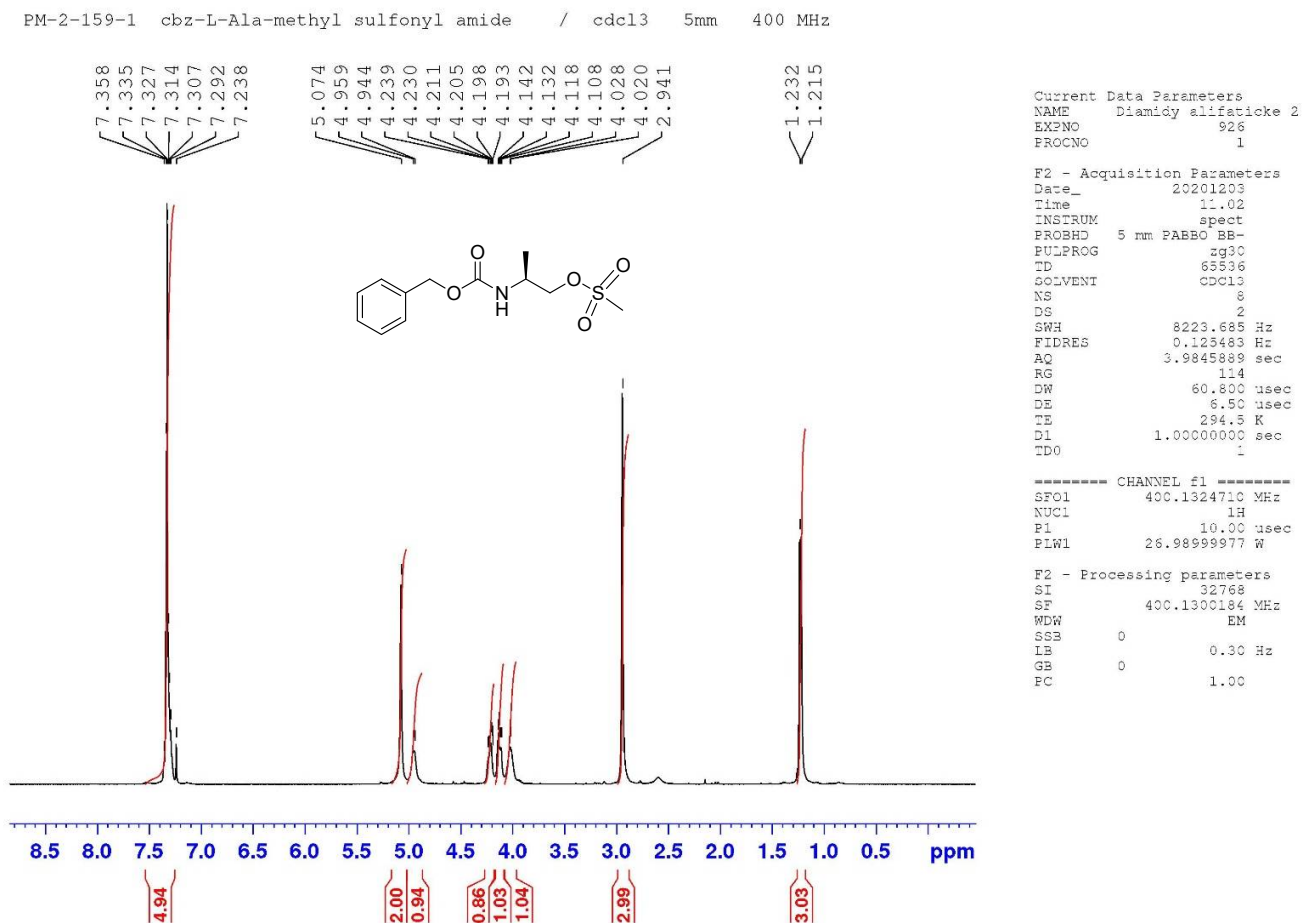
1.9. Copy of ^1H NMR of **3b**



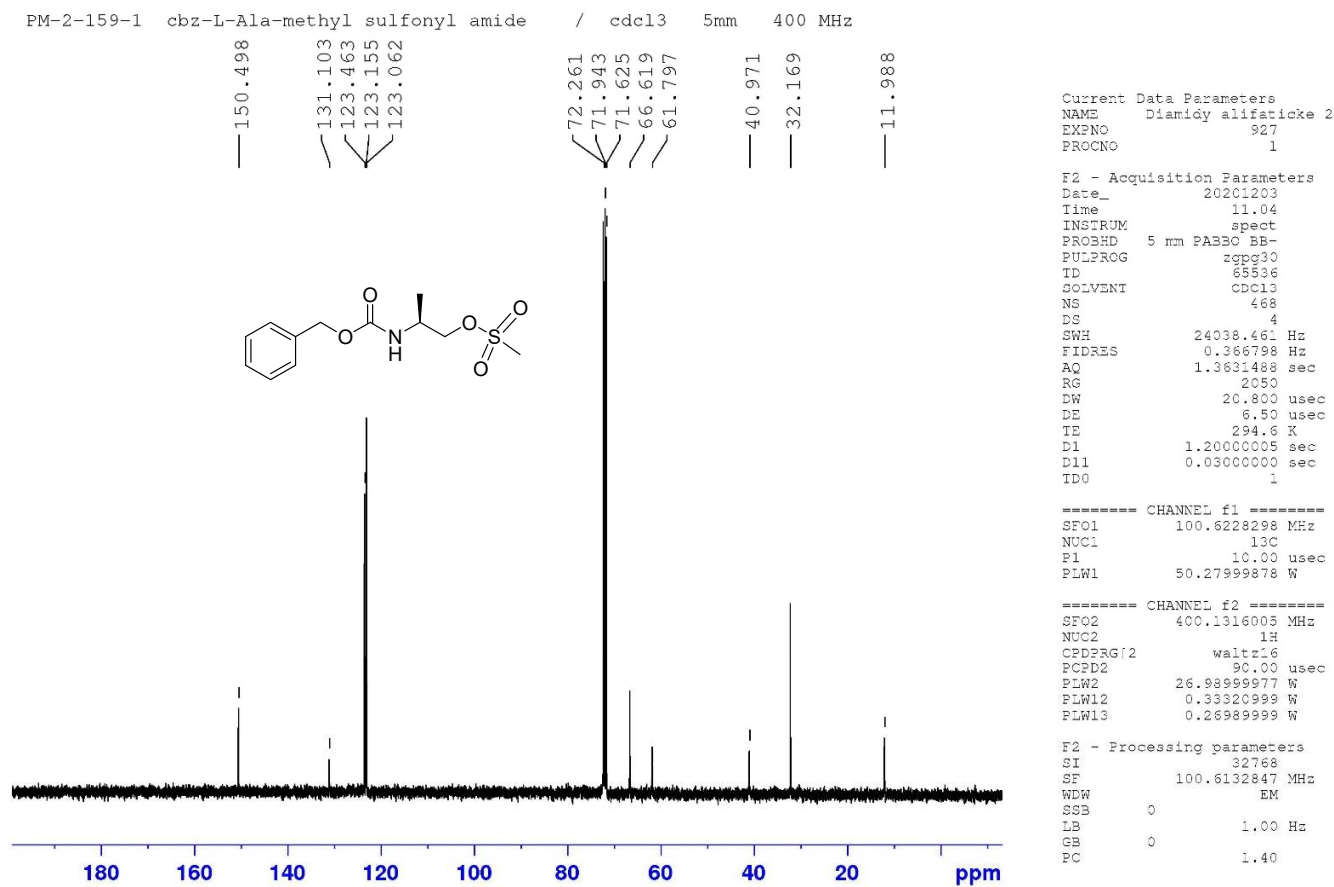
1.10. Copy of ^{13}C NMR of **3b**



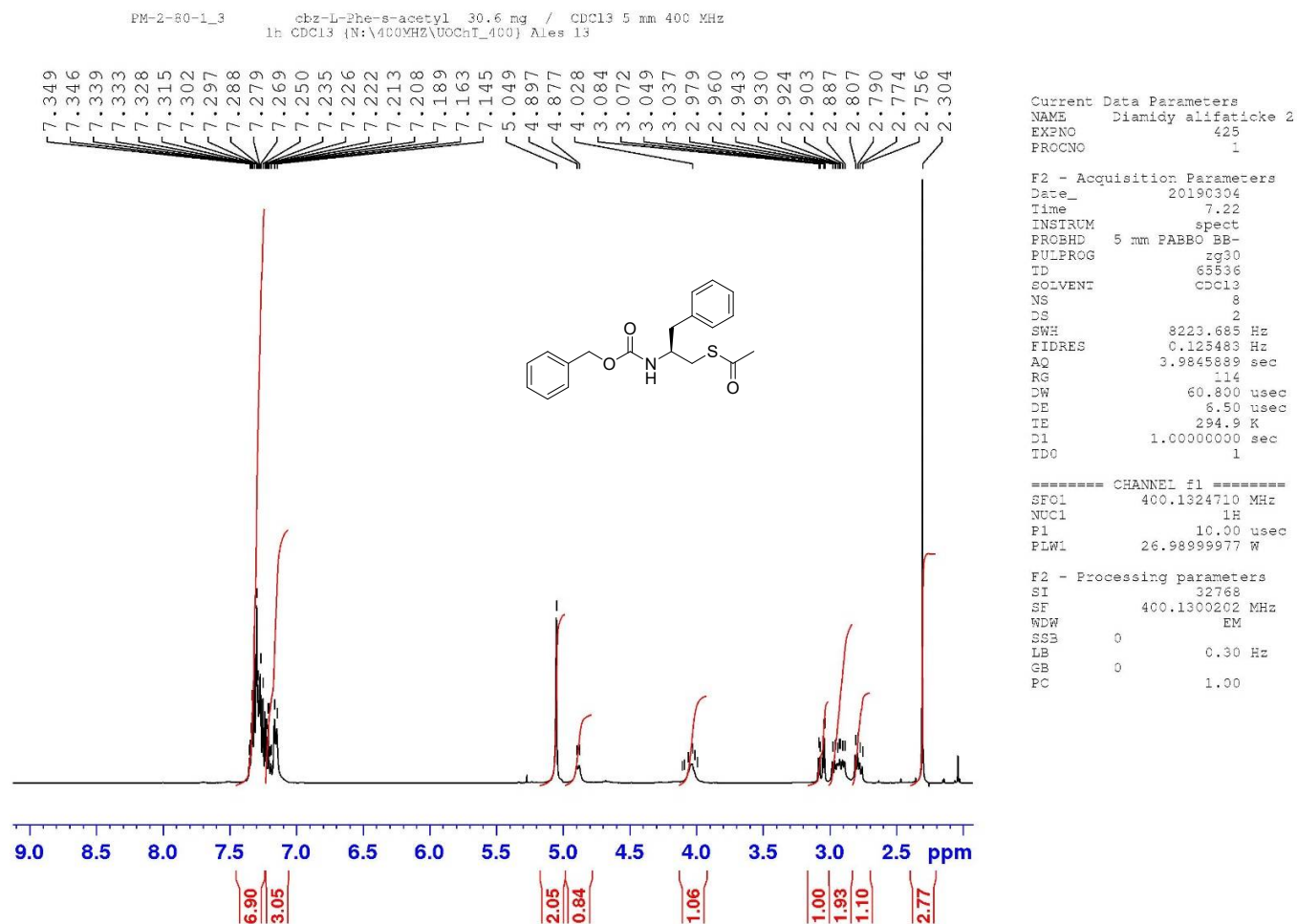
1.11. Copy of ^1H NMR of **3c**



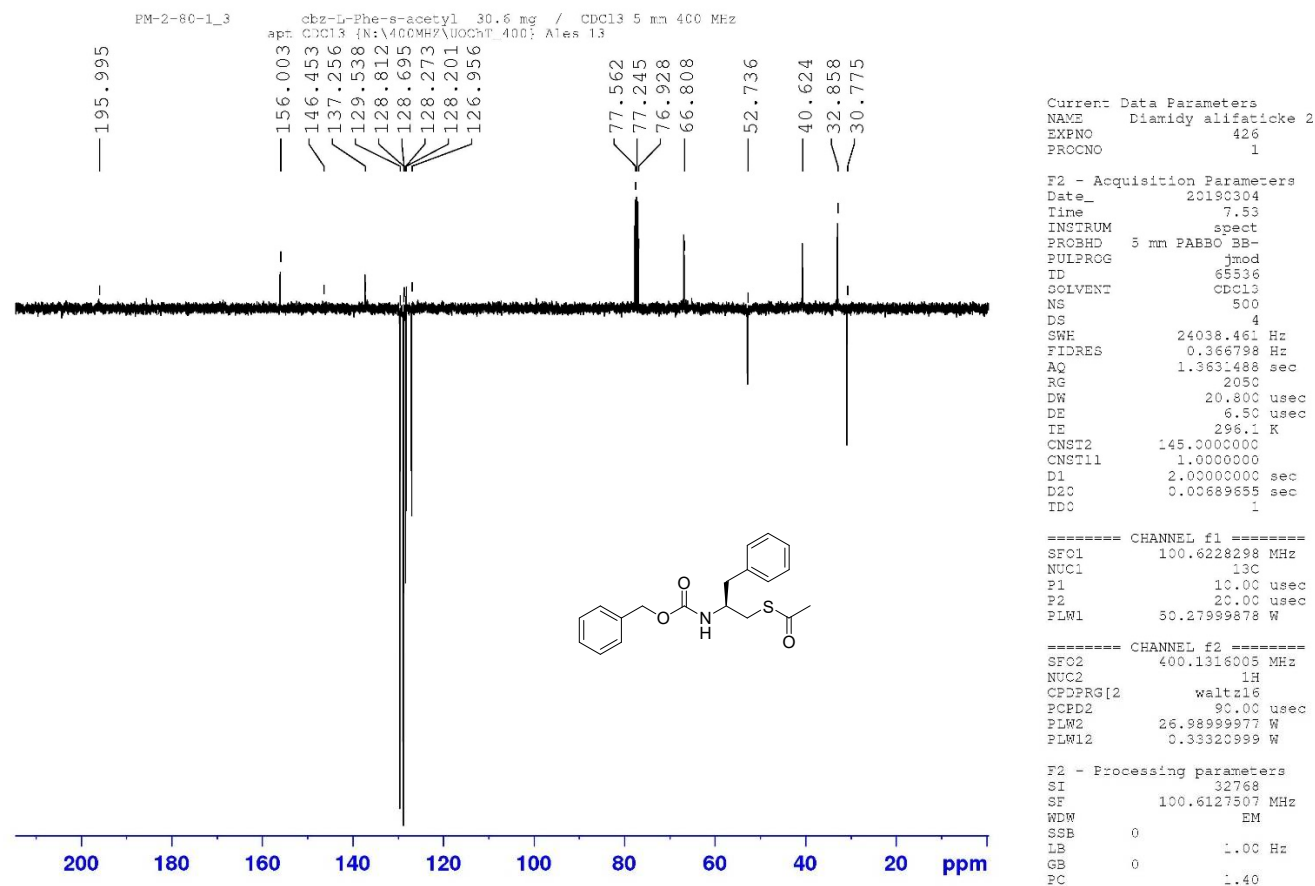
1.12. Copy of ^{13}C NMR of **3c**



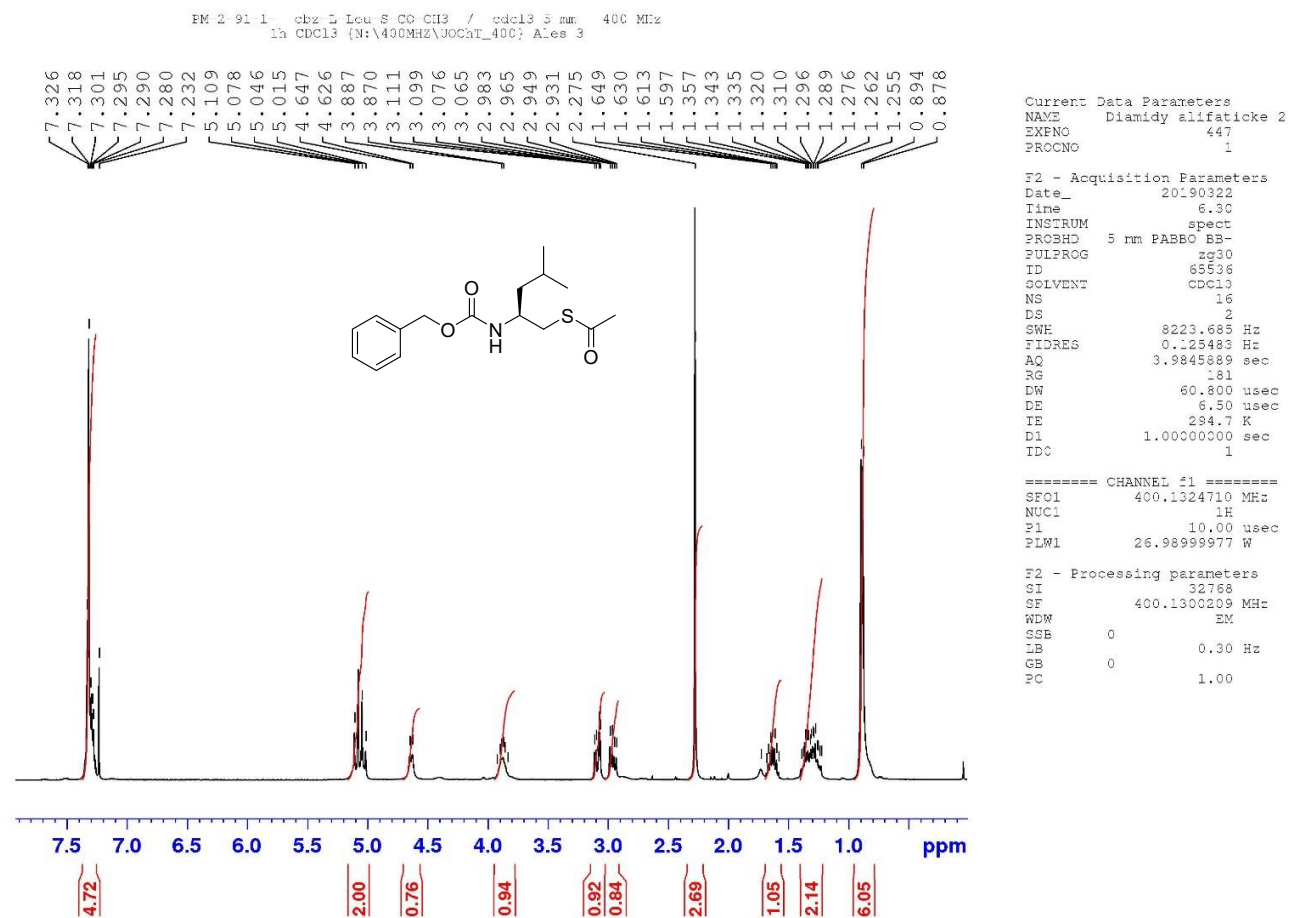
1.13. Copy of ^1H NMR of **4a**



