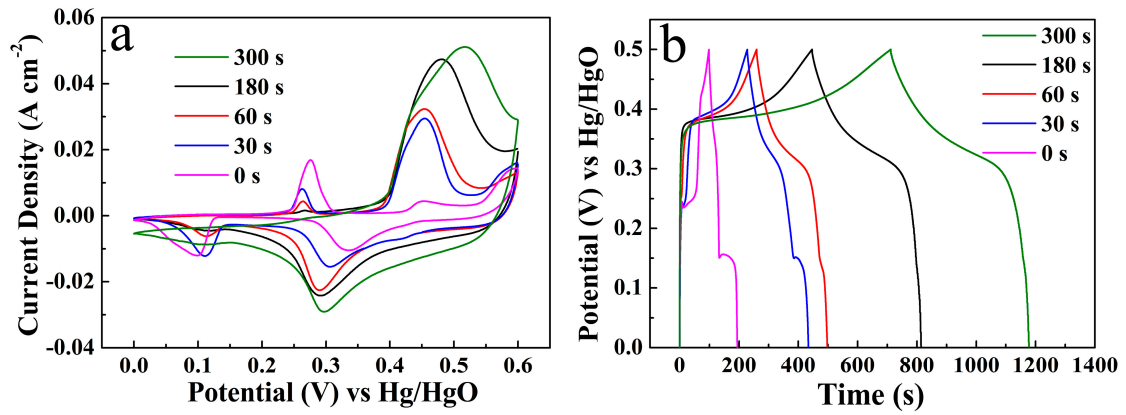


# Hierarchical Core/Shell Structured Ag@Ni(OH)<sub>2</sub> Nanospheres as Binder-Free Electrodes for High Performance Supercapacitors

Sa Lv \*, Xuefeng Chu, Fan Yang, Huan Wang, Jia Yang, Yaodan Chi and Xiaotian Yang \*

Jilin Provincial Key Laboratory of Architectural Electricity & Comprehensive Energy Saving, School of Electrical Engineering and Computer, Jilin Jianzhu University, Changchun 130118, China;  
stone2009@126.com; ctpnxxn@163.com; whuan@ciac.ac.cn; yangjia@jlju.edu.cn; chiyaodan@jlju.edu.cn

\* Correspondence: lvsa82@163.com (S.L.); hanyxt@163.com (X.Y.); Tel.: +86-0431-8456-6181 (S.L.)



**Figure S1.** Ag@Ni(OH)<sub>2</sub> electrode corresponding to different Ni(OH)<sub>2</sub> deposition time: (a) CV curves; (b) GCD curves.

The specific capacitance of Ag@Ni(OH)<sub>2</sub> electrode can be obtained by the following equation:

$$C_s = \frac{\int I(V)dV}{2vS\Delta V}$$

where  $C_s$  (F cm<sup>-2</sup>) is the specific capacitance,  $I(V)$  (A) is the response current,  $V$  (V) is the potential vs. Hg/HgO,  $v$  (V s<sup>-1</sup>) is the scan rate,  $S$  (cm<sup>2</sup>) is the effective area of the electrode and  $\Delta V$  (V) is the working potential. So the specific capacitance decreases with the increase of scan rate.

The specific capacitance of Ag@Ni(OH)<sub>2</sub> electrode can be obtained by the following equation:

$$C_s = \frac{I \times \Delta t}{S\Delta V}$$

where  $C_s$  (F cm<sup>-2</sup>) is the specific capacitance,  $I$  (A) is the charge-discharge current,  $\Delta t$  (s) is the discharging time,  $S$  (cm<sup>2</sup>) is the effective area of the electrode and  $\Delta V$  (V) represents the potential drop during discharge.

