

Article

A Quasi-Intramolecular Solid-Phase Redox Reaction of Ammonia Ligands and Perchlorate Anion in Diamminesilver(I) Perchlorate

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Citation: Fogaça, L.A.; Bereczki, L.; Petruševski, V.M.; Barta-Holló, B.; Franguelli, F.P.; Mohai, M.; Béres, K.A.; Sajó, I.E.; Szilágyi, I.M.; Kótai, L.; A Quasi-Intramolecular Solid-Phase Redox Reaction of Ammonia Ligands and Perchlorate Anion in Diamminesilver(I) Perchlorate. *Inorganics* **2021**, *9*, 38. <https://doi.org/10.3390/inorganics9050038>

Academic Editor: Kazuyuki Takahashi

Received: 3 April 2021
Accepted: 7 May 2021
Published: 9 May 2021

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Abstract: The reaction of ammoniacal AgNO₃ solution (or aq. solution of [Ag(NH₃)₂]NO₃) with aq. NaClO₄ resulted in [Ag(NH₃)₂]ClO₄ (compound **1**). Detailed spectroscopic (correlation analysis, IR, Raman, and UV) analyses were performed on [Ag(NH₃)₂]ClO₄. The temperature and enthalpy of phase change for compound **1** were determined to be 225.7 K and 103.04 kJ/mol, respectively. We found the thermal decomposition of [Ag(NH₃)₂]ClO₄ involves a solid-phase quasi-intramolecular redox reaction between the perchlorate anion and ammonia ligand, resulting in lower valence chlorine oxyacid (chlorite, chlorate) components. We did not detect thermal ammonia loss during the formation of AgClO₄. However, a redox reaction between the ammonia and perchlorate ion resulted in intermediates containing chlorate/chlorite, which disproportionated (either in the solid phase or in aqueous solutions after the dissolution of these decomposition intermediates in water) into AgCl and silver perchlorate. We propose that the solid phase AgCl-AgClO₄ mixture eutectically melts, and the resulting AgClO₄ decomposes in this melt into AgCl and O₂. Thus, the final product of decomposition is AgCl, N₂, and H₂O. The intermediate (chlorite, chlorate) phases were identified by IR, XPS, and titrimetric methods.

Keywords: ammine; perchlorate; vibrational spectroscopy; redox reaction; thermal decomposition; disproportionation; chlorite; chlorate

1. Introduction

Continuing our research on thermally initiated solid-phase, quasi-intramolecular redox reactions of compounds containing redox-active cationic and anionic parts [1–4], including ammonia complexes of transition metal with tetraoxometallate anions [5–10], we prepared some diamminesilver(I) complexes (e.g., [Ag(NH₃)₂]XO₄, X = Mn and Cl) and studied them in detail. In this paper, we present our results on the investigation of monoclinic and orthorhombic diamminesilver(I) perchlorate polymorphs (**1-M** and **1-O**).

The hydrated $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4 \cdot \text{H}_2\text{O}$ (**1-H₂O**) complex was described by Scagliari and Marangoni [11], but Bruni and Levi [12] raised issues against its existence. The occurrence of $[\text{Ag}(\text{NH}_3)_3]\text{ClO}_4$ (compound **2**) in the solid state and the quantum-chemical results about $[\text{Ag}(\text{NH}_3)_2(\text{H}_2\text{O})]^+$ ion [13] prompted us to study the existence of hydrated forms of compound **1** $[\text{Ag}(\text{NH}_3)_2(\text{H}_2\text{O})]\text{ClO}_4$ (compound **3**) and $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4 \cdot \text{H}_2\text{O}$ (**1-H₂O**)).