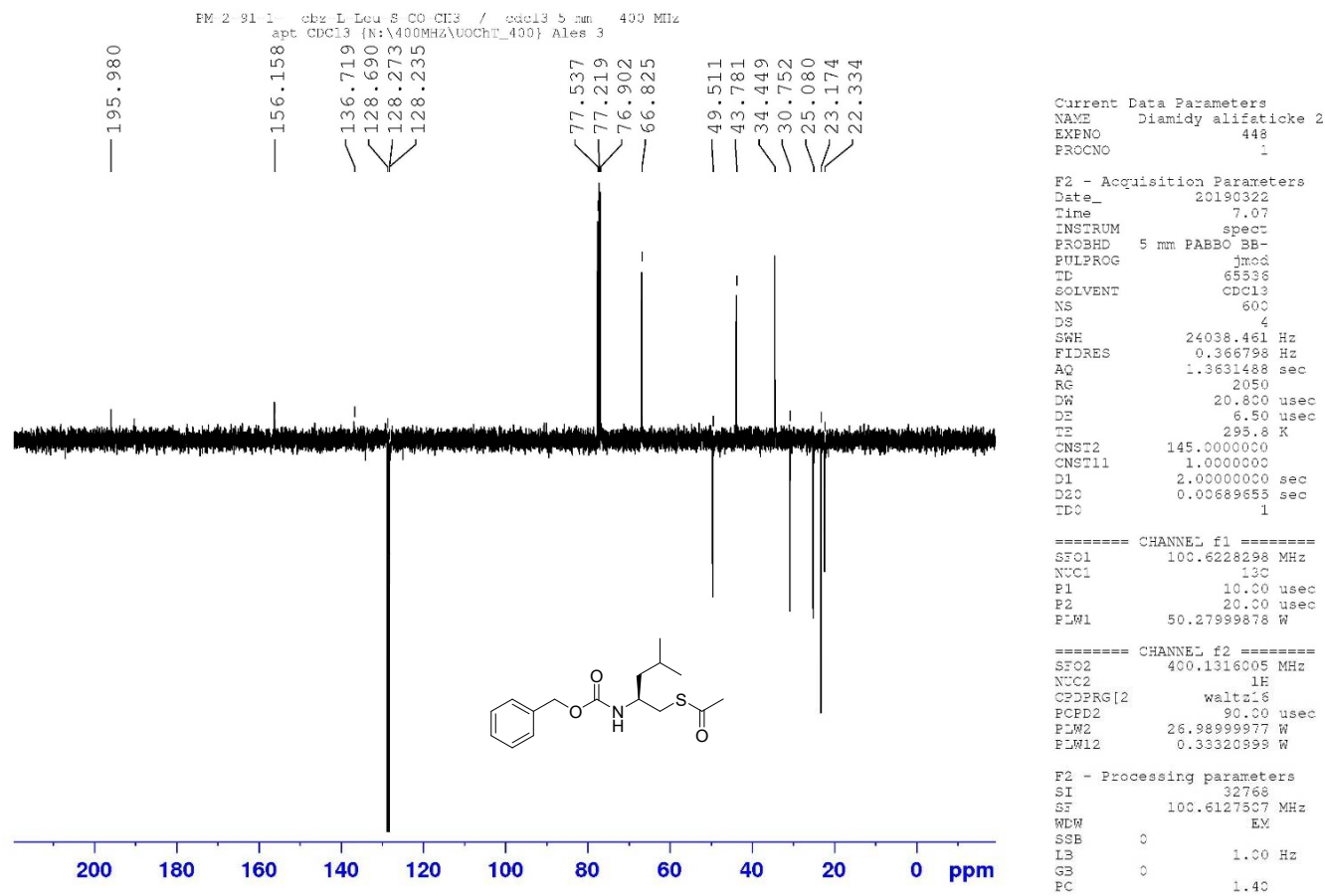
1.14. Copy of ^{13}C NMR of **4a**



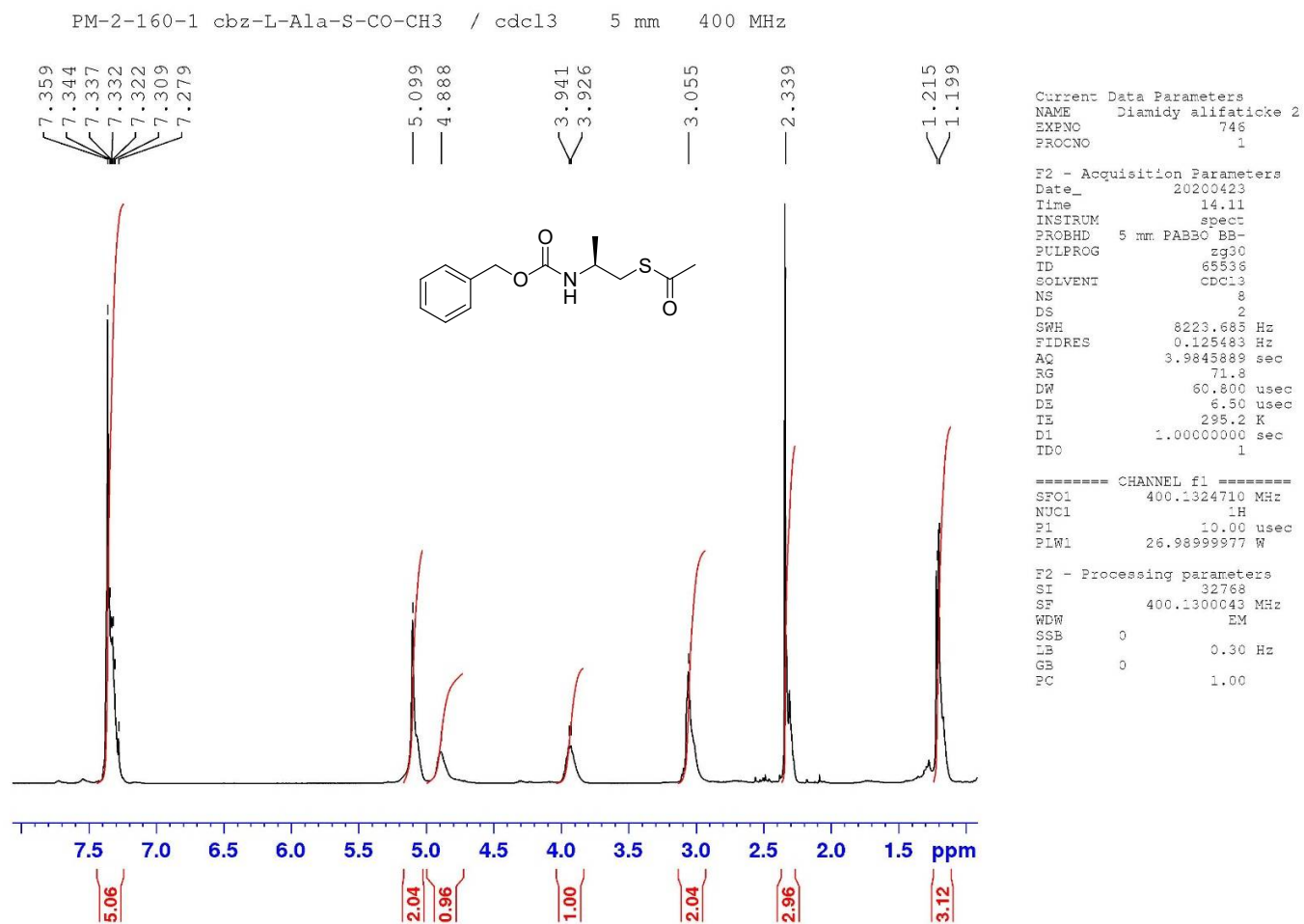
1.15. Copy of ¹H NMR of **4b**



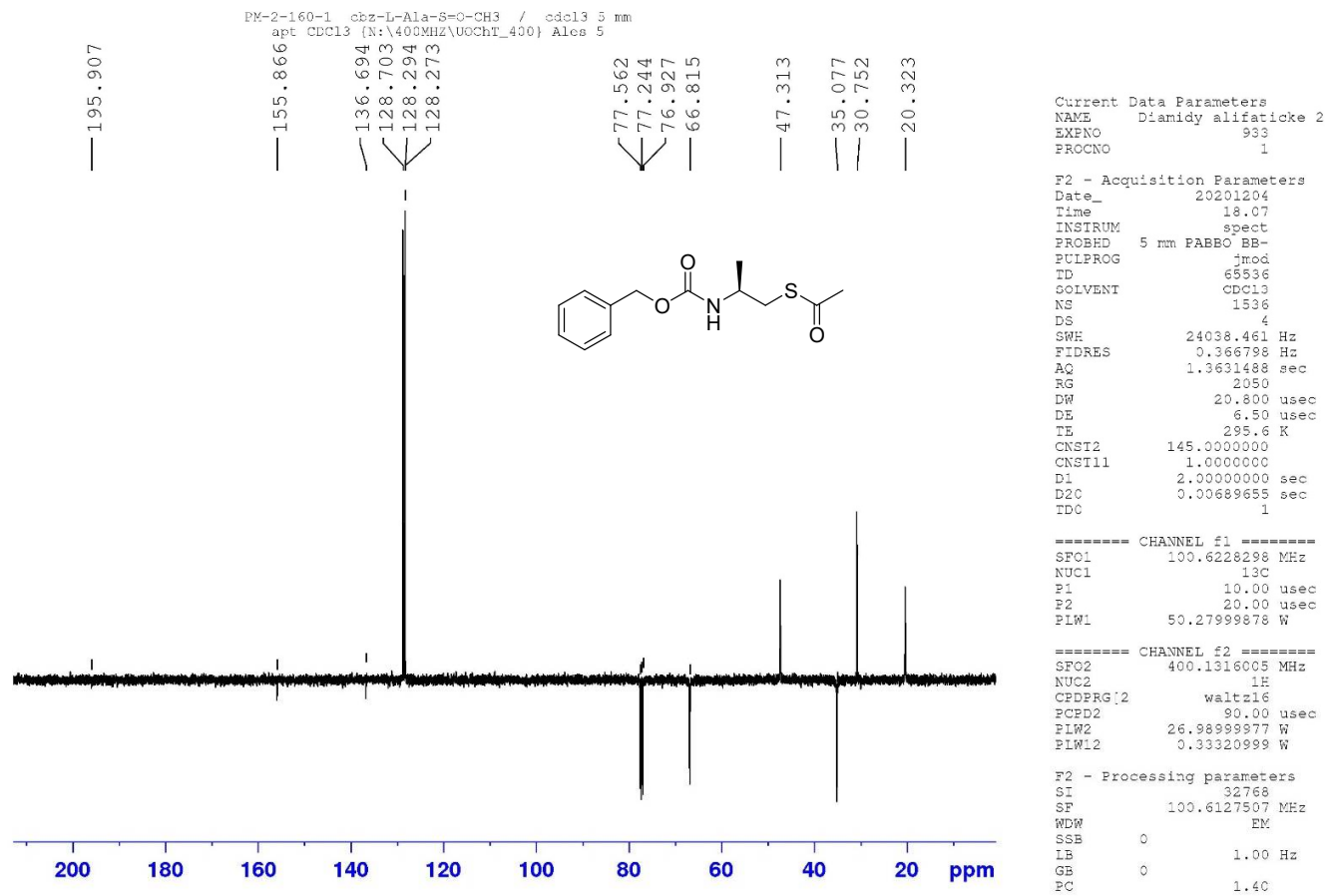
1.16. Copy of ^{13}C NMR of **4b**



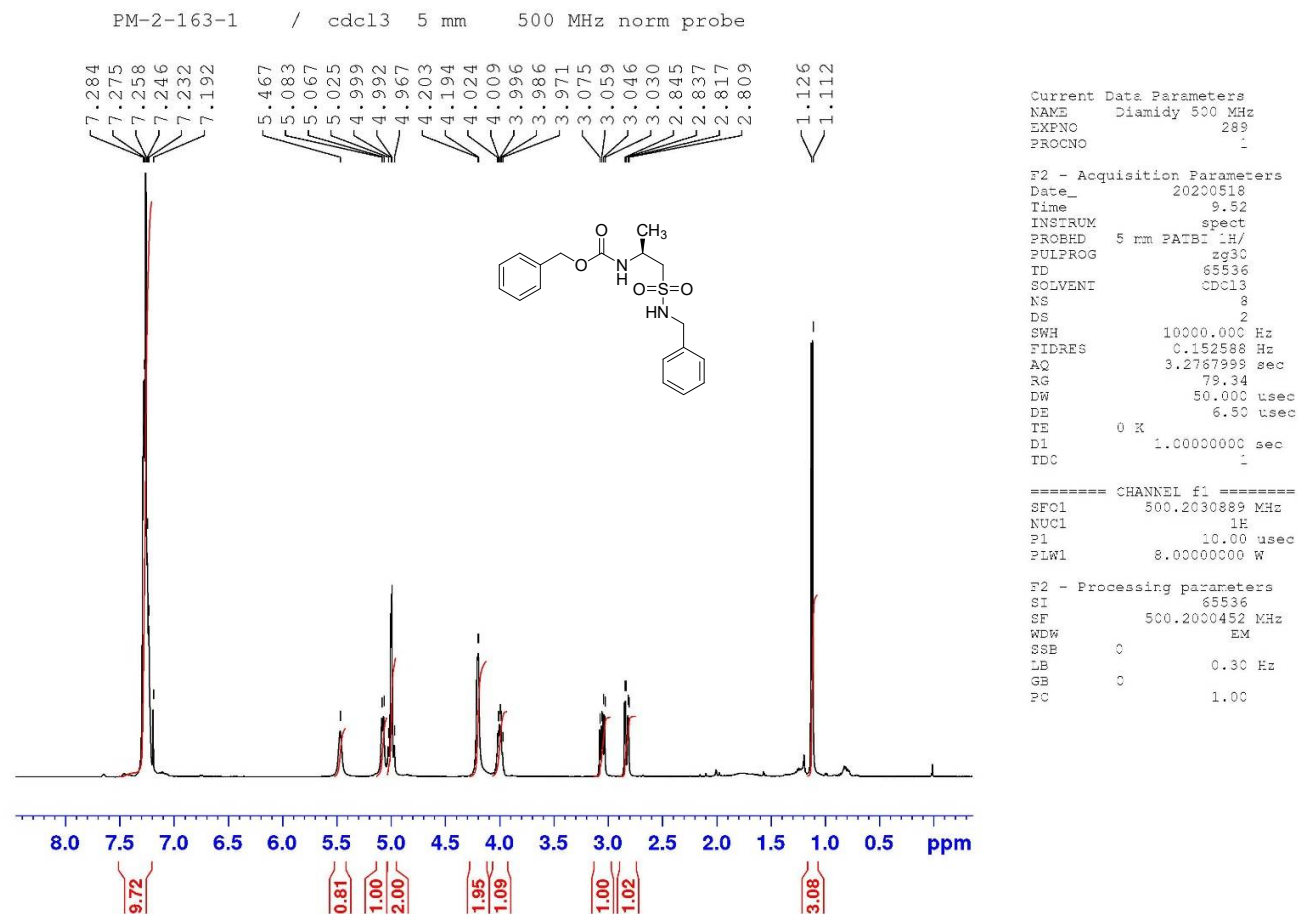
1.17. Copy of ^1H NMR of **4c**



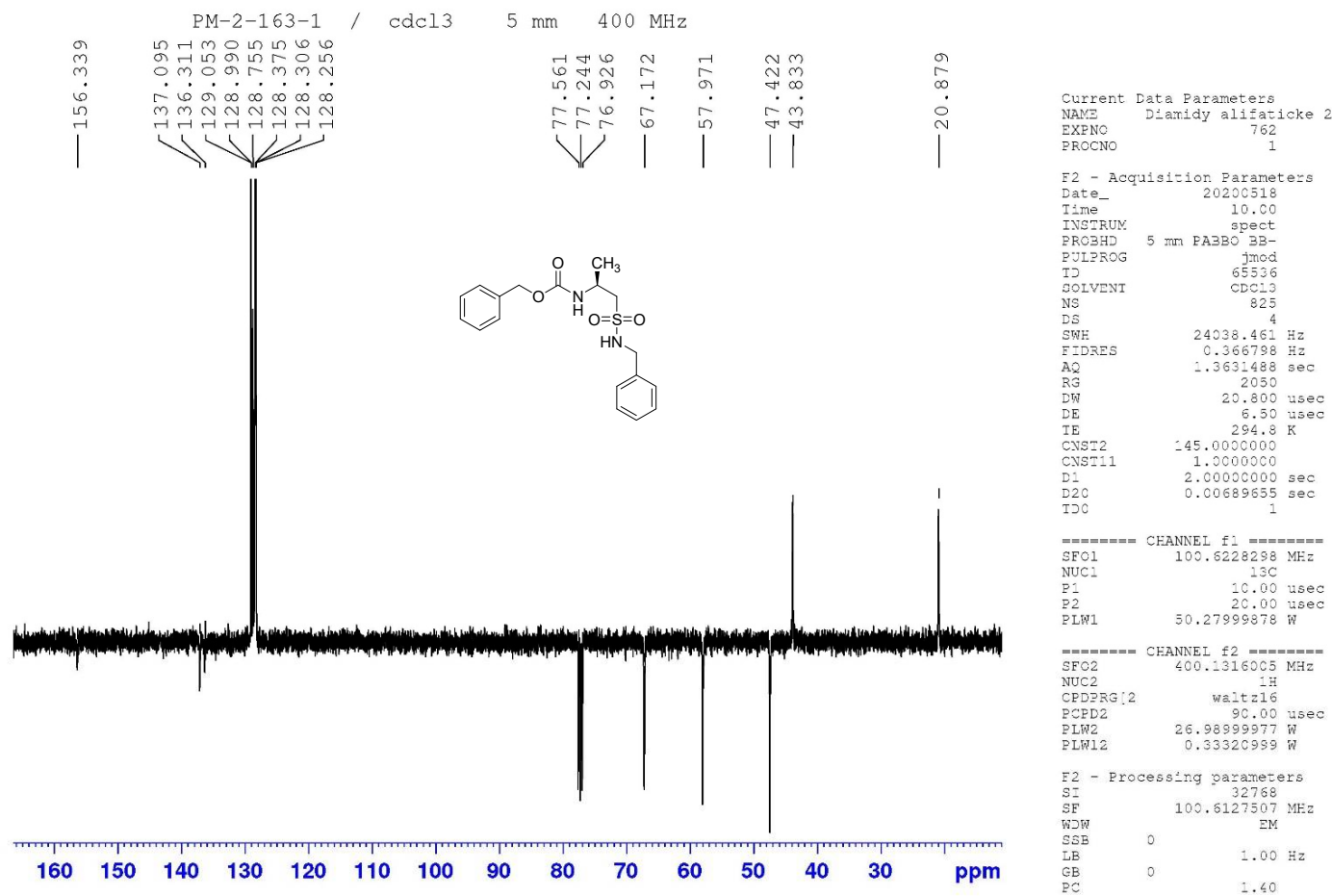
1.18. Copy of ^{13}C NMR of **4c**



