## **Supplementary information**

## Anti-proliferative and anti-migration activity of arene-ruthenium(II) complexes with azole therapeutic agents

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Table of contents

Fig S1. Spectrofluorometric titration spectra of HSA with the ruthenium compounds **2-4**.

Fig S2. <sup>1</sup>H NMR spectrum of guanosine and complex **3** at different times (only resonances of  $H_8$ , NH and NH<sub>2</sub> are assignments).

Fig S3. Hoechst 33258 (H33258) displacement assay for metal complexes 2-4.

Fig S4.  ${}^{31}P{}^{1}H{}$  NMR spectrum in the mixture 70:30 DMSO: Culture medium (DMEM) of complex **3** at different times.

Equation description:

Classical Stern–Volmer equation (1) was used to determine the Stern–Volmer quenching constant (Ksv):

$$F_0/F = 1 + Kq\tau_0[Q] = 1 + Ksv[Q]$$
 (1)

where  $F_0$  and F are the fluorescence intensities in the absence and presence of quencher, respectively, [Q] is the quencher concentration.

Binding constant (Kb) and number of binding sites (n) were determined by plotting the double log graph of the fluorescence data using the equation (2):

 $\log [(F_0-F)/F] = \log Kb + n\log[Q]$  (2)

The thermodynamic parameters were calculated from equations (3) and (4):

$$\ln (K_2/K_1) = [(1/T_1) - (1/T_2)]\Delta H/R (3)$$

where  $K_1$  and  $K_2$  are the binding constants at temperatures  $T_1$  and  $T_2$ , respectively, and enthalpy ( $\Delta H$ ) and R is the gas constant.

Additionally, the change in free energy ( $\Delta G$ ) and entropy ( $\Delta S$ ) were calculated from the following equation:

$$\Delta G = -RT \ln K = \Delta H - T\Delta S (4)$$

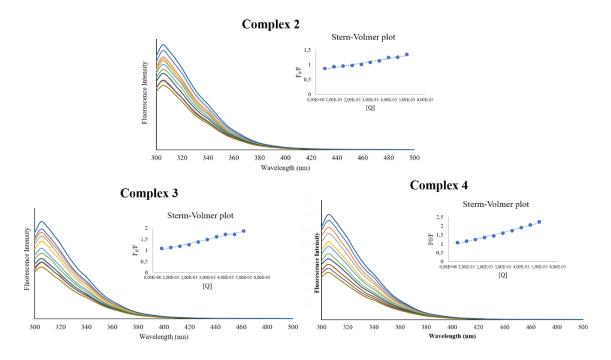


Fig S1. Spectrofluorometric titration spectra of HSA with the ruthenium compounds 2-4

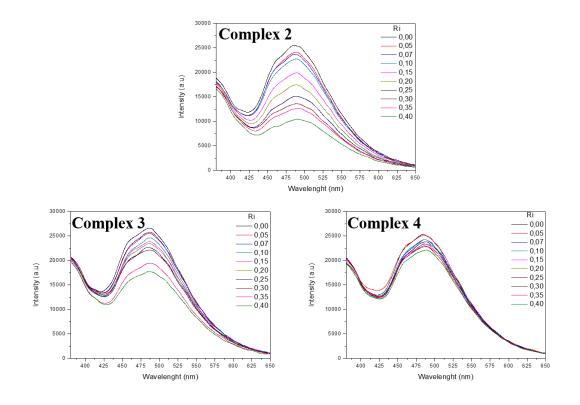


Fig S2. Hoechst 33258 (H33258) displacement assay for metal complexes 2-4

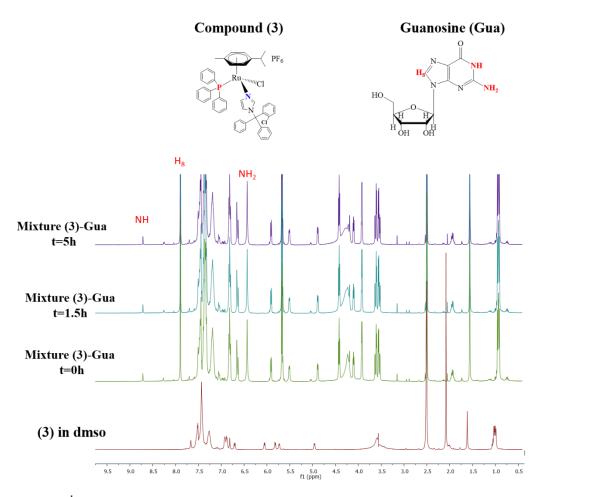


Fig S3. <sup>1</sup>H NMR spectrum of guanosine and complex **3** at different times (only resonances of  $H_8$ , NH and NH<sub>2</sub> are assignments)

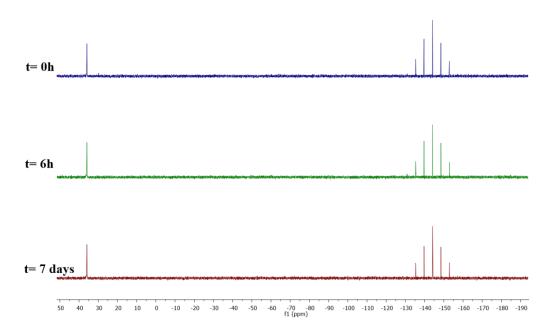


Fig S4.  ${}^{31}P{}^{1}H$  NMR spectrum in the mixture 70:30 DMSO:Culture medium (DMEM) of complex **3** at different times