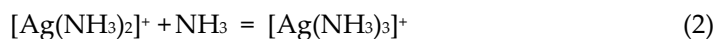
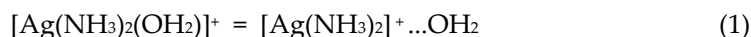
Diamminesilver(I) perchlorate has two known polymorphs [14], a low-temperature monoclinic (**1-M**) and a high-temperature orthorhombic one (**1-O**); however, the enthalpy of phase transition between them has not been studied yet. Furthermore, no spectroscopic data are available about the low-temperature polymorph. Although IR spectroscopic data are available for **1-O**, their peak assignments are questionable. Furthermore, the thermal decomposition of compound **1** (**1-O**) has not been investigated previously. Therefore, in our present work, we supplied the available spectroscopic data about polymorph **1-O** with low-temperature Raman results, compared their measured Raman spectroscopic data, and performed a correlation analysis for both polymorphs. The thermal analysis data, including DSC and TG-MS, were also evaluated to follow the thermal decomposition characteristics of compound **1** (**1-O**). Based on these results, we propose the possibility of solid-phase quasi-intramolecular redox reactions between ammonia ligand and perchlorate anion.

2. Results and Discussion

2.1. Efforts to Synthesize of Compound **1** and Its Hydrate(s)

Diamminesilver(I) perchlorate was first prepared by E. Mitscherlich [15] as a colorless crystalline material, when silver perchlorate was dissolved in ammonia, and the solution was left to evaporate. He did not determine the composition of the product. However, Miles et al. [16], and later Nockeman and Meyer [14], used the same procedure and identified the crystallized compounds as $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$ (compound **1**). Scagliari and Marangoni reacted an ammoniacal silver nitrate solution with aq. sodium perchlorate at +5 °C [11] and analyzed the resulting crystalline mass as $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4 \cdot \text{H}_2\text{O}$ (compound **1-H₂O**). Bruni and Levi repeated this experiment [12] and found that the product of the reaction was anhydrous compound **1**. To solve this controversy, we performed a series of experiments to study the effect of reaction conditions (ammonia concentration, temperature) on the composition of products prepared in the reactions of AgClO_4 and ammonia or $[\text{Ag}(\text{NH}_3)_2]\text{NO}_3$ and NaClO_4 . In all cases, only anhydrous compound **1** was formed. Figure 1 shows the XRD of the reaction products according to the original methods of Mitscherlich and Bruni. The XRD patterns of the products of the two reactions were the same; thus, the formation of compound **1-H₂O** or compound **2** in these reactions can be excluded (Figure 2). However, several researchers managed to prepare compound **2** from AgClO_4 and dry ammonia [12,17].

Quantum chemical calculations showed that there was no stable $[\text{Ag}(\text{NH}_3)_2(\text{OH}_2)]^+$ structure [3 + 0] (three coordinated and zero non-coordinated ligands), and this species always converges into a linear $[\text{Ag}(\text{NH}_3)_2]^+ \dots \text{OH}_2$ hydrogen-bound structure [2 + 1] (two coordinated and one non-coordinated ligands), although a small interaction between the oxygen (in water) lone pair and the silver cation was detected [13]. The $[\text{Ag}(\text{NH}_3)_3]^+$ [3 + 0] cation, however, has stable configurations [13,18].



Comparison of the Gibbs energy changes during the substitution of one of the ammonia molecules with water in the $[\text{Ag}(\text{NH}_3)_3]^+$ cation (+7.5 kJ/mol [13]) (Equation (1)) and the addition of the third ammonia to an $[\text{Ag}(\text{NH}_3)_2]^+$ cation (−30.5 kJ/mol [19])

(Equation (2)) suggest that the formation of hydrated diamminesilver(I) complexes is unfavorable.