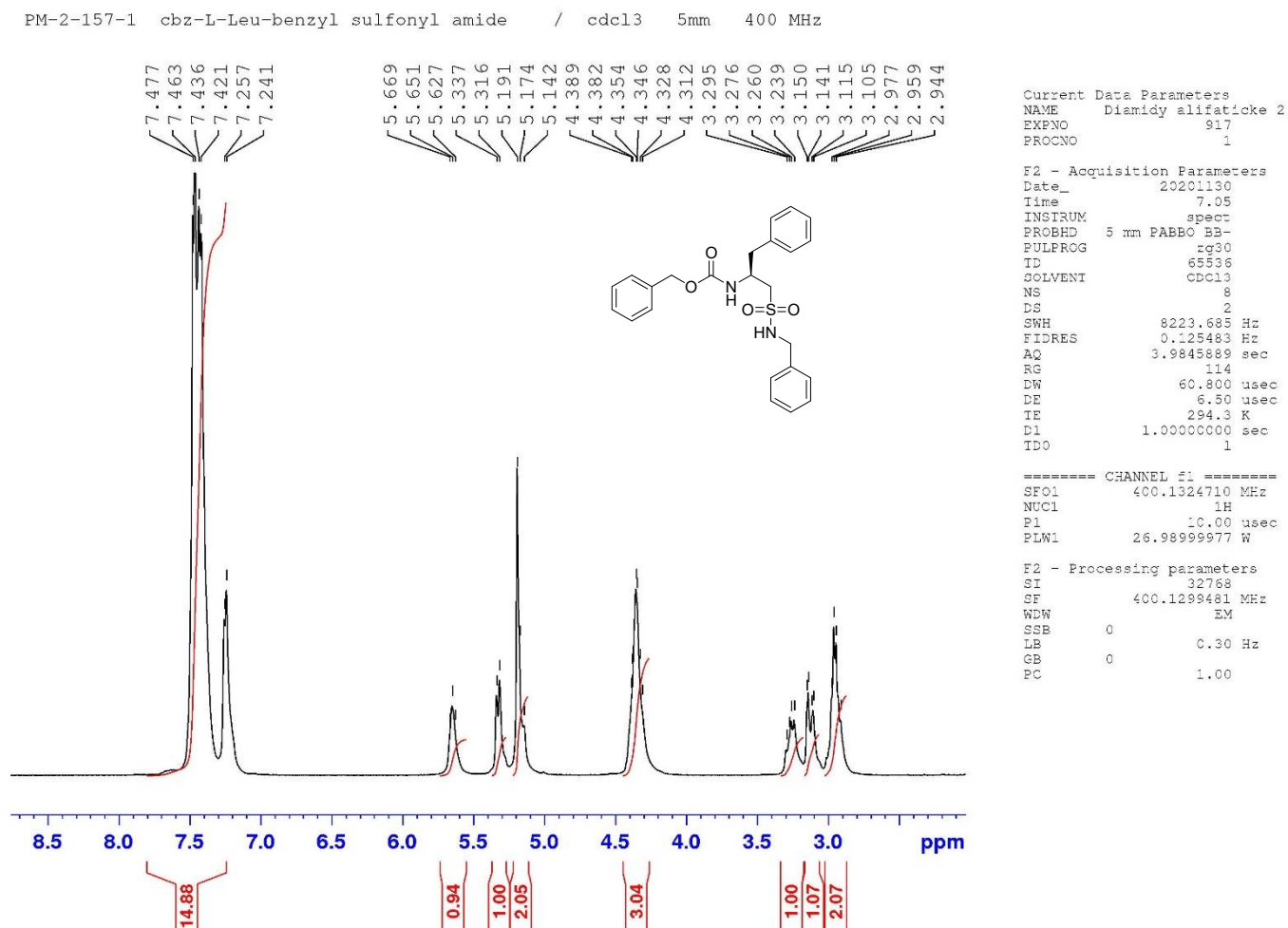
1.19. Copy of ¹H NMR of **5a**



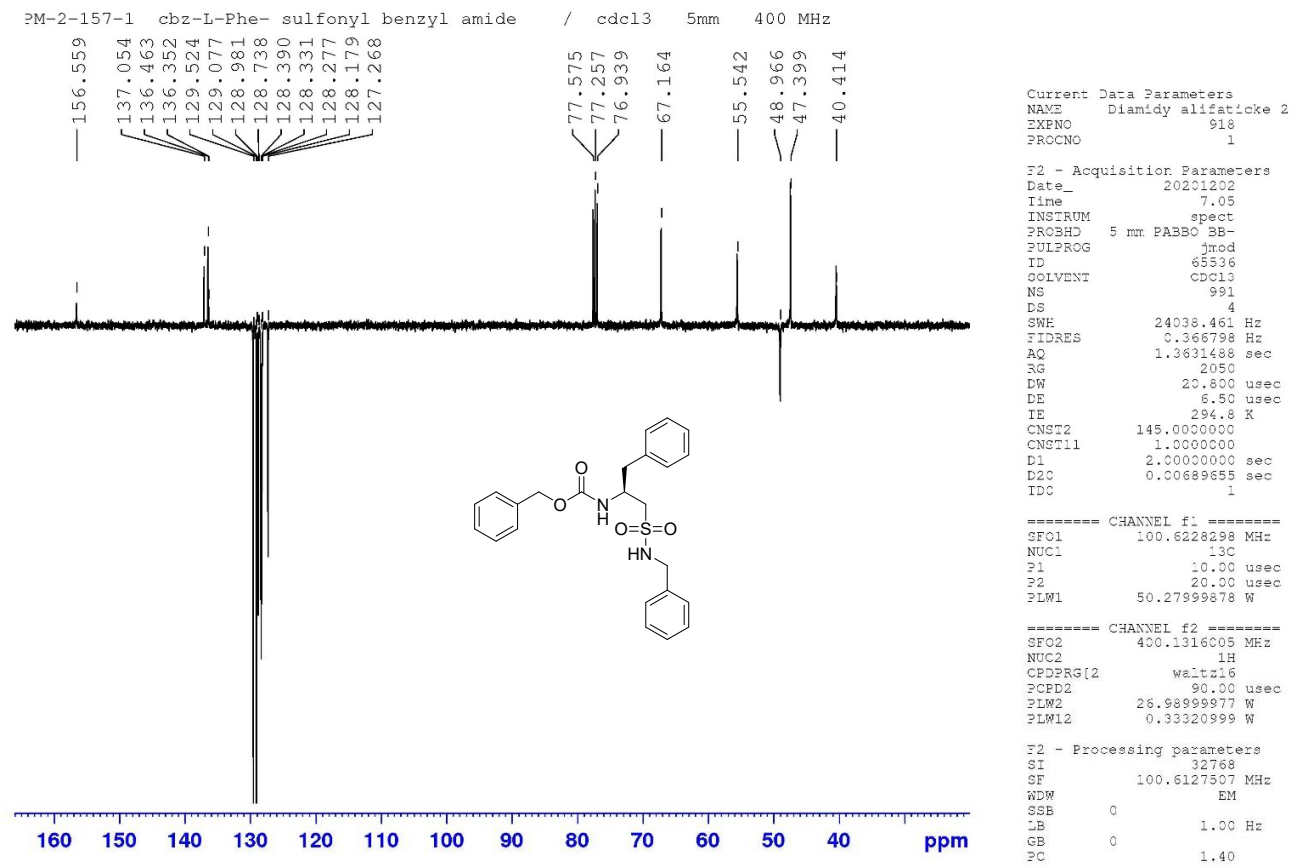
1.20. Copy of ^{13}C NMR of **5a**



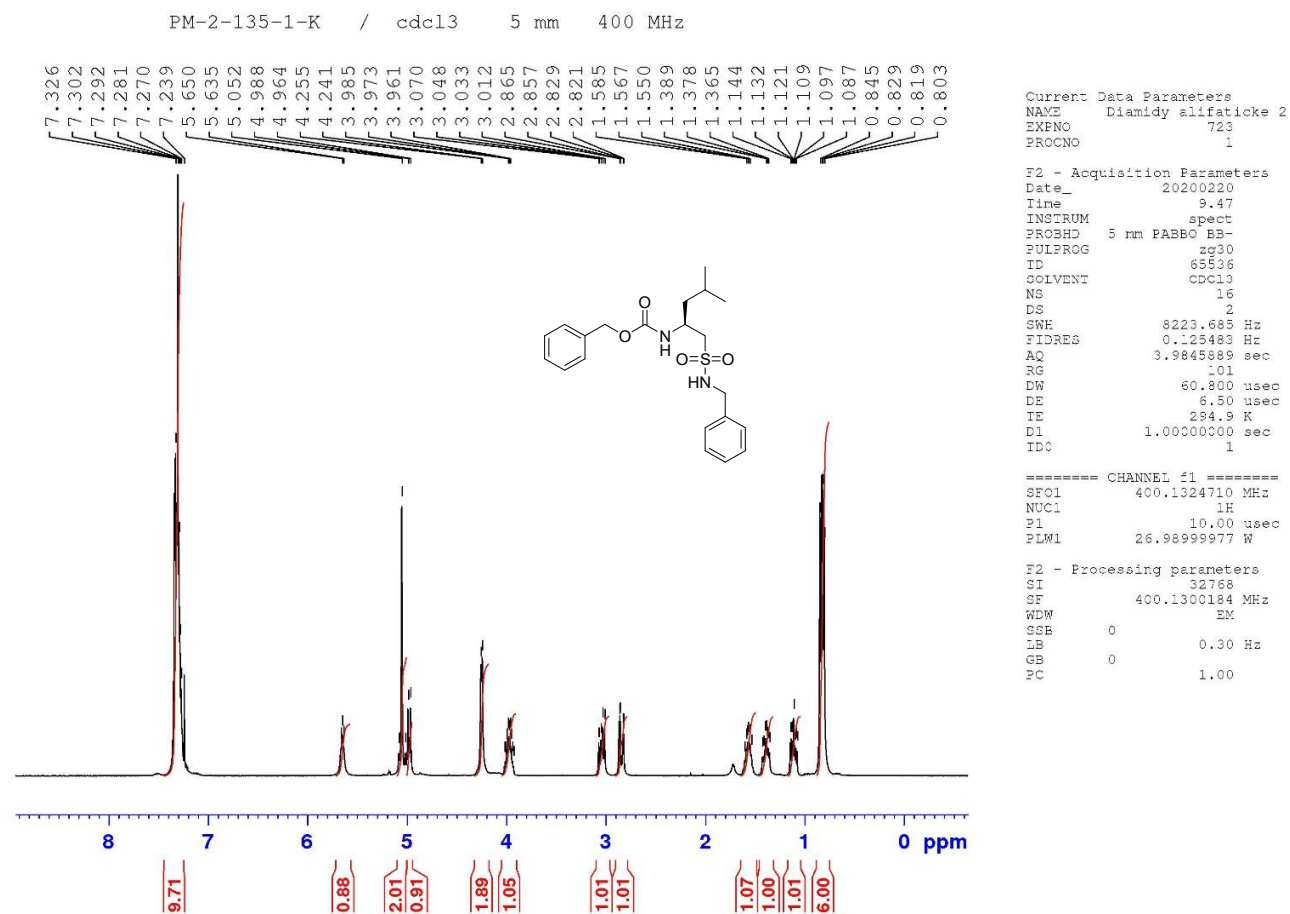
1.21. Copy of ^1H NMR of **5b**



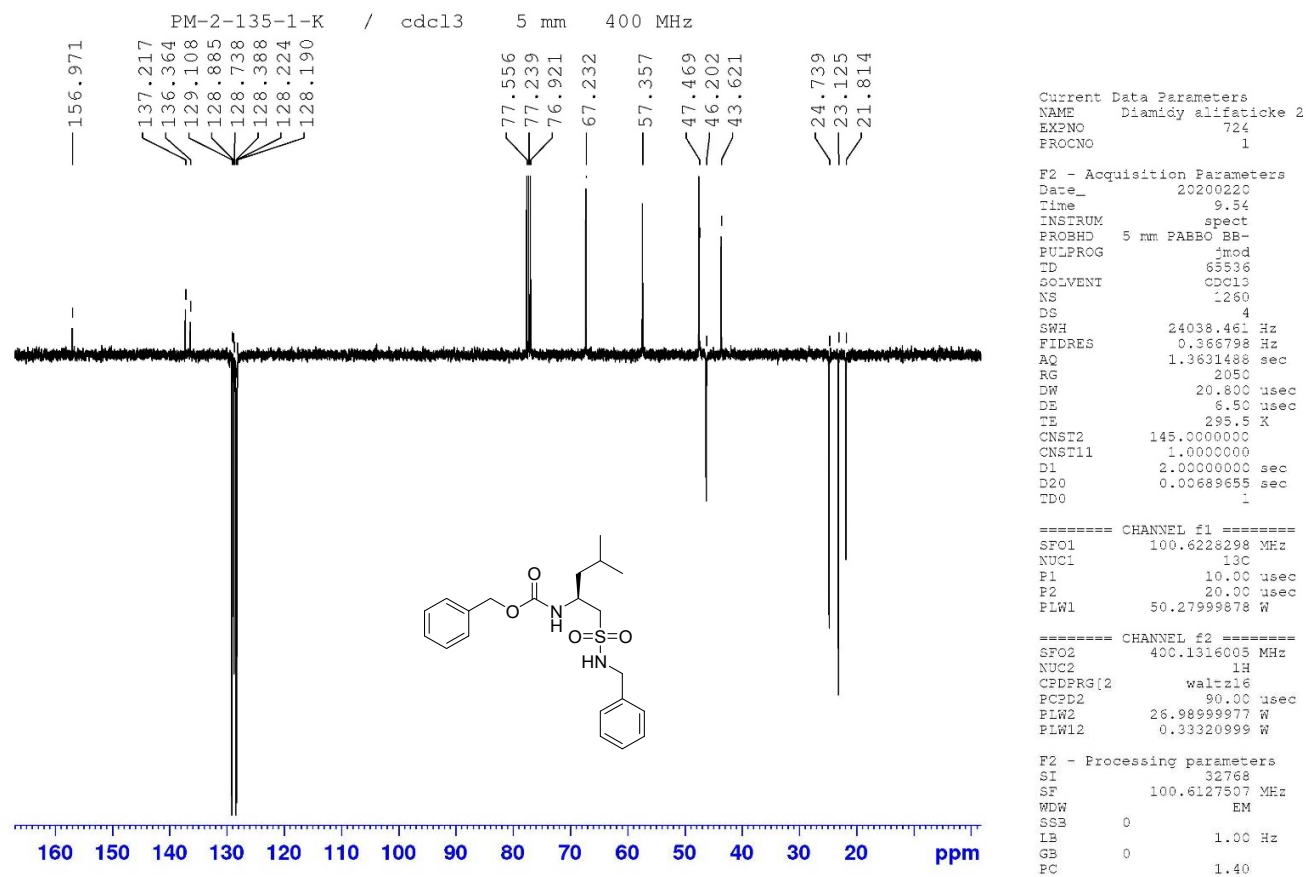
1.22. Copy of ^{13}C NMR of **5b**



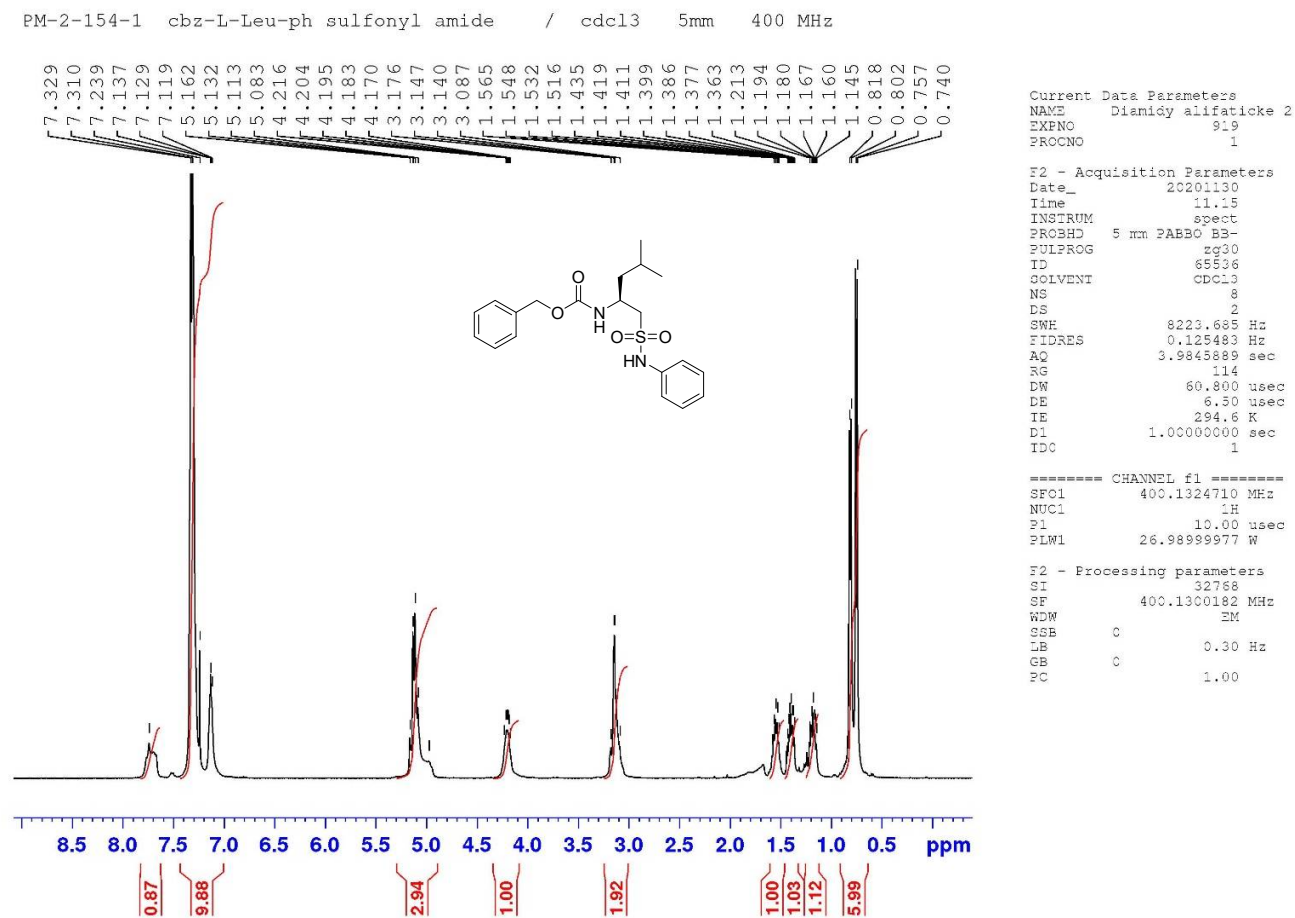
1.23. Copy of ^1H NMR of **5c**