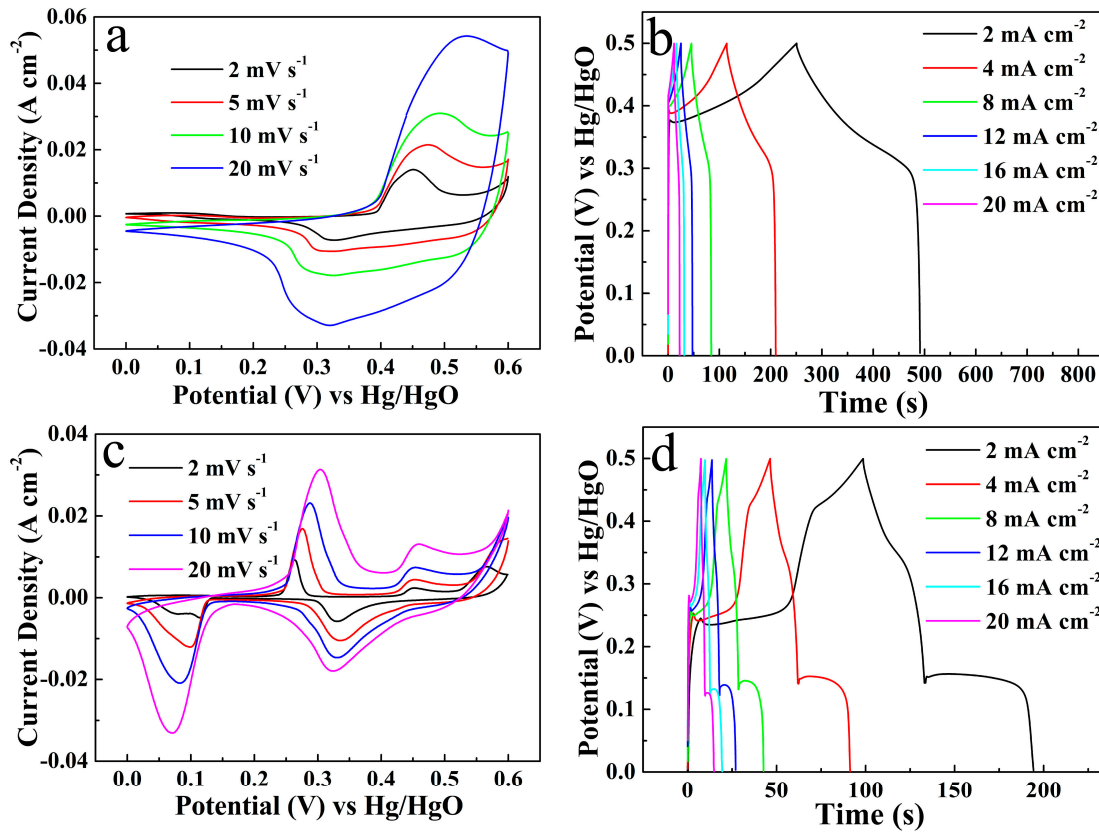


Figure S2. (a,c) The CV curves and (b,d) GCD curves of Ni(OH)<sub>2</sub> and the Ag electrode, respectively.

Table S1. The specific capacitance versus current density of the three electrodes.

Electrode $C_s$ (F cm <sup>-2</sup> )	2 mA	4 mA	8 mA	12 mA	16 mA	20 mA
Ag@Ni(OH) <sub>2</sub>	1.864	1.395	1.098	0.946	0.838	0.760
Ni(OH) <sub>2</sub>	0.962	0.766	0.619	0.540	0.480	0.436
Ag	0.388	0.359	0.336	0.319	0.310	0.296

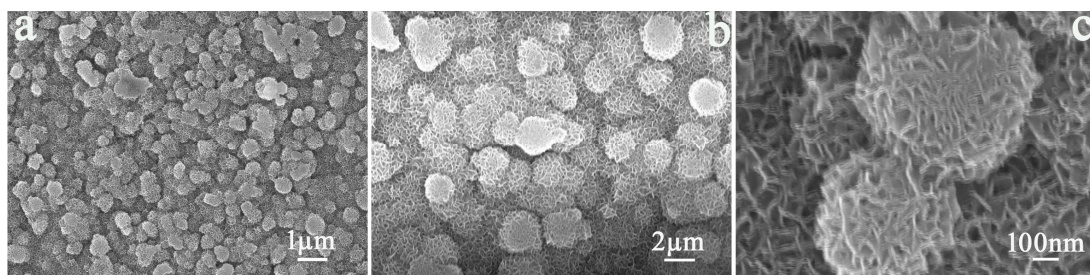
The average equivalent series resistance ( $R_{ESR}$ ) are derived from the equation:

$$R_{ESR} = \frac{V_{drop}}{2I}$$

where,  $V_{drop}$  (V) is the abrupt voltage drop at the beginning of the discharging curve and  $I$  (A) is the corresponding current.

Table S2. The  $R_{ESR}$  and voltage drop versus current density of the Ag@Ni(OH)<sub>2</sub> electrode.

$I$ (mA)	2 mA	4 mA	8 mA	12 mA	16 mA	20 mA
$V_{drop}$ (V)	0.0043	0.0081	0.0156	0.0230	0.0305	0.0380
$R_{ESR}$ ( $\Omega$ cm <sup>-2</sup> )	1.075	1.015	0.975	0.958	0.953	0.950
Average $R_{ESR}$ ( $\Omega$ cm <sup>-2</sup> )	0.988					



**Figure S3.** FE-SEM images of the Ag@Ni(OH)<sub>2</sub> composite at different magnifications after 3000 cycles.