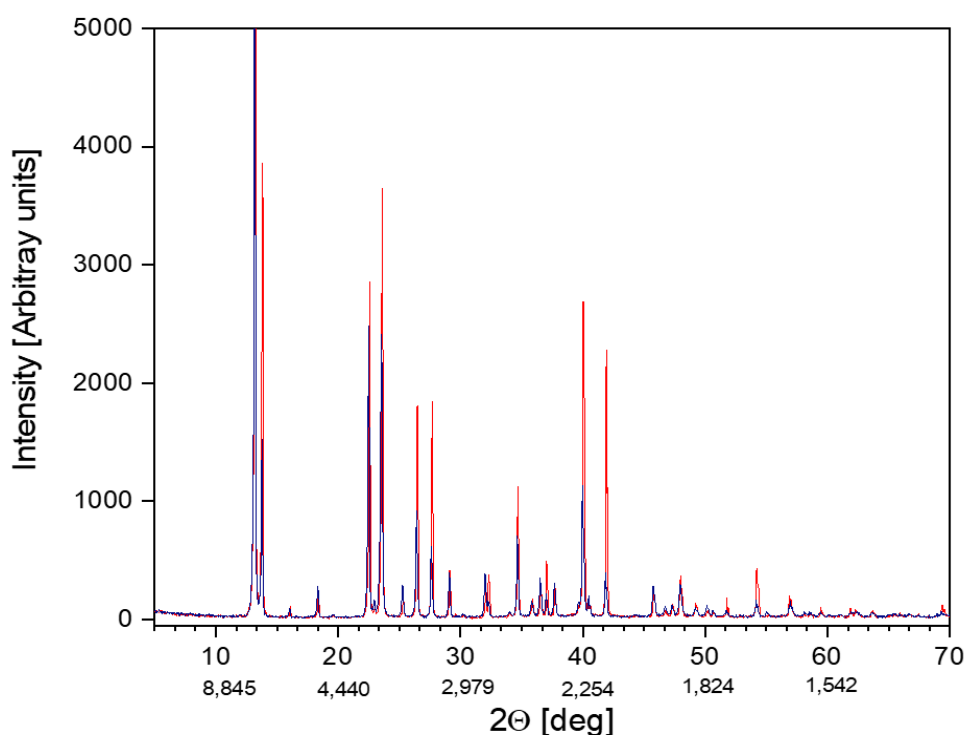


Figure 1. Comparison of the XRD curves of the reaction products prepared with the methods of Mitscherlich [15] (blue) and Scagliari [11] (red).



Figure 2. Schematic representation of phases formed in the $\text{AgClO}_4\text{-NH}_3$ and $\text{AgClO}_4\text{-NH}_3\text{-H}_2\text{O}$ systems.

We managed to isolate $[\text{Ag}(\text{NH}_3)_3]\text{ClO}_4$ in the solid phase. The heat that formed in the solid phase (-363.7 kJ/mol) and in the aqueous solutions (-318.5 kJ/mol) did not differ much, but the hydration of liberated ammonia in the aqueous solution may change the Gibbs energy of the reaction to be positive, which prevents the direct formation of the hydrated diamminesilver(I) perchlorate ($1\text{-H}_2\text{O}$) [20,21].

2.2. Phase Transition of Compound 1

Compound 1 has two polymorphs (monoclinic and orthorhombic), and the phase transition point was between 200 and 210 K [14]. The cryogenic DSC tests performed on the sample (Figure S1) showed that the phase transition temperature was 226 K, and the transition enthalpy was 103.04 kJ/mol.

2.3. General Spectroscopic Characterization and Correlation Analysis of Compounds **1-M** and **1-O**

The solid-state IR and the solid and aqueous solution-phase Raman spectra of the high-temperature orthorhombic polymorph of compound **1** (**1-O**) have already been discussed [14,16]. However, the spectral characteristics of the low-temperature monoclinic form (**1-M**) and the factor group analysis of the modifications have not been done yet. Therefore, correlation analyses were performed for perchlorate-ion, ammonia ligand, and silver ion (site groups were C_s and C_{2h} , and factor groups were C_{2h} and D_{2h} , for all three species in monoclinic (P_2/m) and orthorhombic ($Pnmm$) polymorphs, respectively). The correlation analysis showed the number of bands belonging to normal modes and distinguished them from combination or overtone bands of the spectra. The vibrational spectra of crystals may be quite difficult to interpret in the case of bands due to second-order (and higher-order) transitions. One of the problems is that although the total wave vector of the combined transition, \mathbf{k} , must be 0 ($\mathbf{k} = \mathbf{k}_1 + \mathbf{k}_2 = 0$), the individual values of the component vectors could be anywhere within the first Brillouin zone, which prevents the use of the terminology we are familiar with. In order to simplify things (and keeping in mind that this is not a paper that deals with solid-state physics primarily), we will continue to talk about “overtones” and “combinations” in the sense we use these terms for isolated molecules/ions. However, in reality, things are far more complicated.