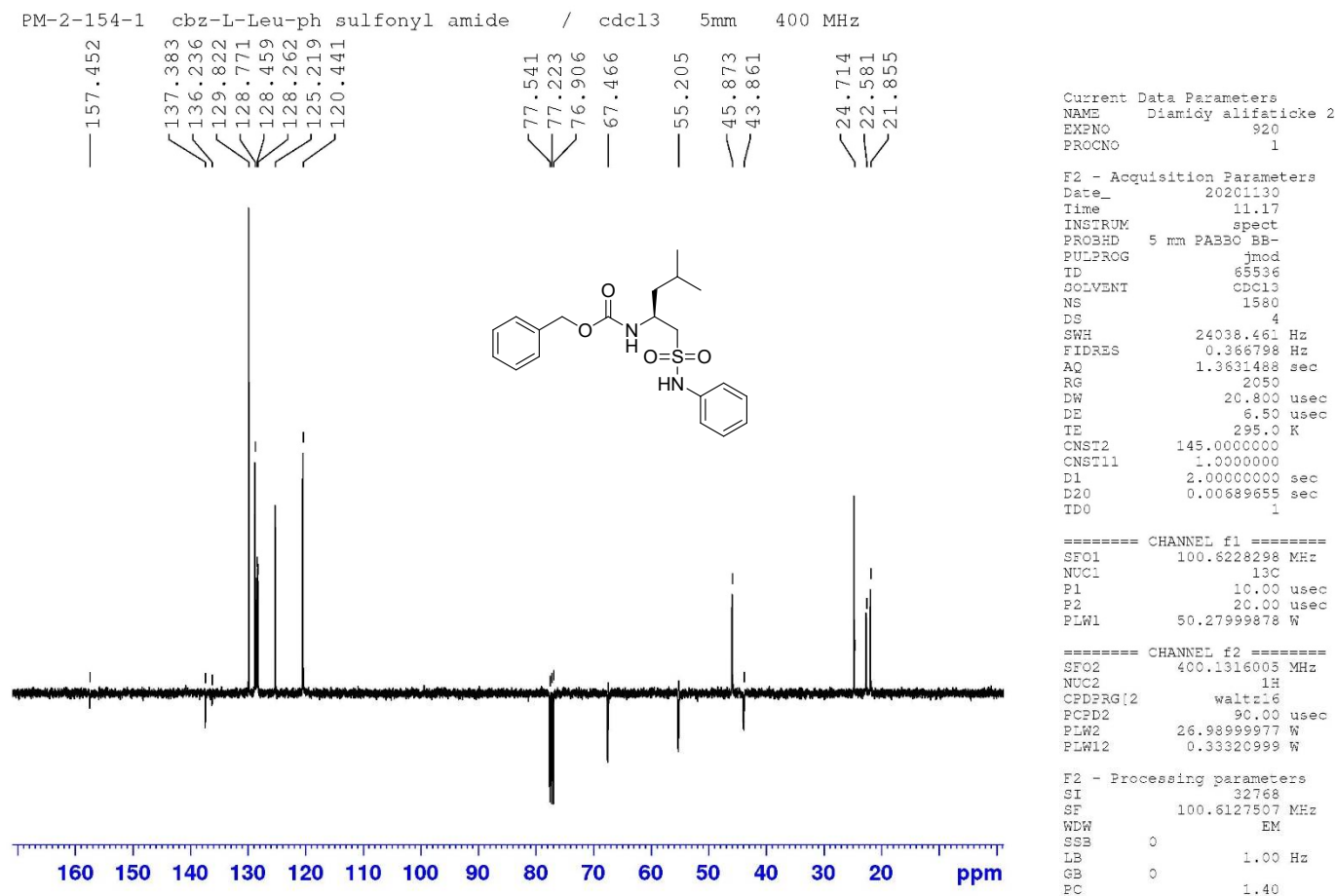
1.24. Copy of ^{13}C NMR of **5c**



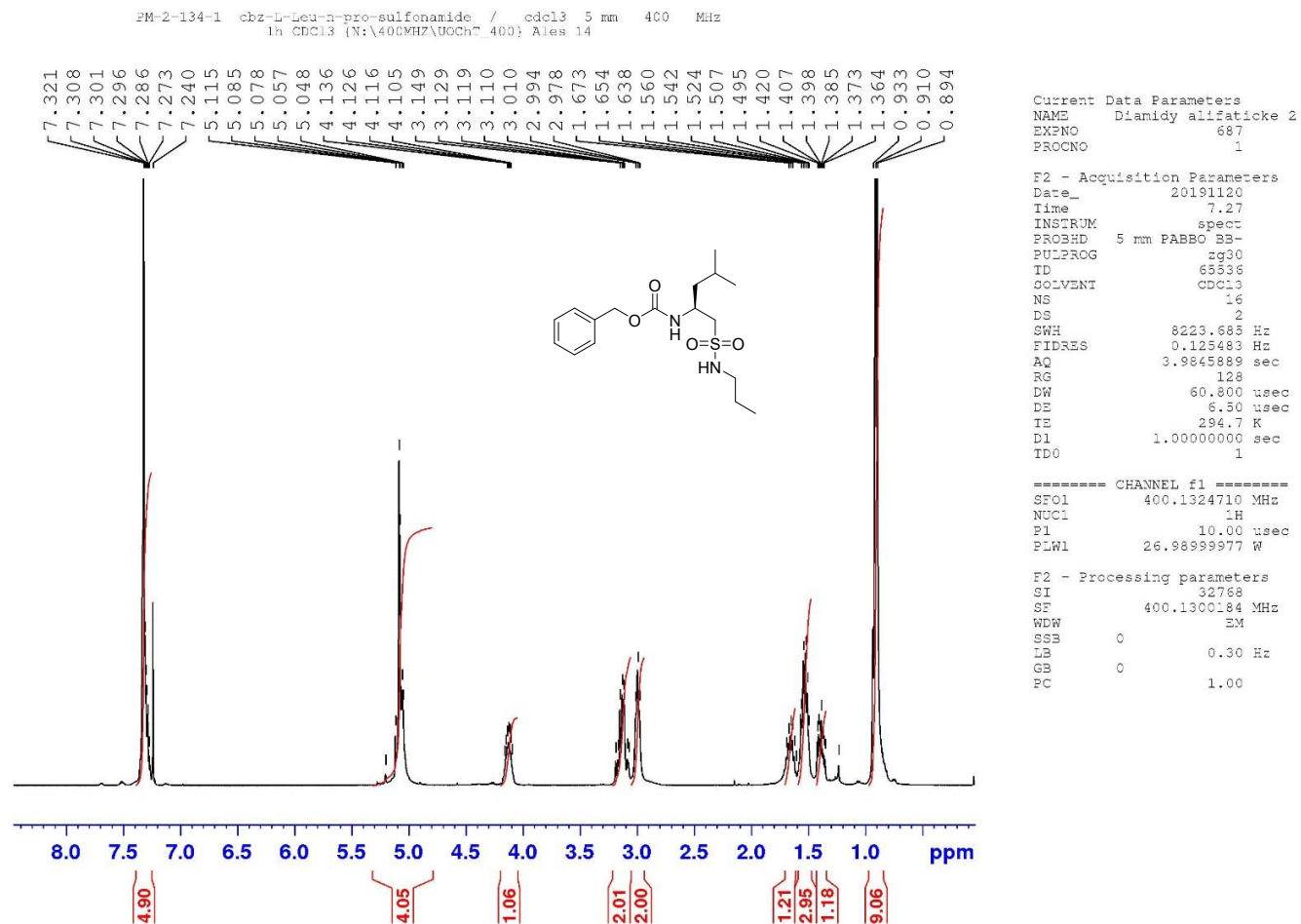
1.25. Copy of ¹H NMR of **5d**



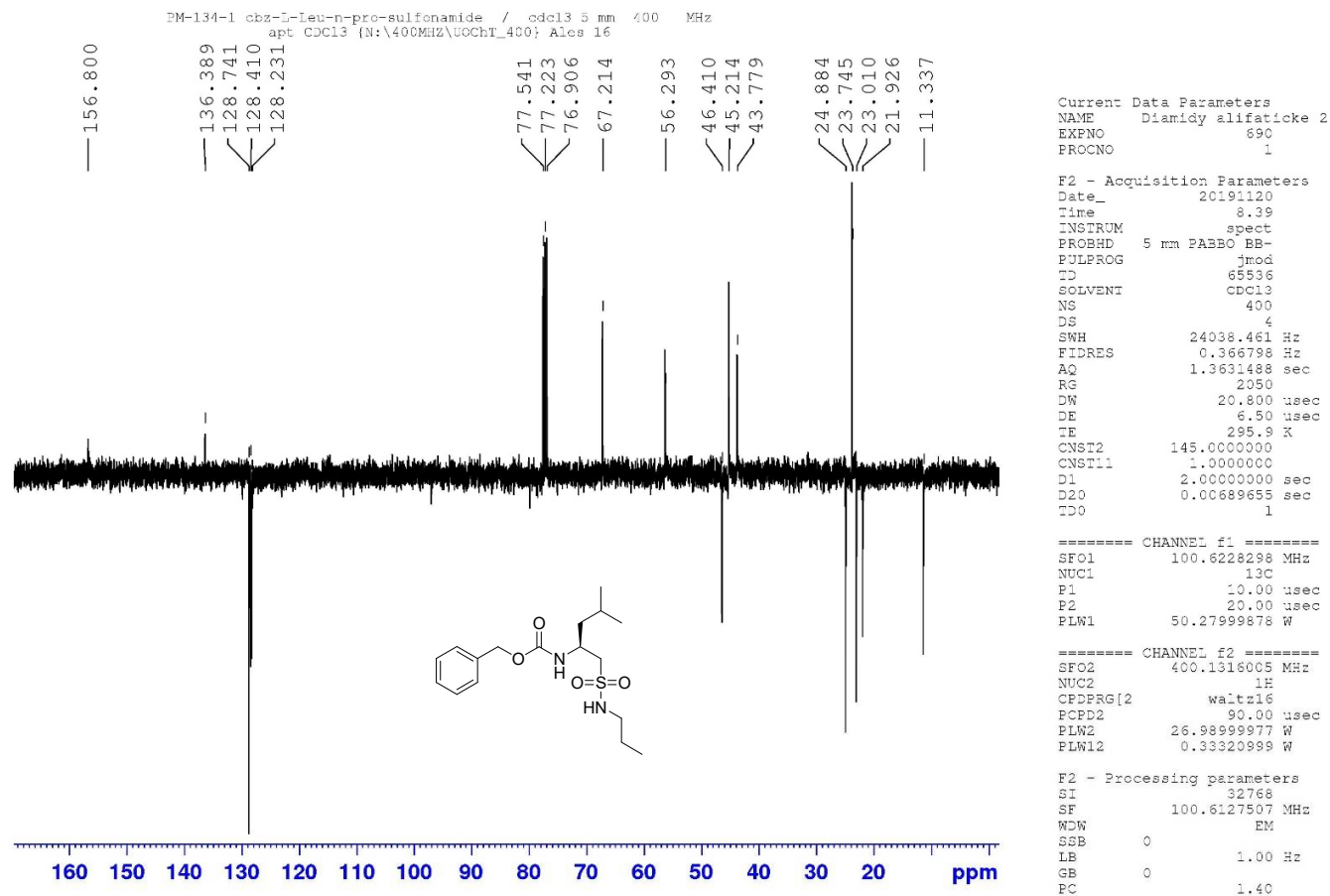
1.26. Copy of ^{13}C NMR of **5d**



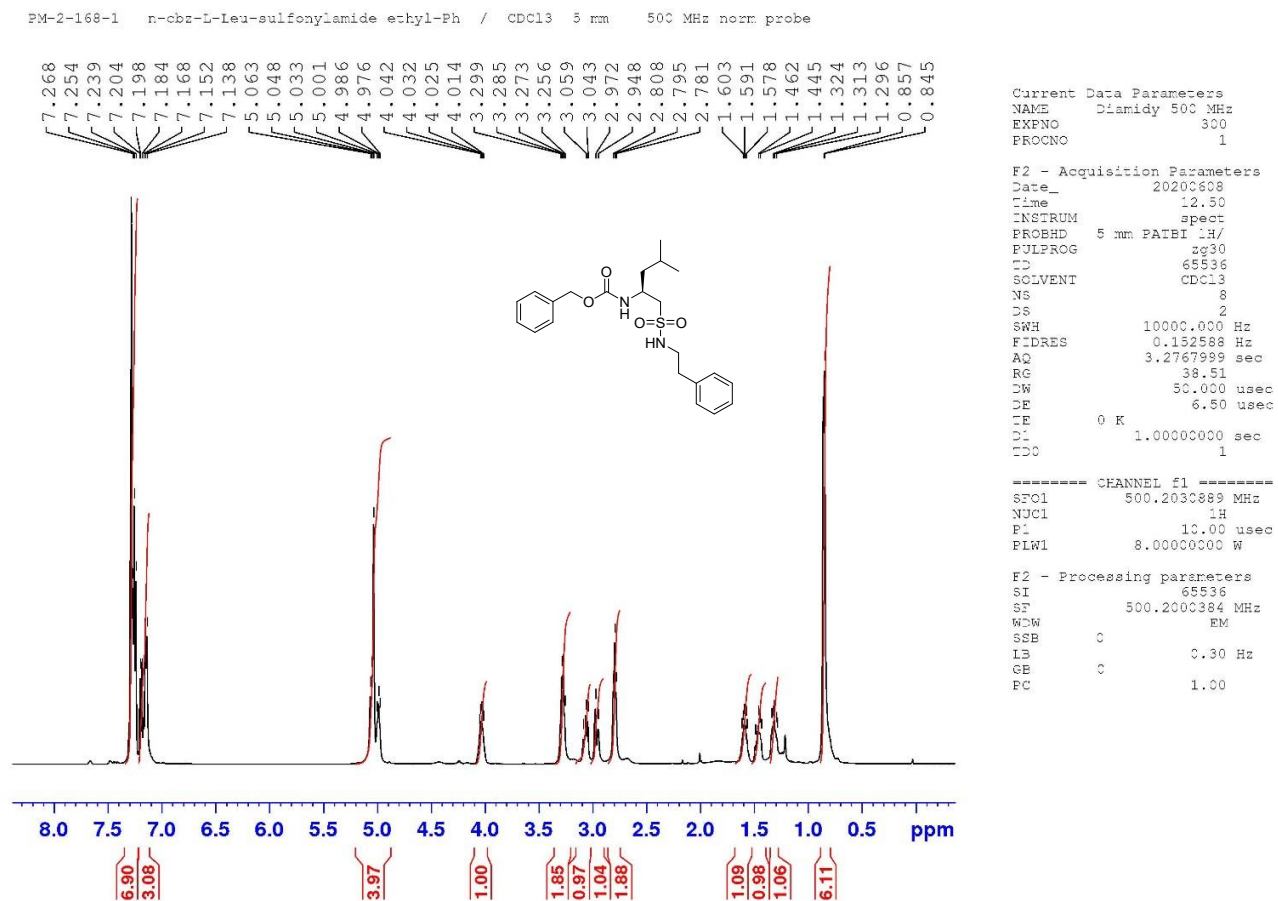
1.27. Copy of ^1H NMR of **5e**



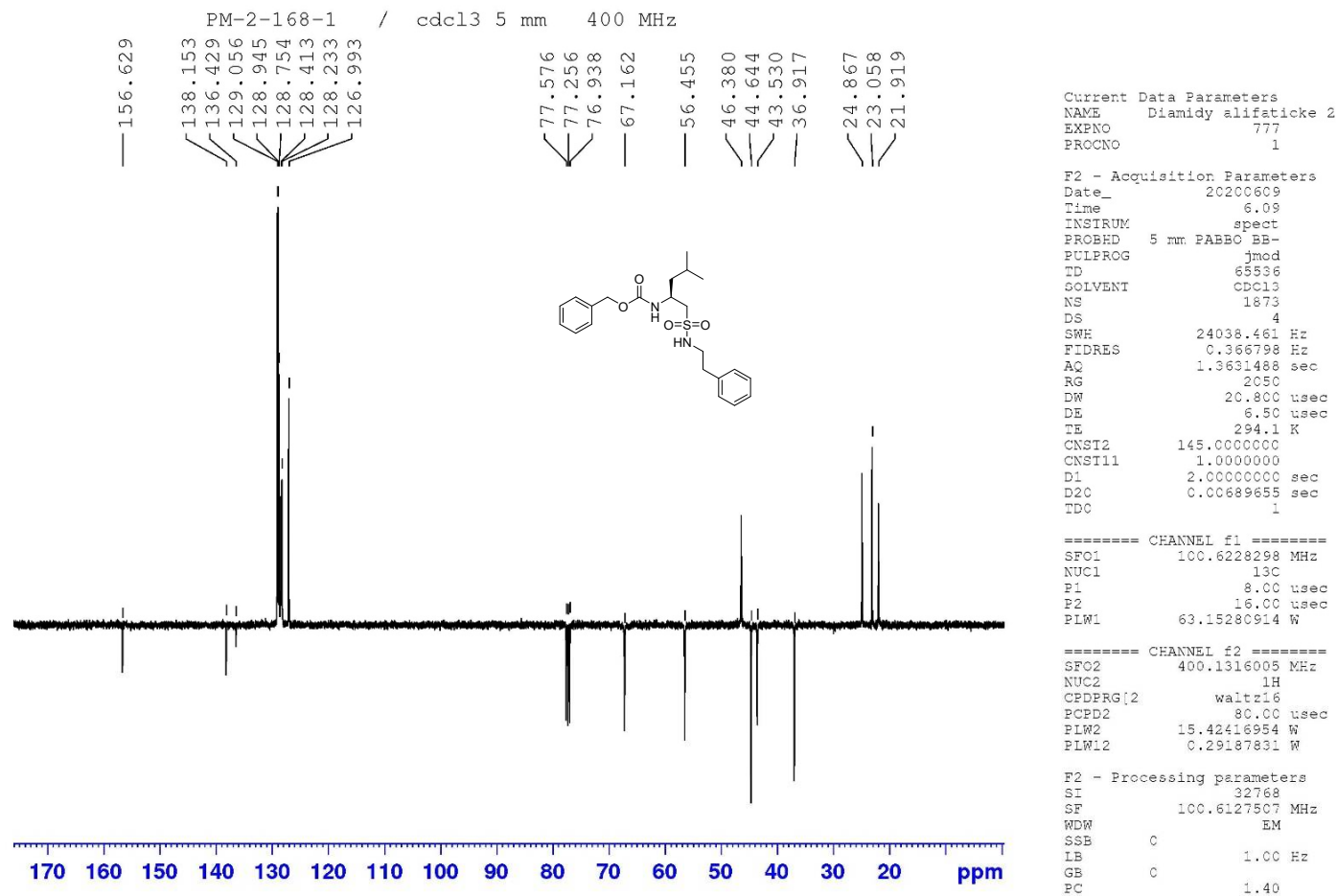
1.28. Copy of ^{13}C NMR of **5e**