2.4. Perchlorate Anion Modes

The isolated tetrahedral perchlorate ion has four normal modes, symmetric stretching ($\nu_1(A_1)$ (R)) and bending ($\nu_2(E)$ (R)), and antisymmetric stretching ($\nu_3(F_2)$ (IR, R)) and bending ($\nu_4(F_2)$ (IR, R)). The total number of factor group modes due to internal perchlorate vibrations is $2 \times 18 = 36$ ($2 \times 9(\text{IR}) + 2 \times 9$ (R)) and 36 ($18(\text{IR}) + 18(\text{R})$) vibrational degrees of freedom, for **1-M** and **1-O**, respectively (the monoclinic polymorph contains 2 crystallographically distinguishable perchlorate ions; thus, two sets of modes are available) (Figure 3). The hindered rotations (R_{xyz})(F_1) and translations (T_{xyz} (F_2)) of perchlorate ions result in $2 \times 12 = 24$ ($2 \times 6(\text{IR}) + 2 \times 6(\text{R})$) and 24 ($12(\text{IR}) + 12(\text{R})$) vibrational degrees of freedom, respectively (Figure S2a,b).

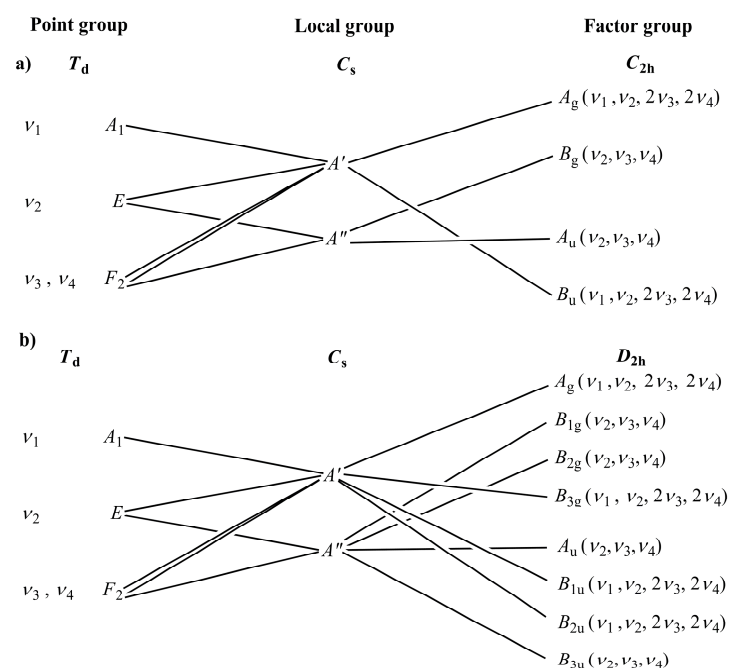


Figure 3. Correlation diagrams of perchlorate ions in (a) monoclinic (**1-M**) and (b) orthorhombic (**1-O**) polymorphs of $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$.

2.5. Ammonia Ligand Modes

The C_{3v} ammonia ligand has two symmetric ($\nu_1(A_1)$ (IR, R) stretching and $\nu_2(A_1)$ (IR, R) bending) and two antisymmetric ($\nu_3(E)$ (IR, R) stretching and $\nu_4(E)$ (IR, R) bending) internal normal modes. The total number of factor group modes due to internal vibrations is $4 \times 12 = 48$ ($4 \times 6(\text{IR}) + 4 \times 6(\text{R})$) (four crystallographically distinguishable ammonia molecules) and $2 \times 24 = 48$ ($2 \times 12(\text{IR}) + 2 \times 12(\text{R})$) (two crystallographically distinguishable ammonia molecules) (Figure 4).

The external vibrational modes ($T_z(A_1)$, $R_z(A_2)$ and $T_{xy}(E)$ or $R_{xy}(E)$) must be quadrupled due to four different ammonia molecules in the structure, which results in $4 \times 12 = 48$ ($4 \times 6(\text{IR}) + 4 \times 6(\text{R})$) and $2 \times 24 = 48$ ($2 \times 12(\text{IR}) + 2 \times 12(\text{R})$) vibrational degrees of freedom (Figure S3a,b).