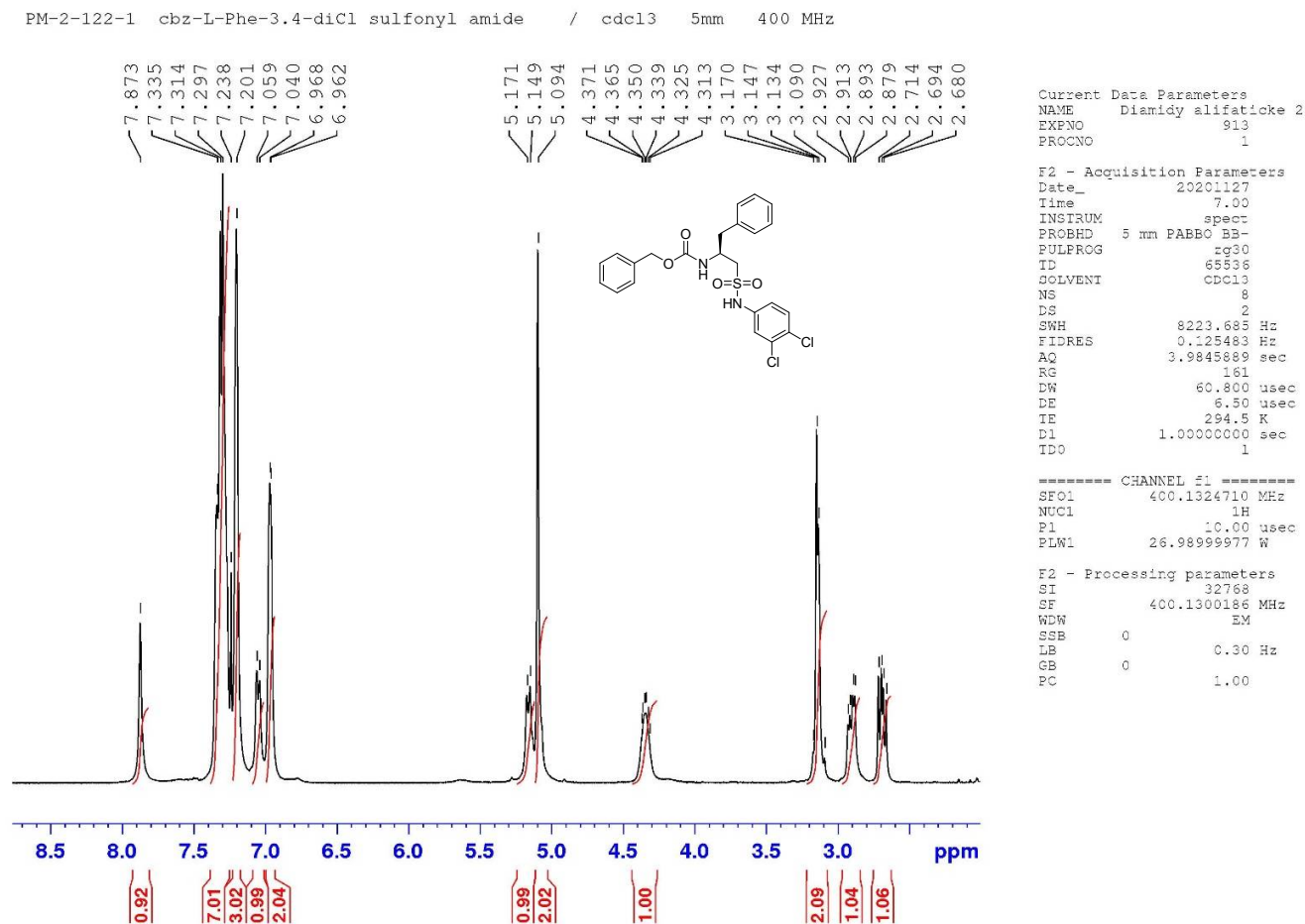
1.29. Copy of ^1H NMR of **5f**



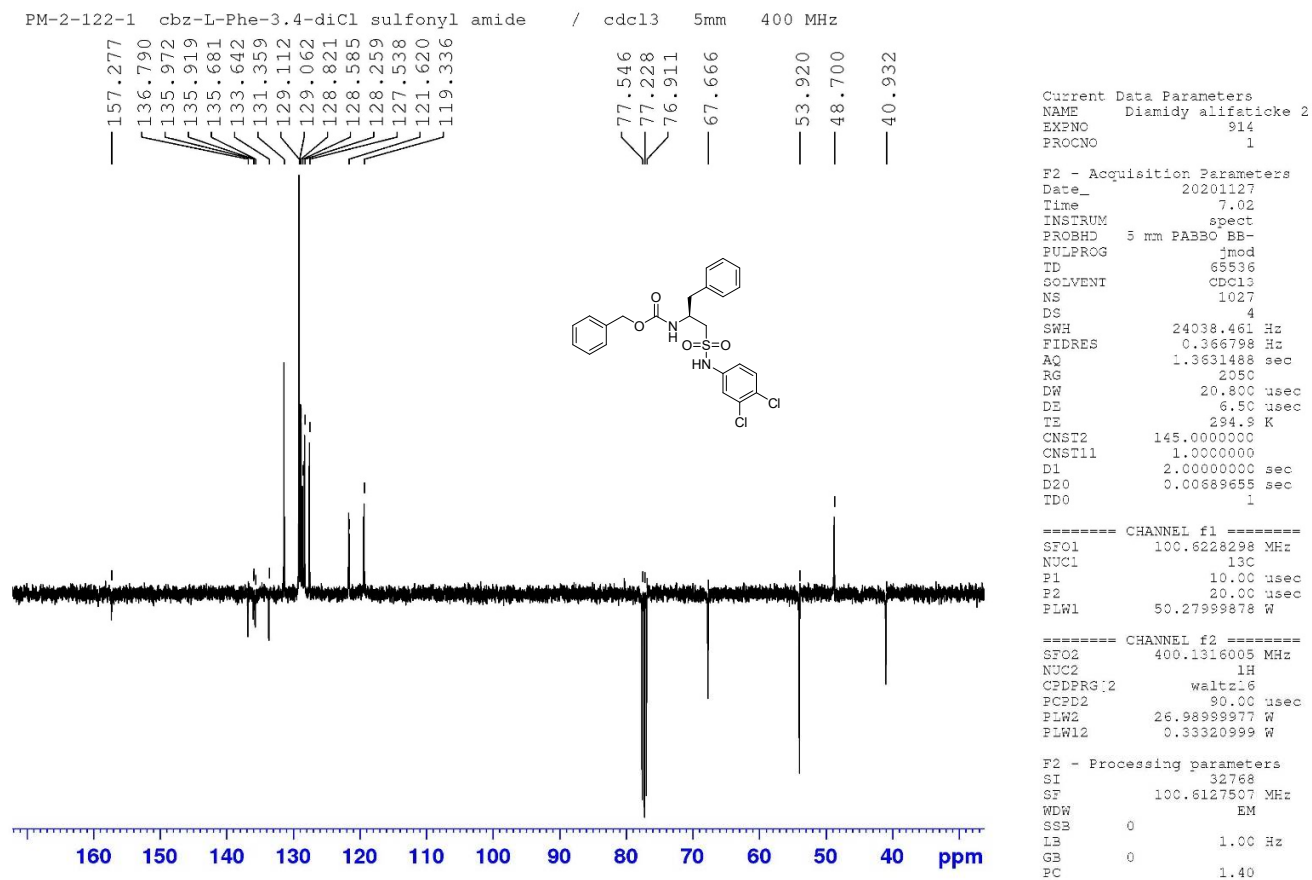
1.30. Copy of ^{13}C NMR of **5f**



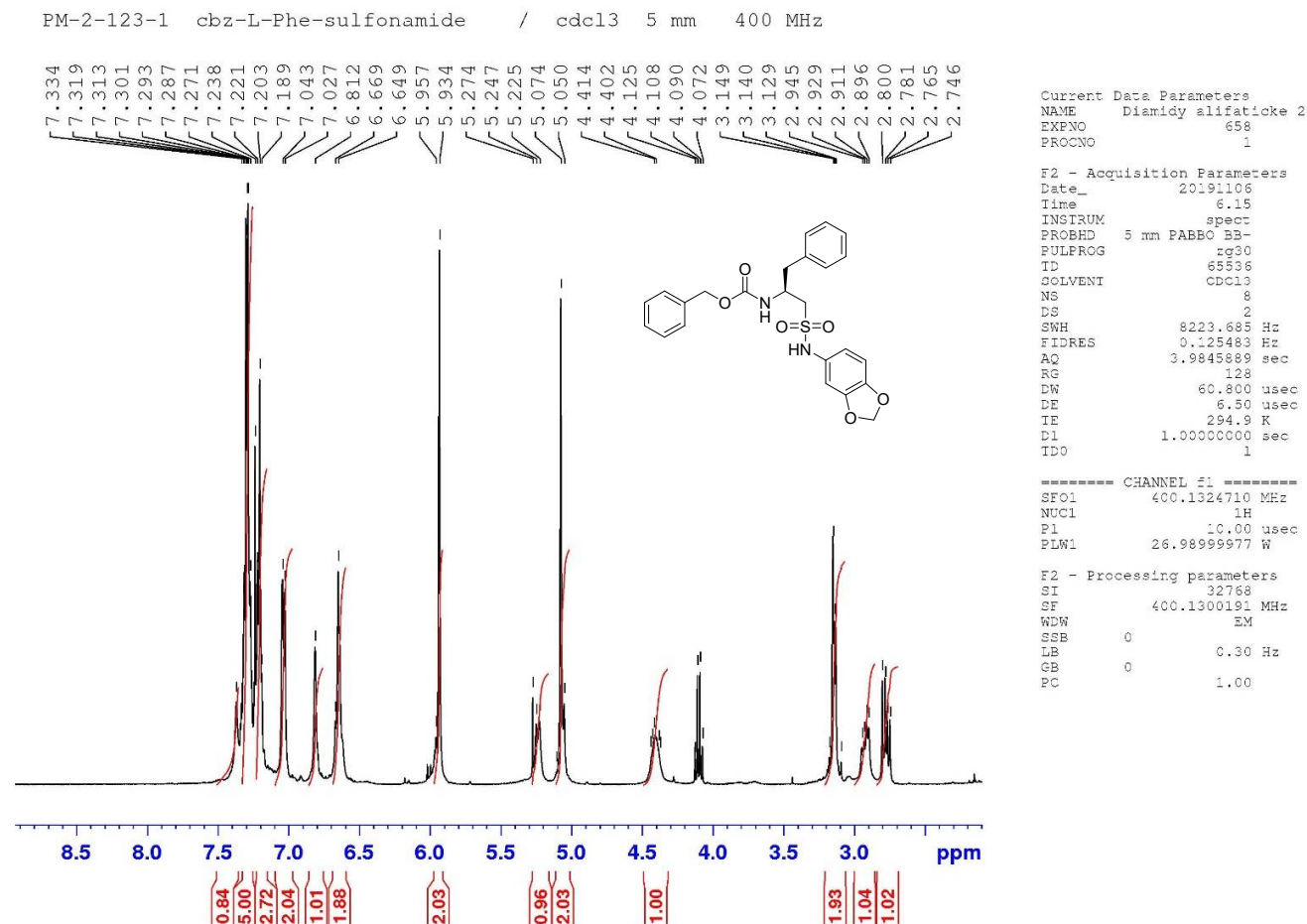
1.31. Copy of ^1H NMR of **5g**



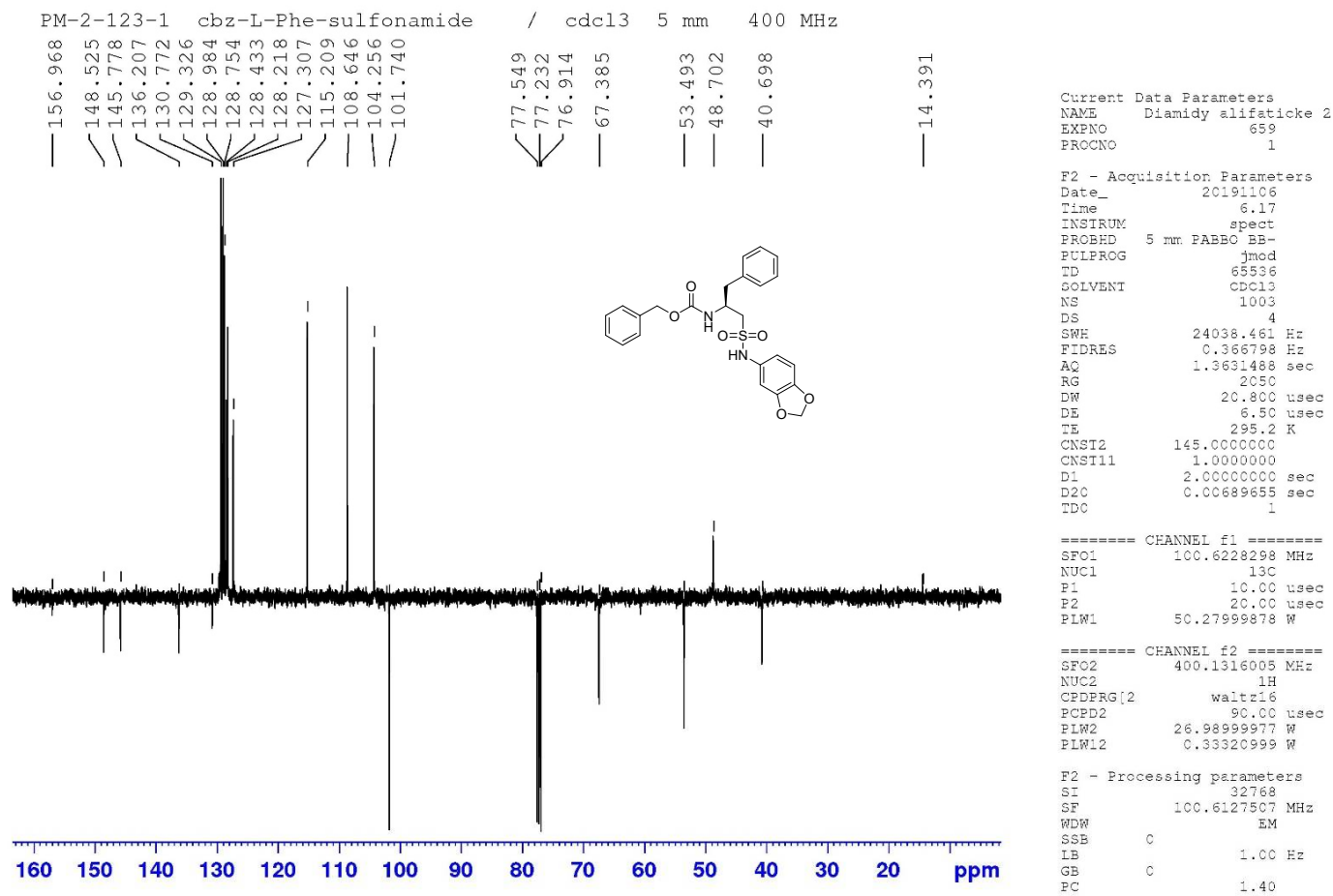
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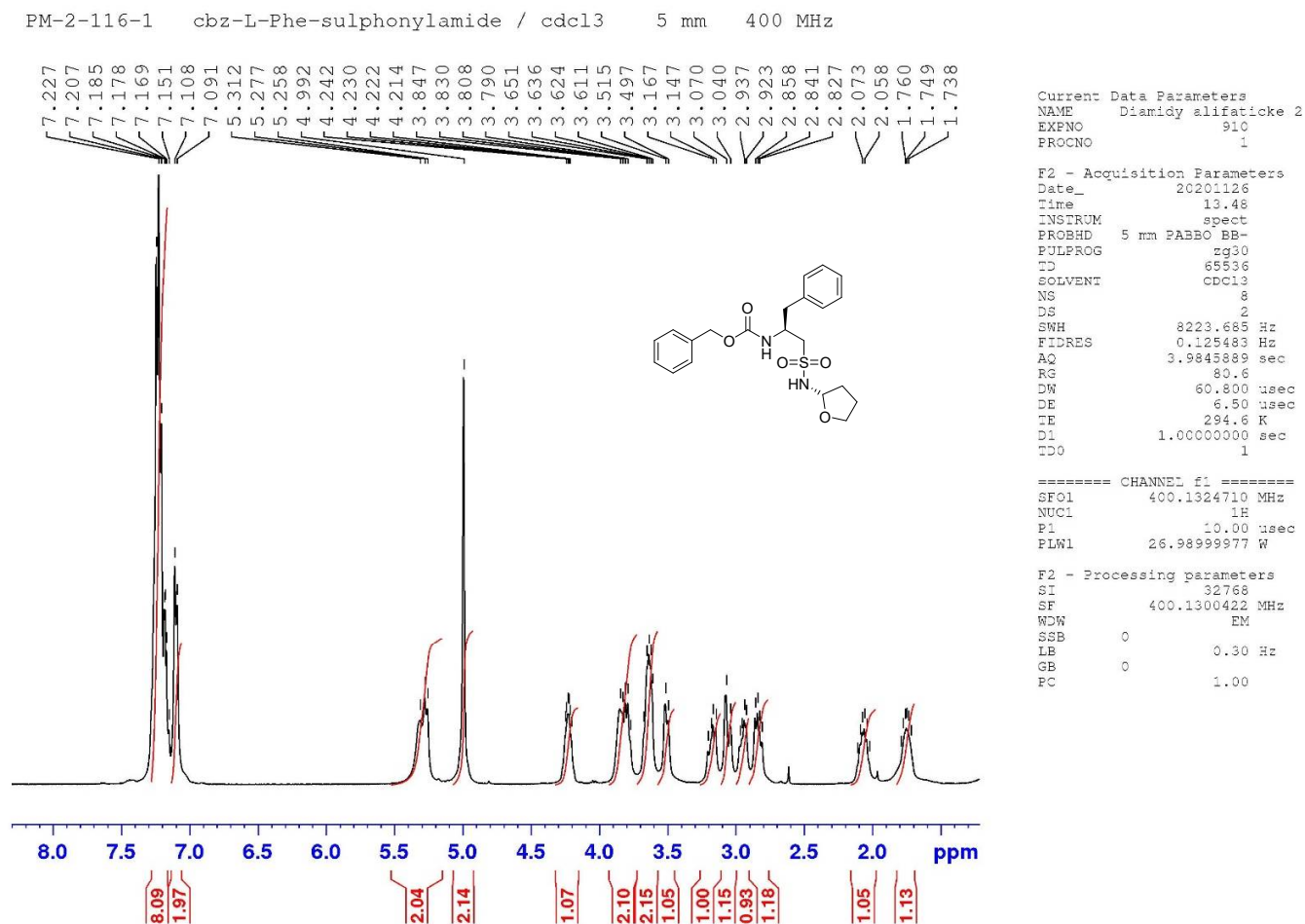
1.33. Copy of ^1H NMR of **5h**



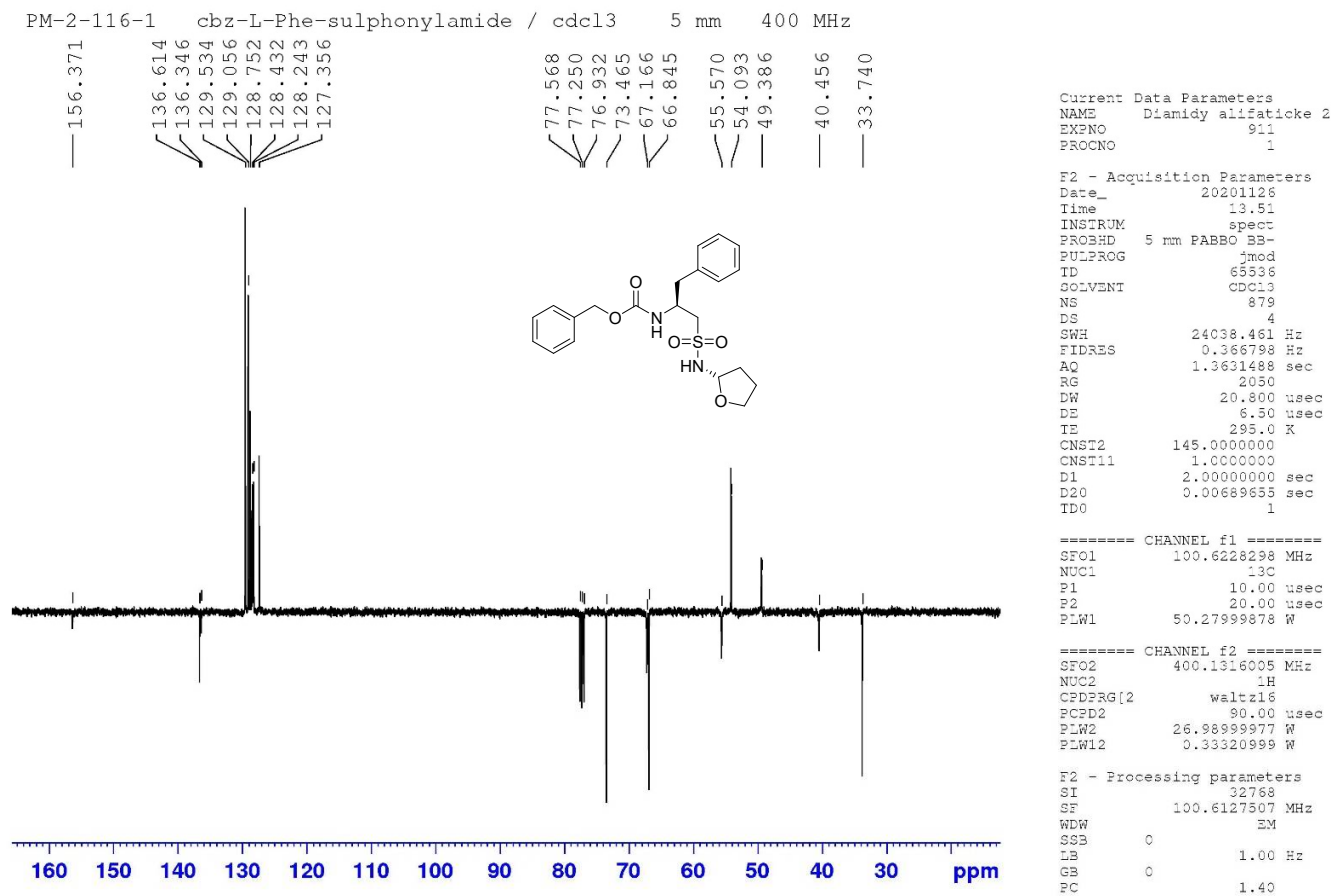
1.34. Copy of ^{13}C NMR of **5h**



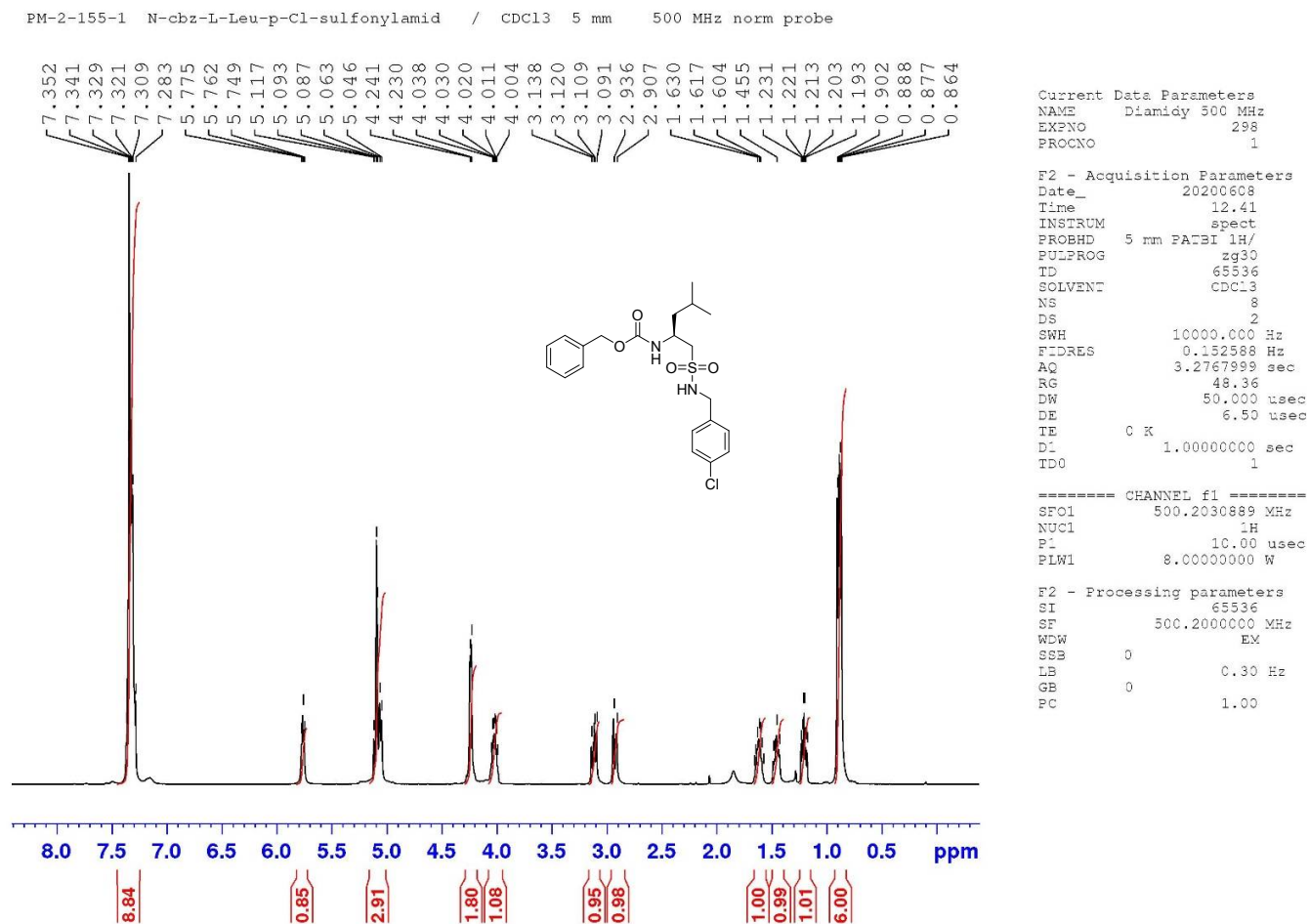
1.35. Copy of ^1H NMR of **5i**



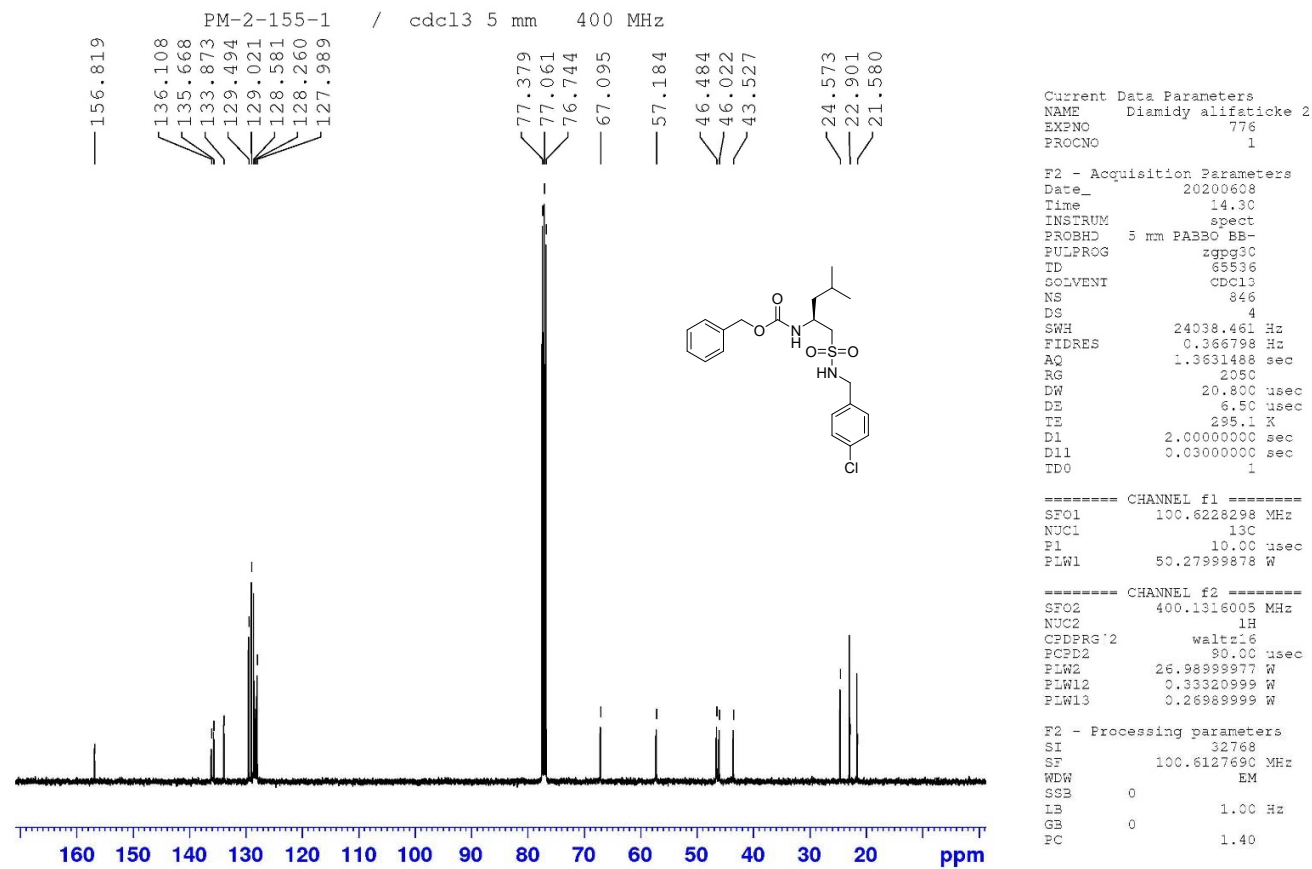
1.36. Copy of ^{13}C NMR of **5i**



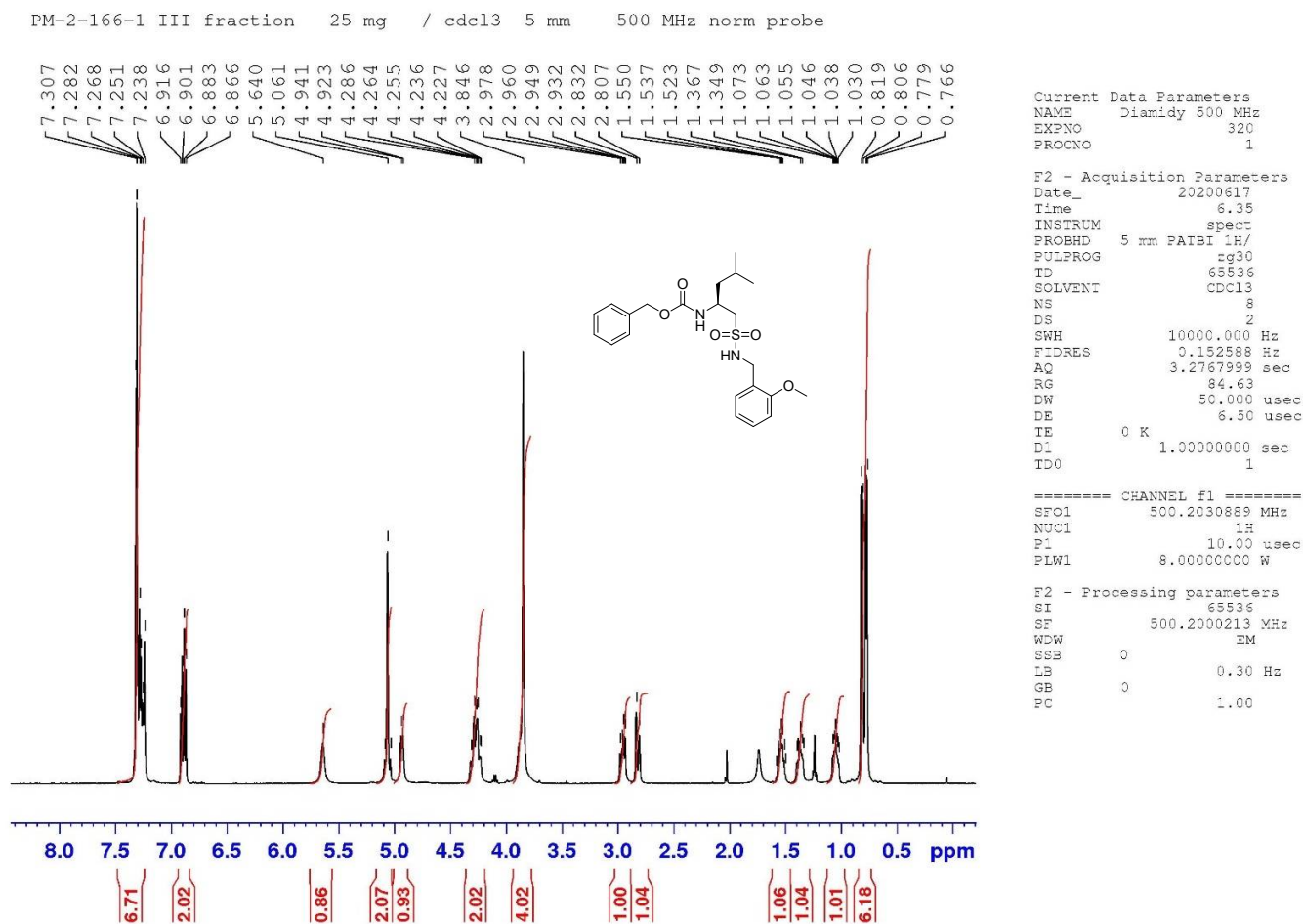
1.37. Copy of ^1H NMR of **5j**



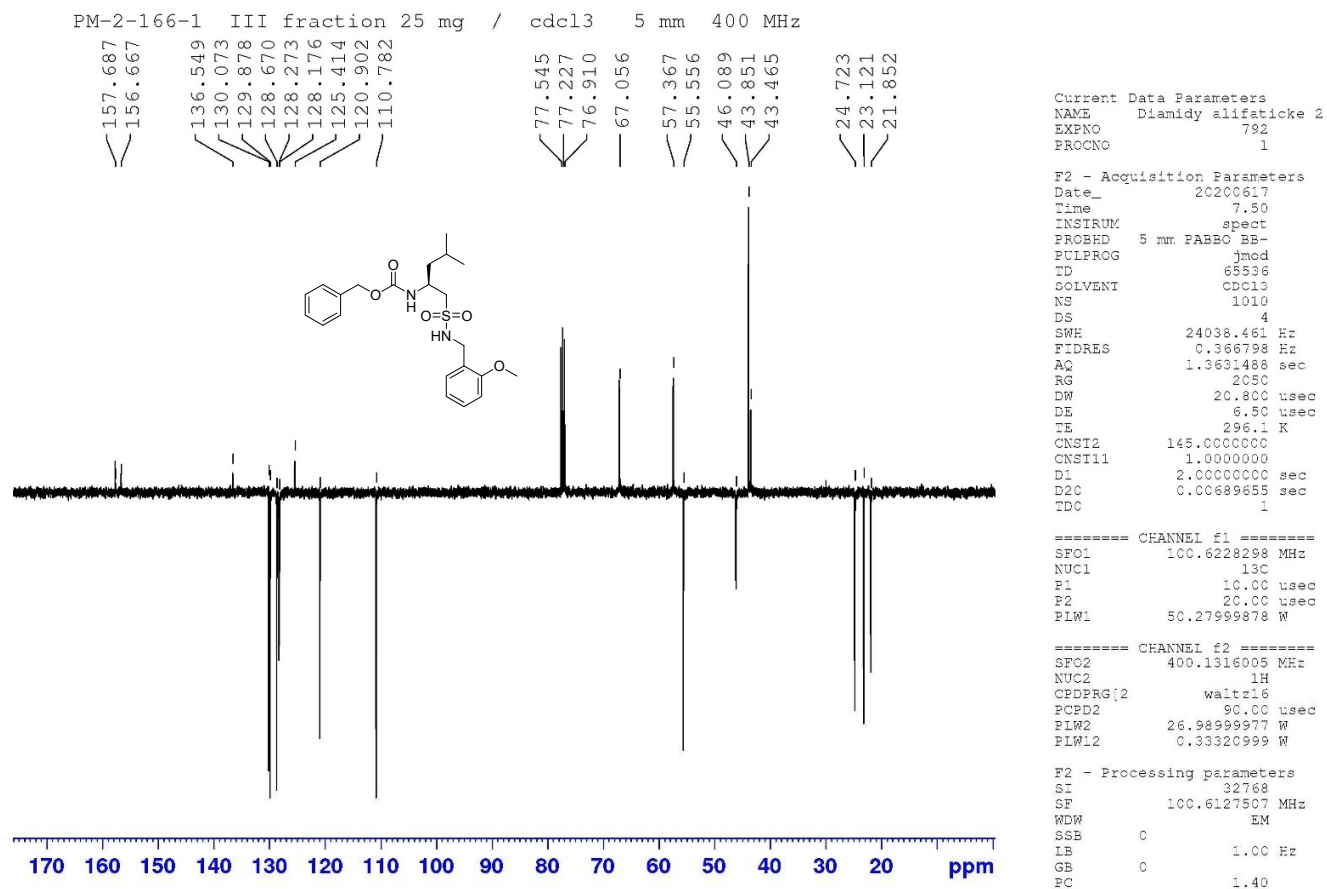
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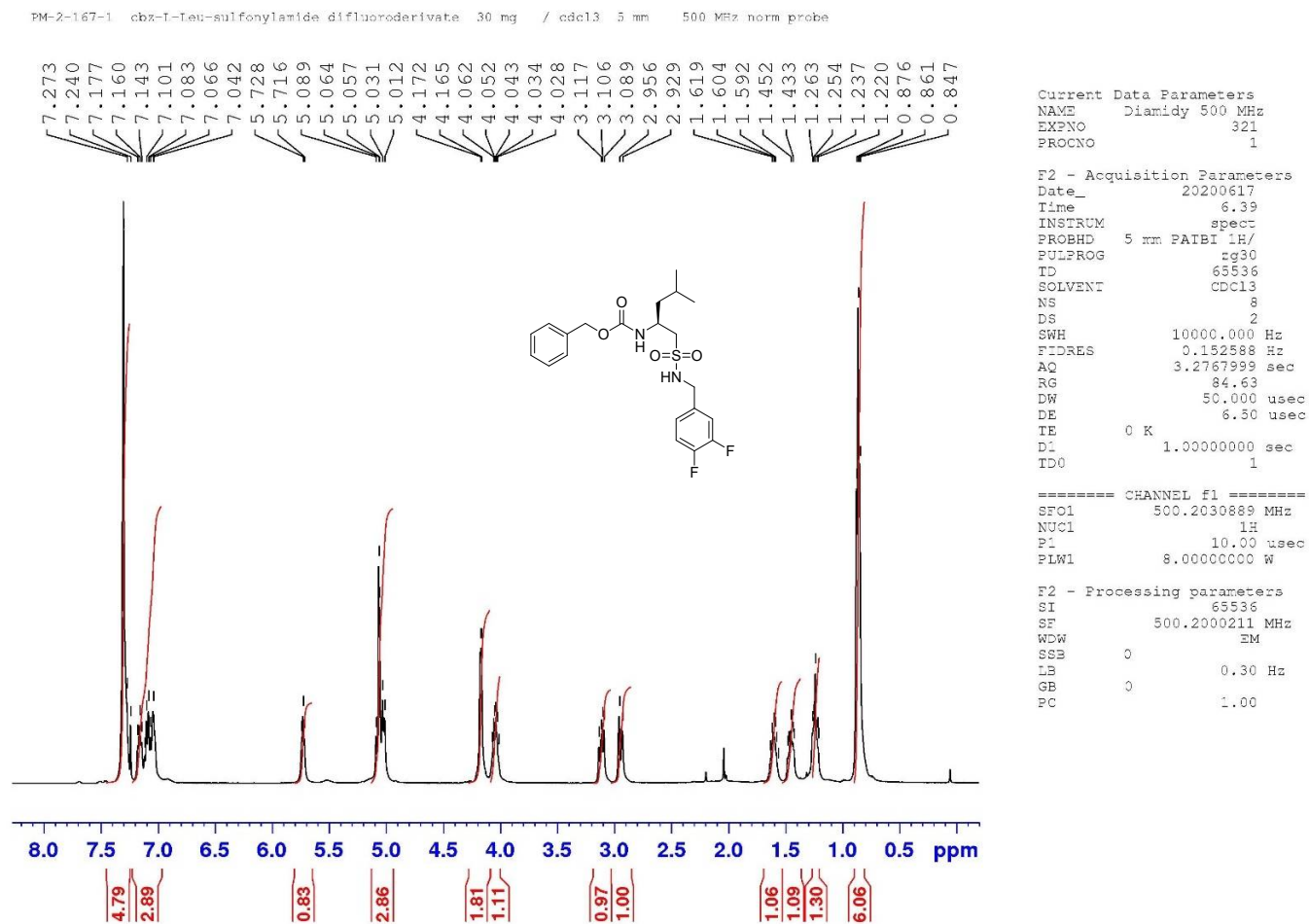
1.39. Copy of ^1H NMR of **5k**