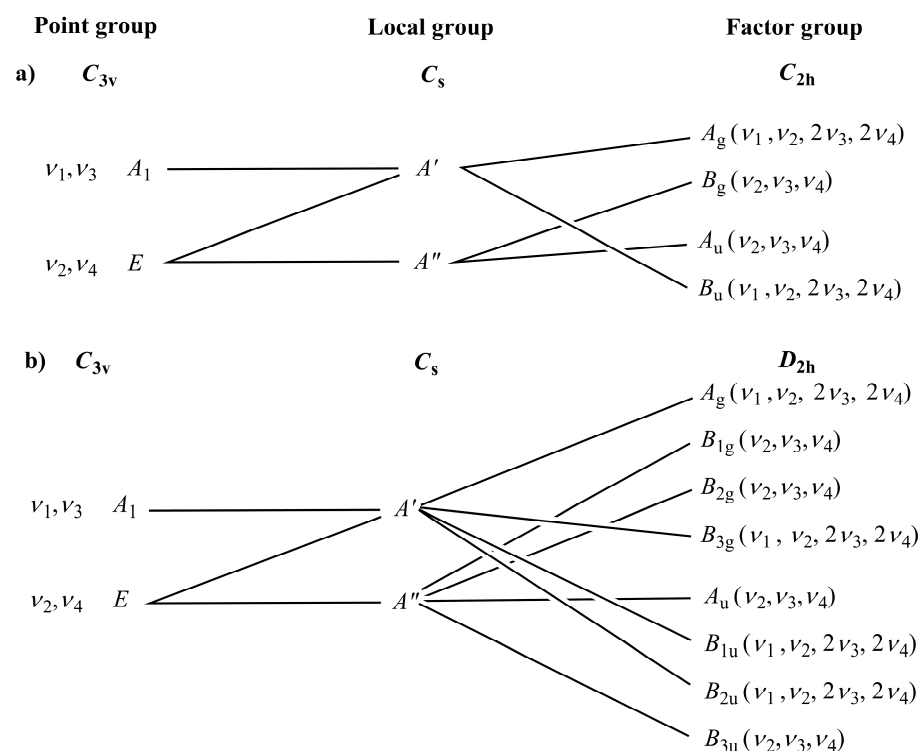


Figure 4. Correlation diagrams of ammonia ligand in (a) monoclinic (**1-M**) and (b) orthorhombic (**1-O**) polymorphs of $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$.

Silver cations, four and two crystallographic types in the **1-M** and **1-O** polymorphs, have 4×3 (T_z (A_u), T_x , T_y ($2B_u$)) and 2×6 (T_z (A_u), T_z (B_{1u}), T_x, T_y ($2B_{2u}$), and (T_x, T_y ($2B_{3u}$)) vibrational degrees of freedom, respectively. All of these are only infrared active modes. For both monoclinic and orthorhombic polymorphs, among 48 modes of translational origin, there are 3 acoustic modes that belong to $A_u + 2B_u$ modes, and the further 45 are optical modes (Figure 5). There are also 36 and 84 optical modes of rotational origin and due to internal vibrations, respectively.

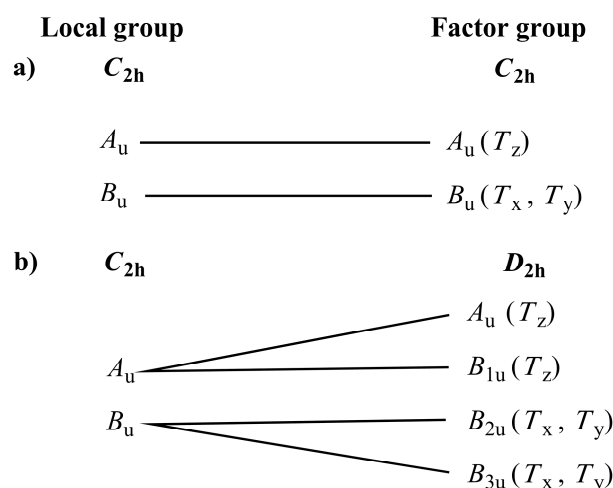


Figure 5. Correlation diagrams of silver in (a) monoclinic (**1-M**) and (b) orthorhombic (**1-O**) polymorphs of $[Ag(NH_3)_2]ClO_4$.