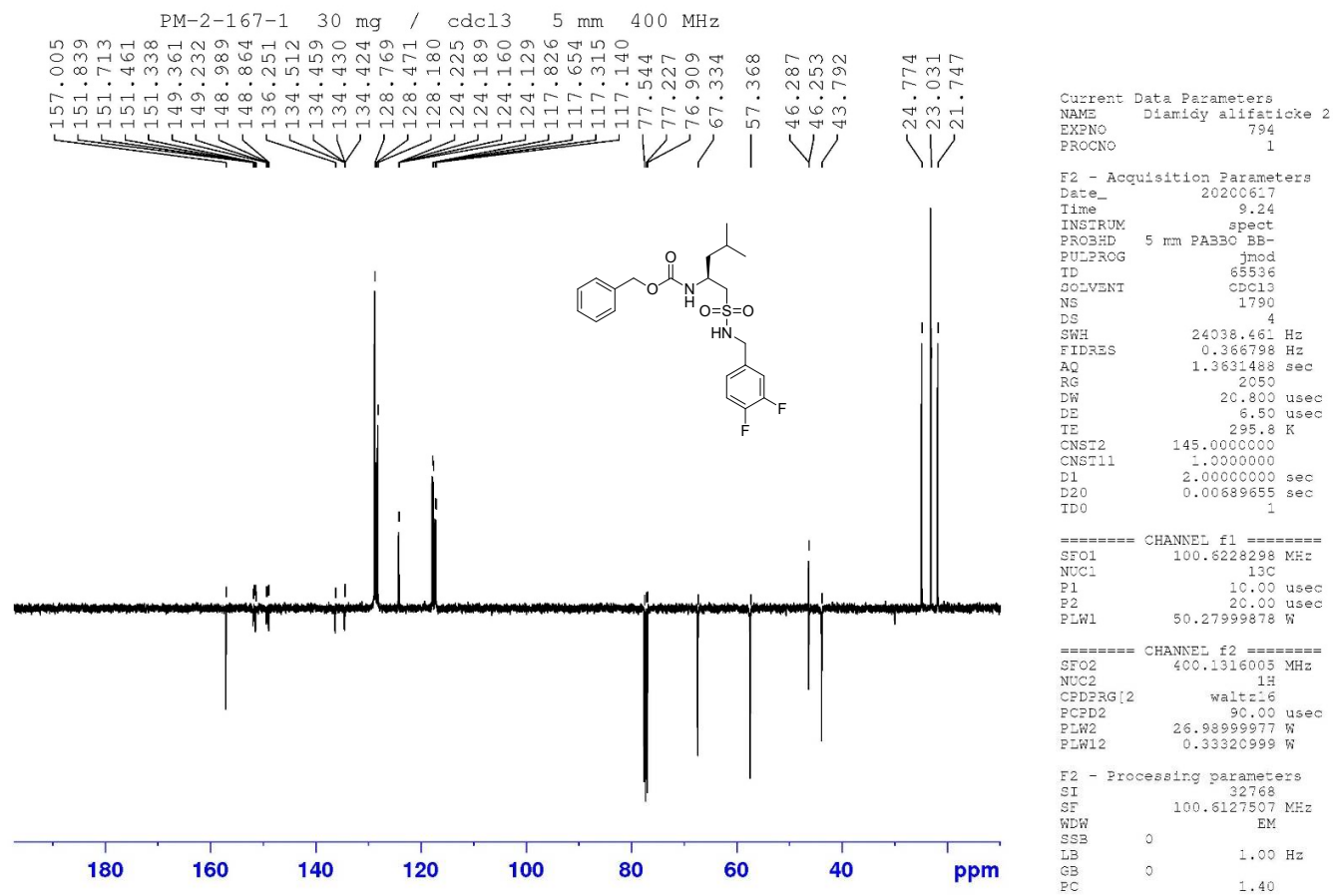
1.40. Copy of ^{13}C NMR of **5k**



1.41. Copy of ^1H NMR of **5l**

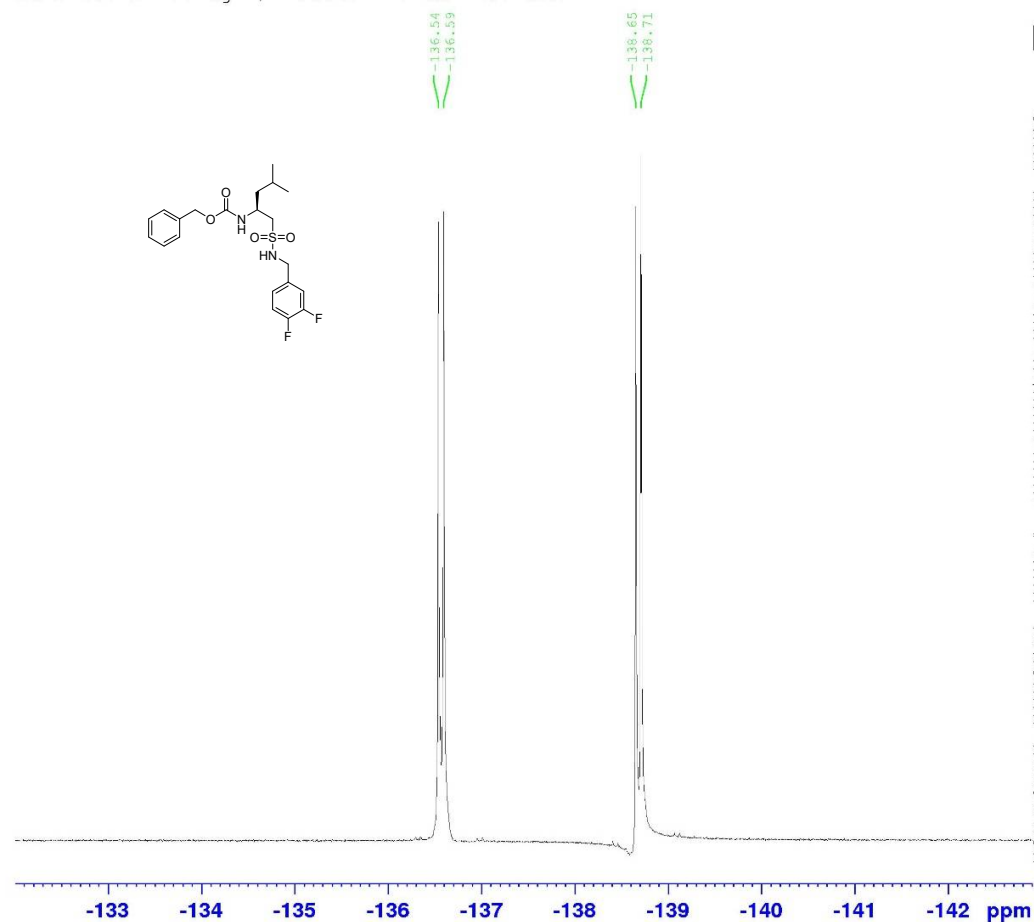
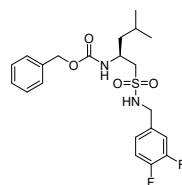


1.42. Copy of ^{13}C NMR of **5l**



1.43. Copy of ^{19}F NMR of **5I**

PM-2-167-1 30 mg / cdcl_3 5 mm 400 MHz



Current Data Parameters
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EXPNO 793
PROCNO 1

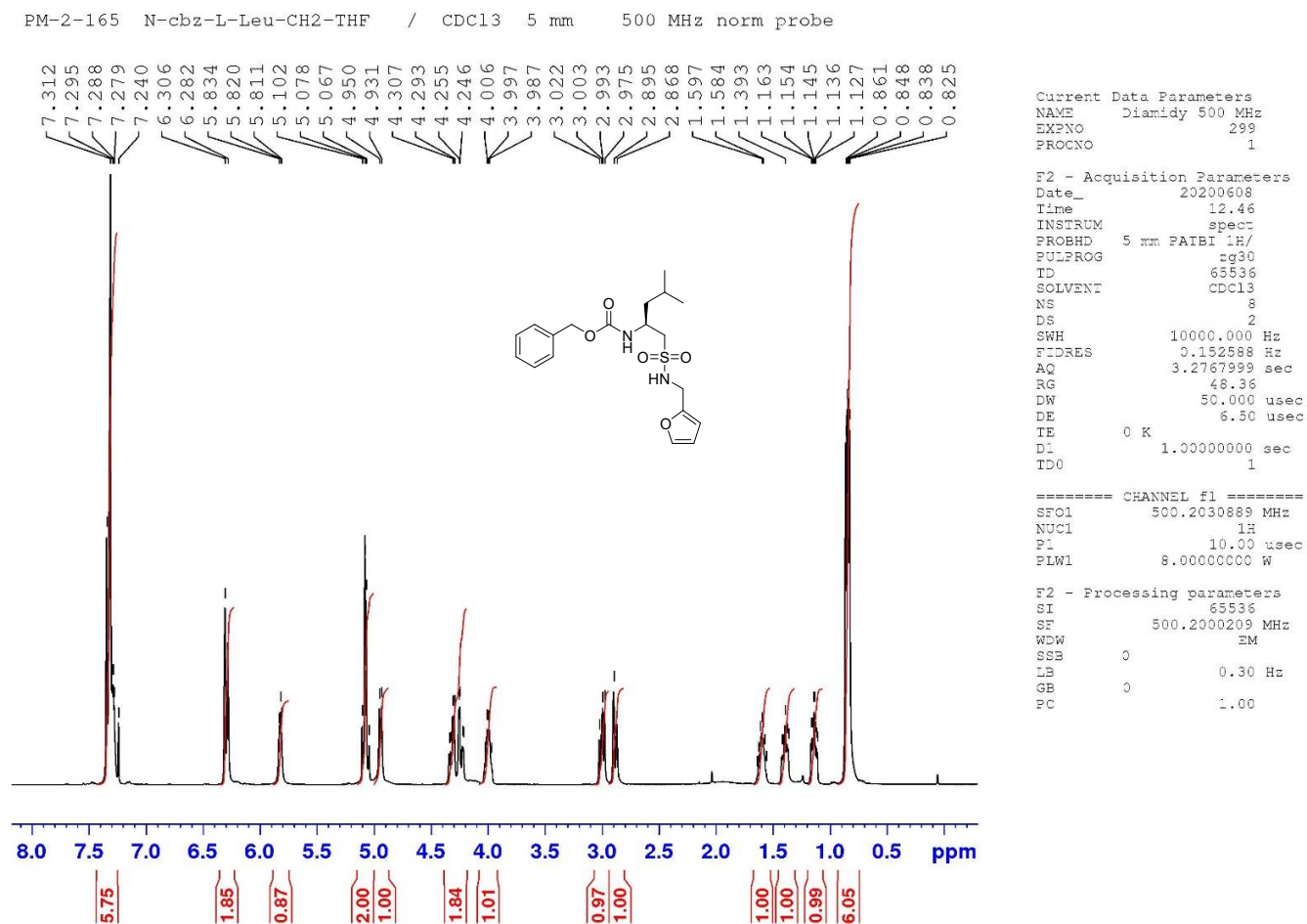
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RG 2050
DW 5.600 usec
DE 6.50 usec
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D11 0.03000000 sec
D12 0.00002000 sec
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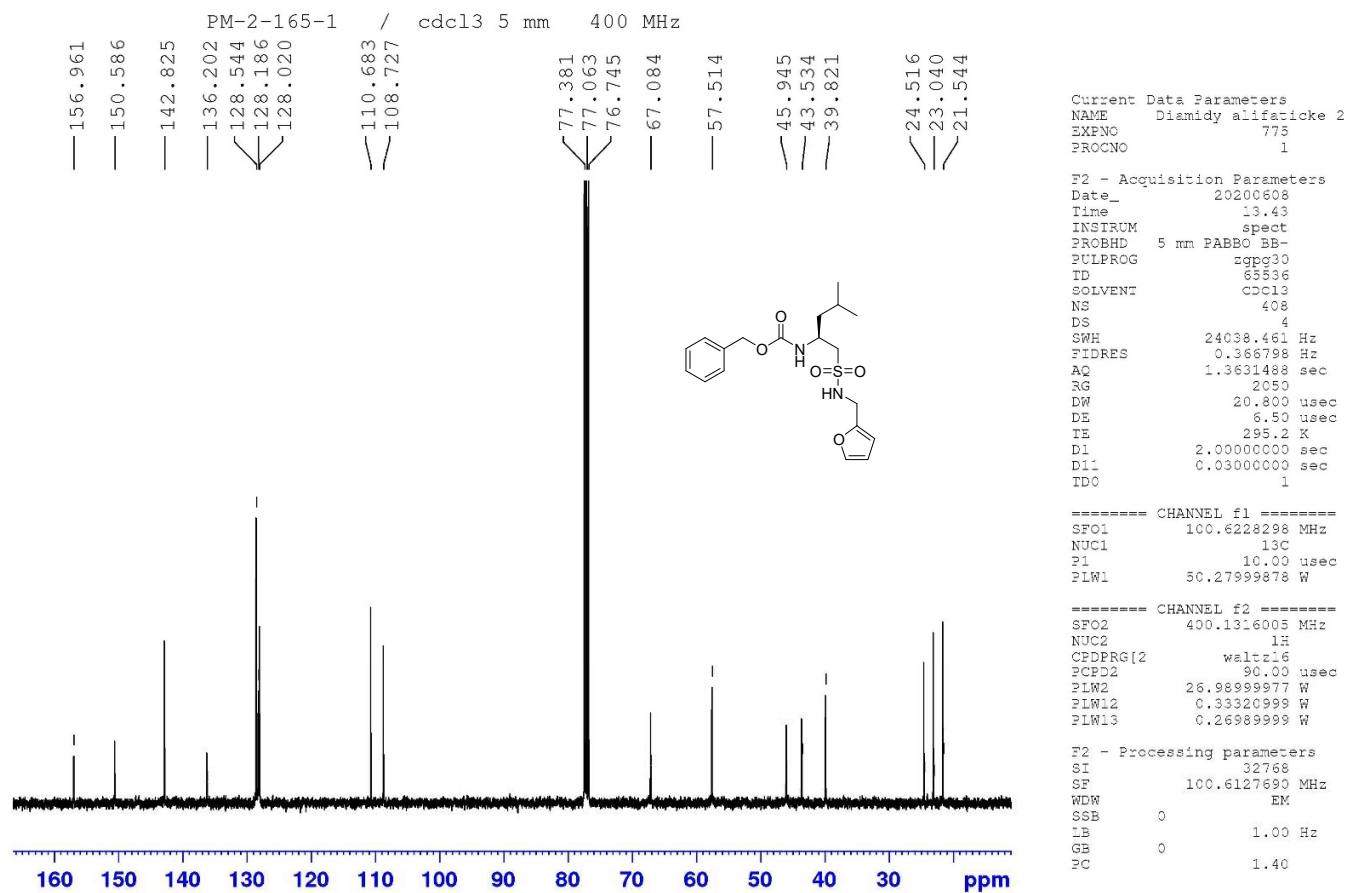
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F2 - Processing parameters
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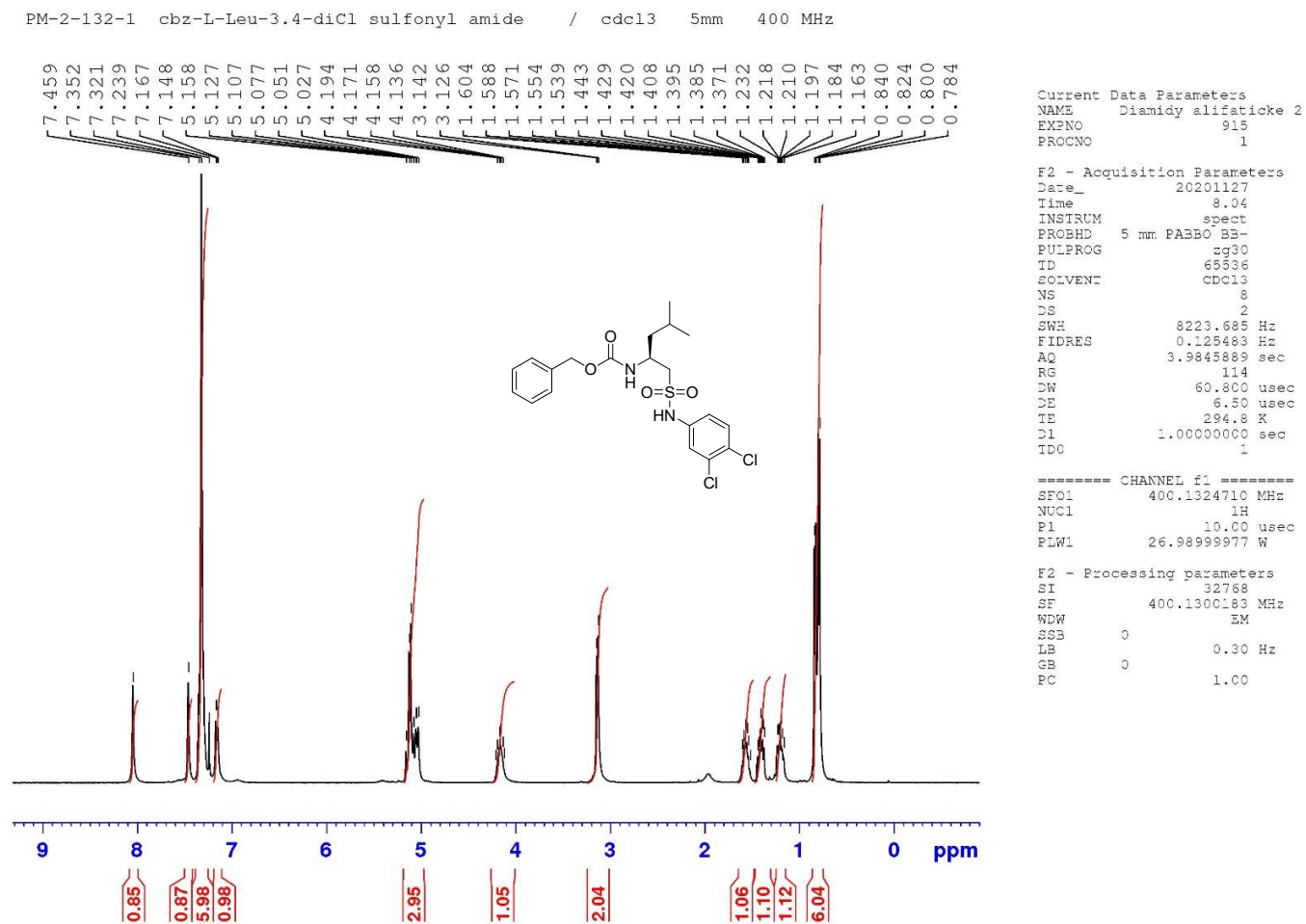
1.44. Copy of ^1H NMR of **5m**



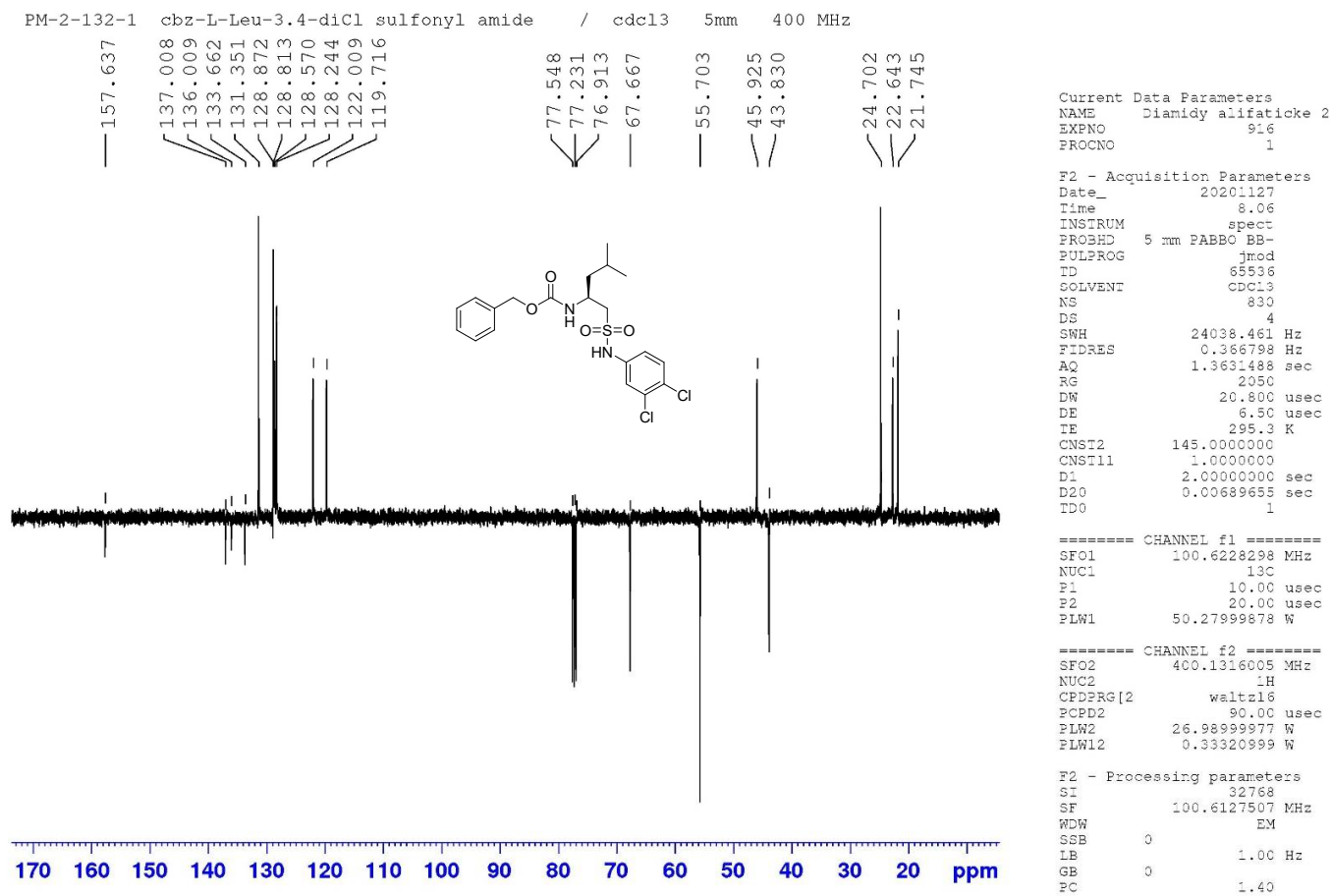
1.45. Copy of ^{13}C NMR of **5m**



1.46. Copy of ^1H NMR of **5n**



1.47. Copy of ^{13}C NMR of **5n**



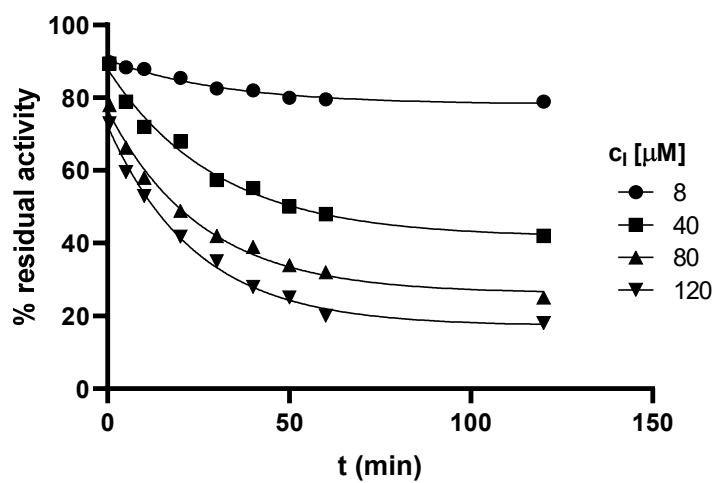


Figure S2 The dependence % residual BChE activity vs. time in the presence of **5k**. The points are expressed as the mean of at least two separate experiments.