2.6. Assignment of Vibrational Modes in Compound 1 Polymorphs

2.6.1. Anion Modes

The assignment of perchlorate IR bands and Raman shifts found in IR and Raman spectra of compound **1** can be seen in Table 1. The IR data were measured only at room temperature (**1-O**), but Raman shifts were recorded at 123 K (**1-M**) and 298 K (**1-O**) as well. The IR inactive symmetric stretching mode of perchlorate ion (ν_1) appears as a single peak at 933 cm^{-1} in the spectrum of **1-O**, which unambiguously shows the distortion of symmetric T_d geometry. The antisymmetric stretching mode, as the most intense Cl–O band, is centered around 1050 cm^{-1} . This band has a strongly asymmetric shape, and the shoulder detected on the left side shows the splitting of triple (F_2) degeneracy as the consequence of the distortion of the geometry. The doubly and triply degenerated relatively weak $\nu_2(\text{Cl–O})(E)(450\text{ cm}^{-1})$ and $\nu_4(\text{Cl–O})(F_2)(615\text{ cm}^{-1})$, respectively, were not split in the room temperature spectrum. The band system between 2050 and 1950 cm^{-1} contains $\nu_1 + \nu_{3a,b,c}$ combinations and $2\nu_{3a,b,c}$ type (a,b,c —each of three bands set from F_2 ($A_u, 2B_{1u}, 2B_{2u}$, and B_{3u})) first overtones (Figure S2).

We recorded the Raman spectra at three different temperatures (123, 223, and 298 K) to study the effect of temperature and phase transformation on the numerical value/splitting of Raman shifts of **1-O** (298 and 223 K) and **1-M** (123 K). The spectra in the Cl–O stretching/bending modes range ($1200\text{--}300\text{ cm}^{-1}$) can be seen in Figure 6. At room temperature and at 223 K, the Raman spectra of **1-O** show a similar pattern. There is no pronounced splitting of bands, but a shoulder evolves into a band at 910 cm^{-1} , on the right at the bottom of the very intensive $\nu_1(A_1)$ Raman band (933 cm^{-1} and 929 cm^{-1} for **1-O** and **1-M**, respectively). This small intensity band (shoulder) might be the first overtone of $\nu_2(\text{Cl–O})(E)(457\text{ cm}^{-1})$ at 223 K. This overtone is a wideband system that may contain several components for **1-O** and a band consisting of two components for **1-M**. It is also plausible that isotopomers (^{35}Cl and ^{37}Cl) may also contribute to this band [22]. Overtones of $2\delta_{as}(\text{ClO})(F_2)(A_g, B_g)$ for **1-M** appears together with band(s) of $\delta_s(\text{NH})$.

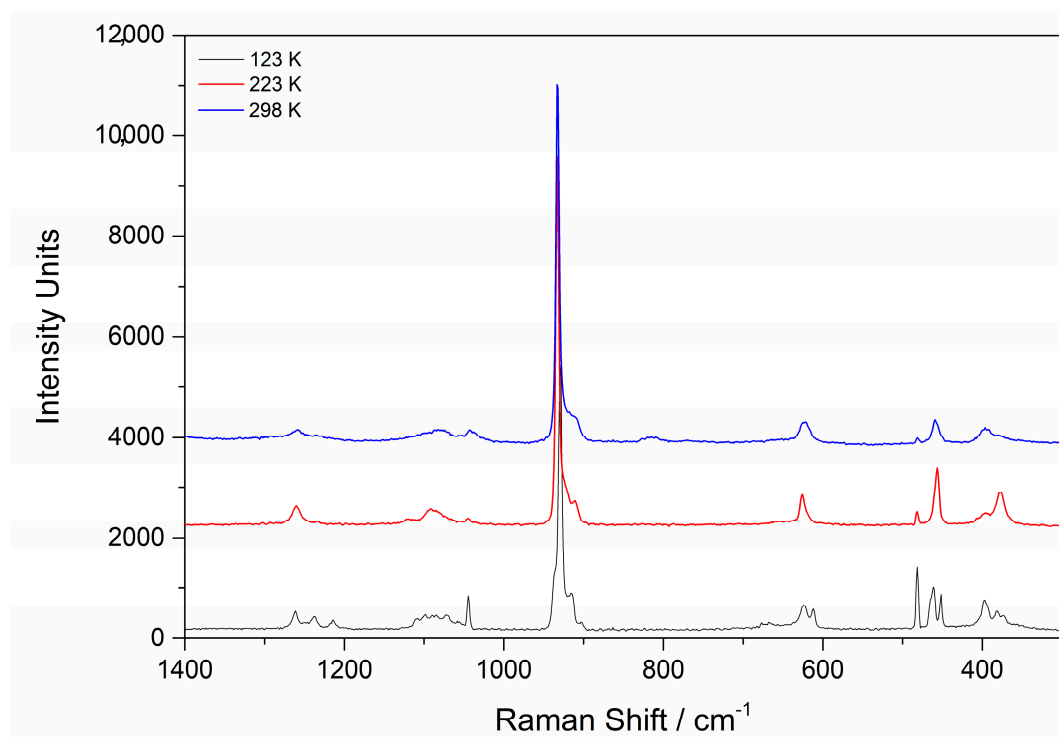


Figure 6. Raman spectra of compounds **1-O** (223 and 298 K) and **1-M** (123 K) in the Cl–O Raman shift range.

The combined effect of low temperature and phase transformation causes the splitting/appearance of new bands in the Raman spectrum of **1-M** at 223 K. The symmetric $\nu_2(\text{Cl-O})(\text{E})$ appears as a four-component band system, including a shoulder, which proves that two different types of perchlorate ions exist. The intensive and sharp $\nu_1(\text{Cl-O})(\text{A}_1)$ band did not split. Overlapping (* in Table 1 and 2) of the $\nu_3(\text{Cl-O})(\text{F}_2)$ bands with the combination band system of $\nu_{4\text{a,b,c}}(\text{Cl-O})(\text{F}_2)$ (2A_g and B_g for **1-M**, or 2A_g , B_{1g} , B_{2g} and 2B_{3g} for **1-O**) and $\nu_{2\text{a,b}}(\text{Cl-O})(\text{E})$ (A_g , B_g for **1-M** and A_g , B_{1g} , B_{2g} and B_{3g} for **1-O**) components resulted in a wide band, which appears at room temperature. There was a poorly separated band system (1107, 1098, 1089, 1084, 1072, 1058, 1054, and 1044 cm^{-1}) at 273 and 223 K for **1-O** and **1-M**, respectively. Combinations of $\nu_2 + \nu_4$ type seem to be located at the higher Raman shift side of the obtained range, while the $\nu_3(\text{Cl-O})(\text{F}_2)$ bands are probably located in the lower or medium Raman shift range. The $\nu_2(\text{Cl-O})(\text{E})$ band system contains two bands at 624 and 611 cm^{-1} , which might be attributed to the two types of perchlorate ion and the doublet nature of the bands as well.

Table 1. Assignment of perchlorate bands in IR and Raman spectra of compound **1**. (*- mixed bands)

IR, cm ⁻¹ (298 K)			Raman Shift, cm ⁻¹ (532 nm Excitation), Our Measurement			Raman Shift, cm ⁻¹		Assignment
Our (1-O)	[16]	[14]	123 K (1-M)	223 K (1-O)	298 K (1-O)	aq. soln. [16]		solid [14]
933	940	-	929	-	933	934	926	$\nu_1(A_1)$ (R)
450	432	-	482,465 sh, 462, 452	-	459	460	460	$\nu_2(E)$ (R)
1073 sh, 1053	1090	1084	1109–1044 *	-	1082, 1044	1107	-	$\nu_3(F_2)$ (IR, R)
615	626	626	624, 612	625	625	629	-	$\nu_4(F_2)$ (IR, R)
2050–1950 wide	-	-	-	-	-	-	-	$\nu_1 + \nu_3^{a,b,c}$ and $2\nu_3^{a,b,c}$
	-	-	1109–1044 *	-	1080 wide	-	-	$\nu_2^{a,b} + \nu_4^{a,b,c}$
	-	-	919, 914	915 sh, 910	915–910 wide	-	-	$2\nu_2^{a,b}$
	-	-	1259	1263, 1239	1263, 1239, 1219	-	-	$4\nu_4^{a,b,c}$