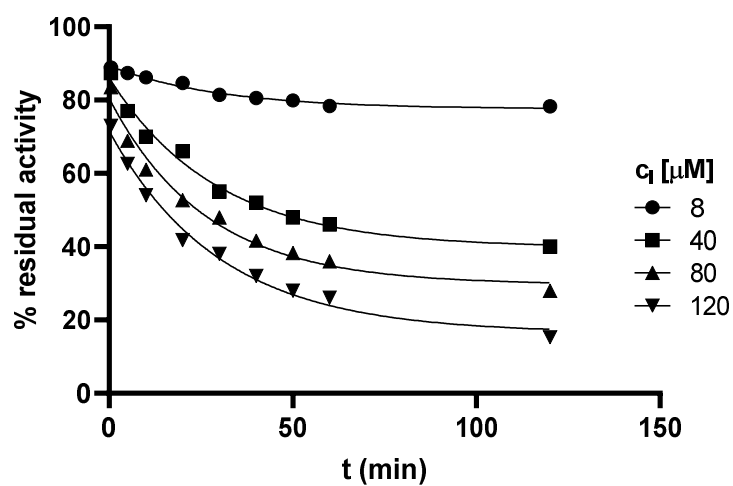


Figure S3 The dependence % residual BChE activity vs. time in the presence of **5j**. The points are expressed as the mean of at least two separate experiments.

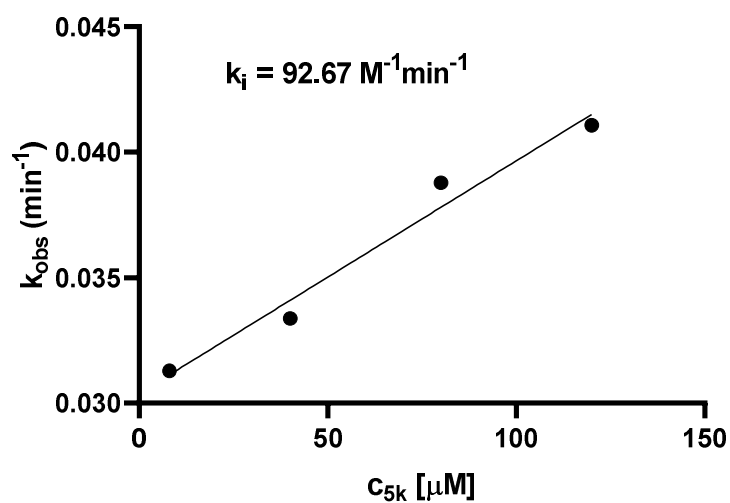


Figure S4 The dependence pseudo-first-order inhibition rates vs. concentration of **5k**.

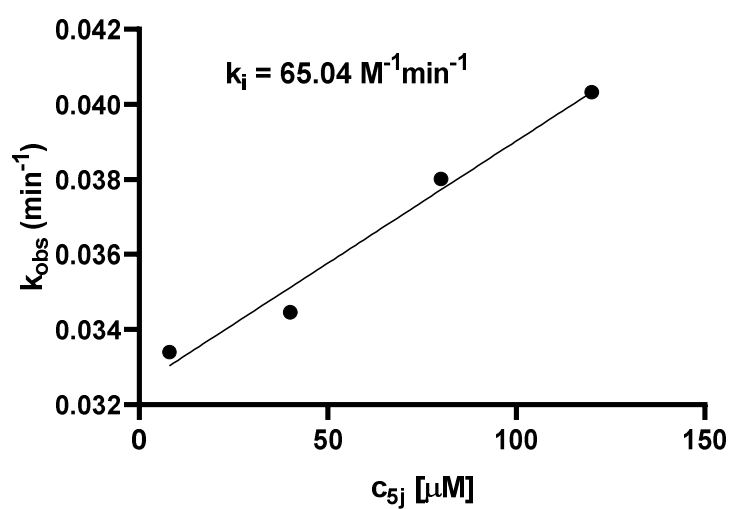


Figure S5 The dependence pseudo-first-order inhibition rates vs. concentration of **5j**.

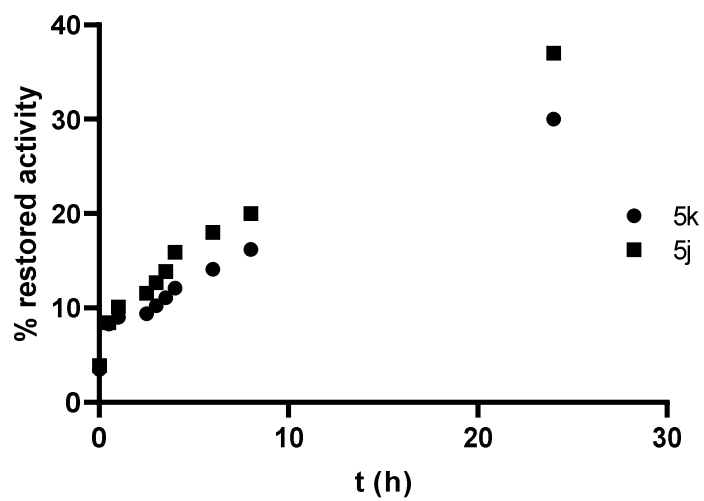


Figure S6 Reactivation of BChE inhibited by **5k** and **5j**. The points are expressed as the mean of at least two separate experiments.

Preliminary stability test of compound **5k** (HPLC results)

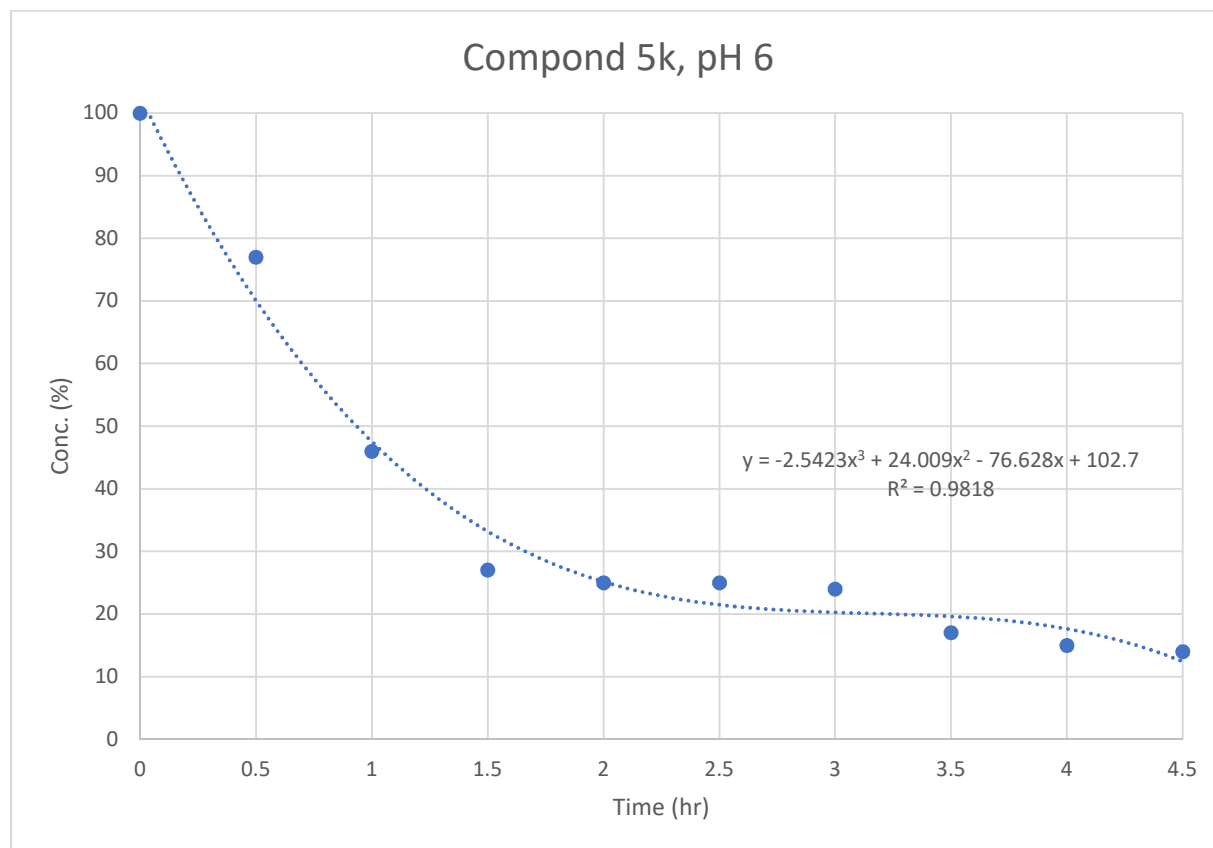


Fig. 1. Decomposition study of compound **5k** at pH 6

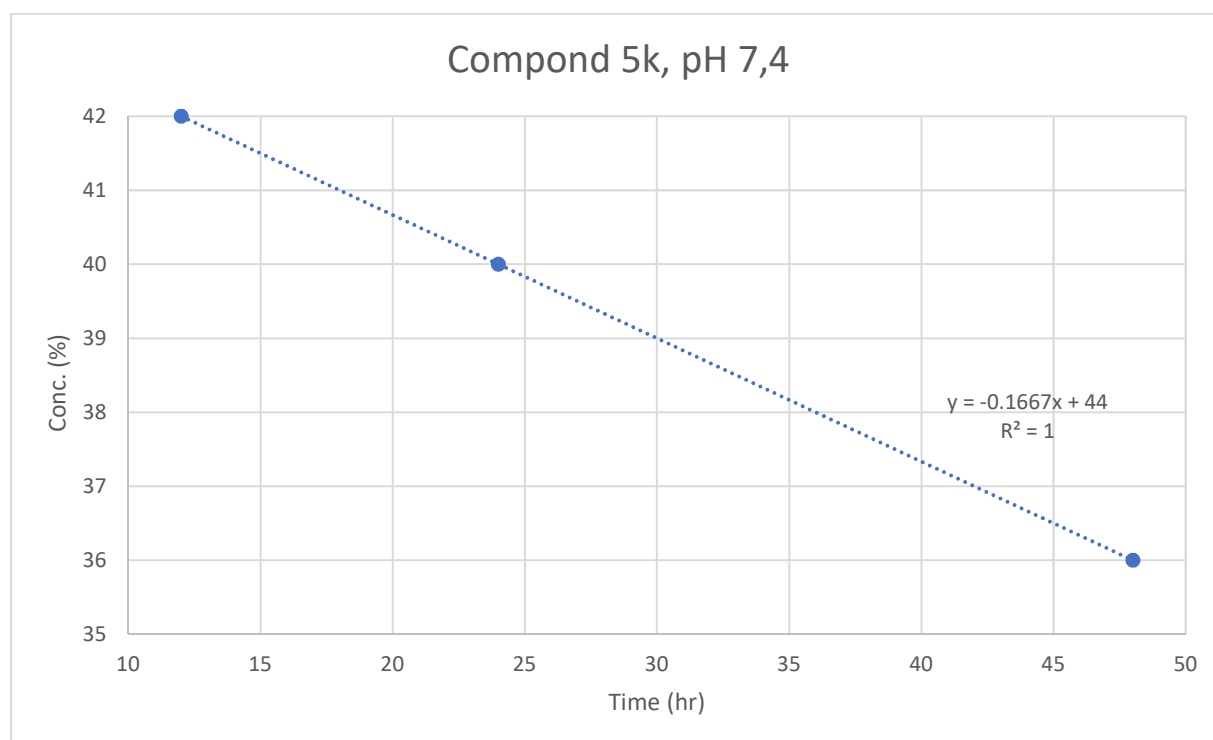
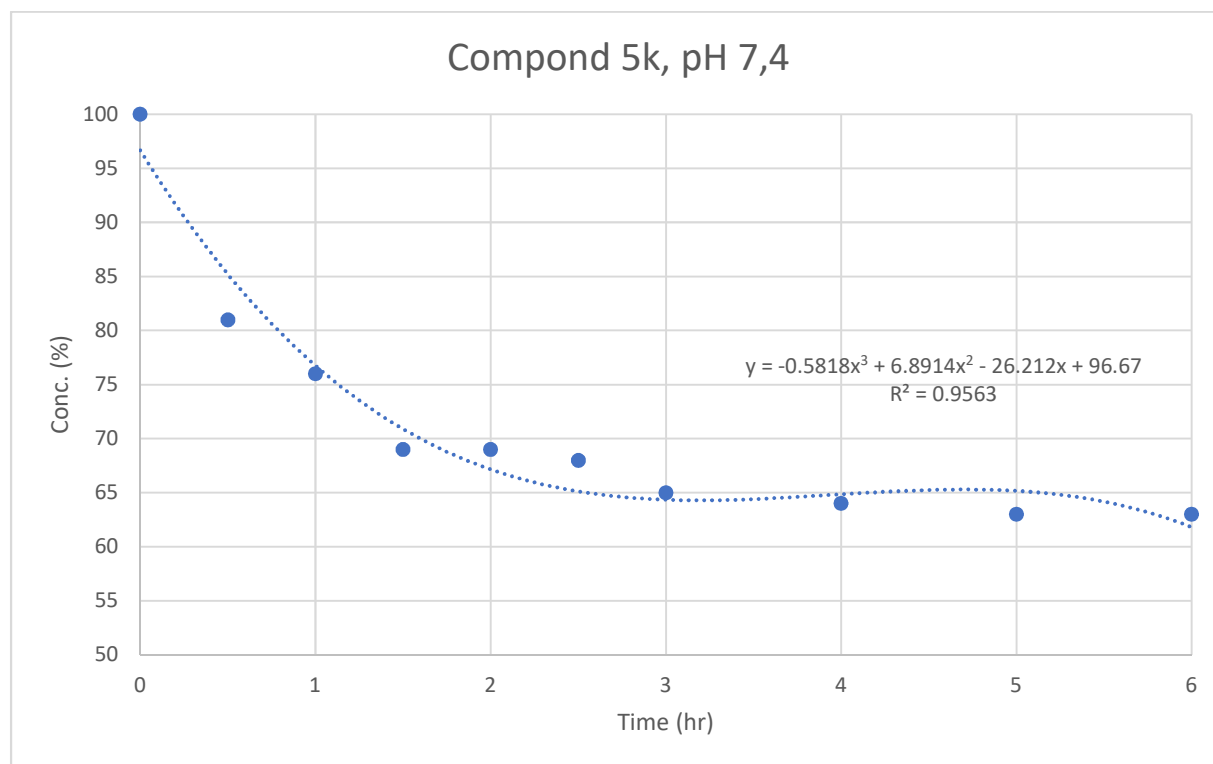


Fig. 2. Decomposition study of compound **5k** at pH 7,4.

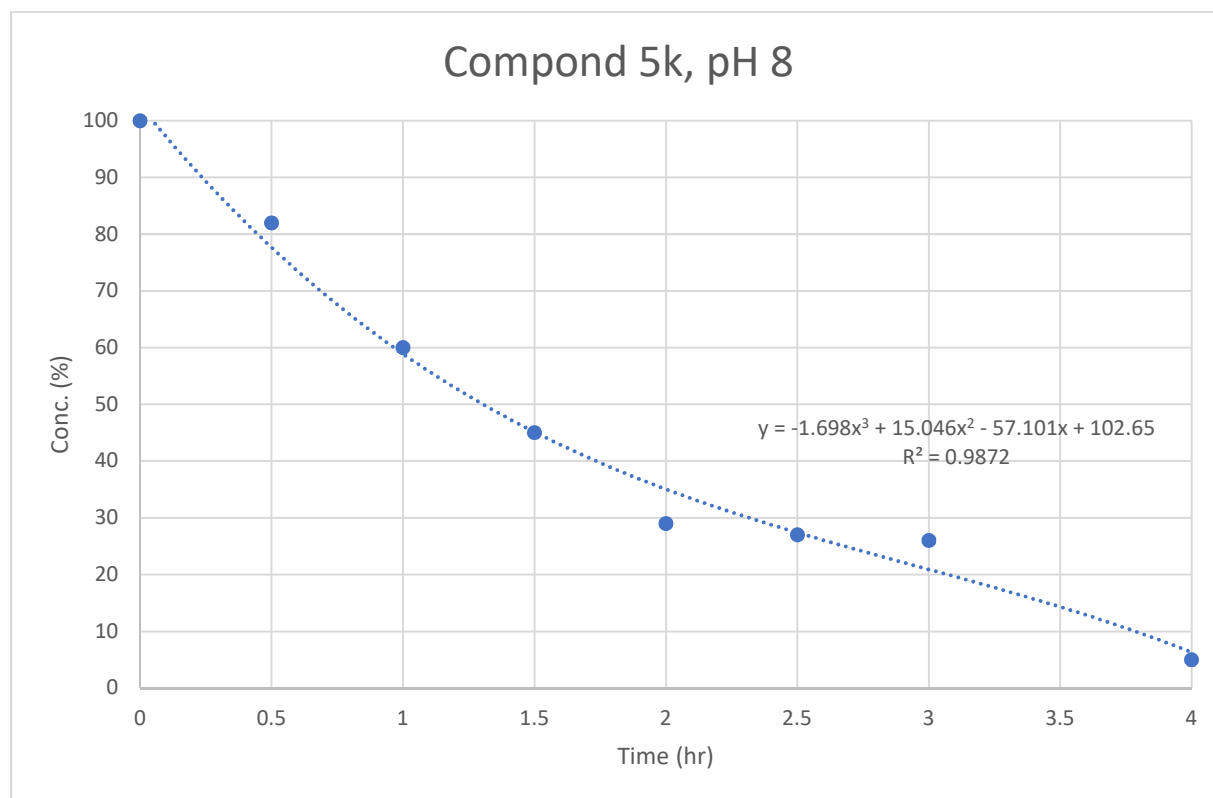


Fig. 3. Decomposition study of compound **5k** at pH 8.

The compound **5k** was chosen for the preliminary study of degradation in three different pH (6, 7.4, 8). Measurement of these degradations was performed by HPLC in reverse phase acetonitrile/water (98:2). The compound was stirred in a tempered buffer (37 °C) of appropriate pH and a sample was taken for HPLC analysis at a specified time. The individual graphs show a clear decomposition of the studied substance as a function of time. In all cases, it is always a pseudo-first order, where one component was in excess and its concentration does not change in time.

Figure 1 represents decomposition study of compound **5k** at pH 6. The graph shows a significant decomposition in the first hour, when 75% of the substance is rapidly decomposed. However after 6 hrs remains only 13% of starting material in a sample.

In second case (**Figure 1**) If the **5k** substance was stirred in a pH 7.4 buffer, its decomposition took twice as long as in the case of pH 6 as it can be seen from the graph. Then there is a gradual decomposition, when only 36% of the starting material remains in 48 hours.

A more interesting situation occurs when the compound **5k** is stirred in a buffer of pH 8. Very rapid decomposition occurs as early as 4 hours, when only 5% of the starting material is present in the mixture.

As already mentioned, this is a preliminary study, which was repeated only twice. This is statistically low and it will require further follow-up experiments.

Procedure:

The substance (5k) was dissolved in a buffer of desired pH in Erlenmayer flask and it was placed to tempered water bath (37 °C). At zero time, a sample is taken immediately from solution (1 ml) and analyzed by HPLC (column C18 for reverse phase, mobil phase AcCN/Water 98/2, column temp. 40 °C, flow rate 1 ml/min, injection 10 µl, detection at 214 and 254 nm). Another samples which were taken from solution were at times: 0,5h, 1h, 1,5h, 2h, 2,5h, 3h, 4h, 5h, 6h, 12h, 24h, 48h, 72h and 96h.