2.6.2. Cation Modes

The normal modes that belong to ammonia and the Ag–N linkages of compounds **1-O** and **1-M** are summarized in Table 2. The antisymmetric and symmetric N–H stretching (3364 and 3286 cm⁻¹, respectively) modes in the IR spectrum of **1-O** partly overlap the first overtone of the $\delta_{as}(NH)$ (~2×1653 cm⁻¹, wide). The antisymmetric deformation of ammonia (ν_4) is a doubly degenerated mode, and for four different ammonias one could expect 4 bands in the IR spectrum. In the Raman spectrum, this band is located at 3370 cm⁻¹ (A_g , B_g), which decomposes into three components at 123 K (**1-M**, A_g , B_{1g} , B_{2g} , B_{3g} , Table 2). The symmetric stretching band has a higher intensity in the Raman spectra than the antisymmetric one (3297 cm⁻¹ for **1-O** (A_g , B_{3g}), and 3299 and 3291sh cm⁻¹ (A_g , B_g) for **1-M**). The first overtone bands of $\delta_{as}(NH)$ gave two IR (3212 and 3185 cm⁻¹, (B_{1u} , B_{2u})) and two Raman bands (3191 and 3183 cm⁻¹ (A_g , B_g)) for **1-M**.

The symmetric $\delta_s(NH)$ Raman bands for **1-O** and **1-M** appear as a singlet band at 1259 cm⁻¹ and 1219 cm⁻¹, respectively. Since the $2\delta_s(NH)$ has only one overtone band at 2439 cm⁻¹, only one band may be associated with this mode, and this is probably the band located at 1219 cm⁻¹. The other two bands located close to this mode (1263 and 1239 cm⁻¹) belonged to $2\delta_{as}(ClO)(F_2)$ ($2A_g$, B_g) overtones. Due to the singlet nature of this mode, two IR bands (1244 and 1210 cm⁻¹) (B_{1u} and B_{2u}) found in the IR spectrum of compound **1-O** belong to the two different ammonia environments.

The wide IR band with the antisymmetric shape at 654 cm⁻¹ belongs to $\rho(NH)$ mode, and one of its higher wavenumber components may result in the wide but weak overtone centered at 1353 cm⁻¹. In the Raman spectrum, this mode was found at 658 cm⁻¹ for **1-O**, and a doublet appeared in the Raman spectrum of **1-M** (679, 666 cm⁻¹).

The silver-nitrogen vibrational modes are in the far-IR region. The far-IR spectra of compound **1-O** can be seen in Figure S4. There was a very wide band at 136 assigned as a mixed band of $\delta(OAgO)$, $\delta(OAgN)$, and $\delta(NAgN)$ modes, and a lower intensity and less wide band belonging to a lattice mode at 67 cm⁻¹.

Table 2. Assignment of bands belonging to cation modes in IR and Raman spectra of compound 1. (*- mixed bands)

IR		Raman		Assignment		
Our, 1-O, 298K	[14]	[16]	Our, 1-O, 298K	Our, 1-M, 123K	[14]	aq. solution [16]
3364		3360	3370	3389, 3362		-
3285	3285	3290	3297	3299, 3291		3293
3193, 3182	3182	3185	3185	3191, 3183		3203
				2439		
1610	1612	1625, 1610				1658, 1634
1351	-	-	-	-	-	-
1213, 1238	1250	1244	1259	1219	1219	1223
835			814	-	-	
653	-	650	658	679, 666	679, 666	
		430				
-			395*	397	397	
-			379*	383, 373	372	
-			269			
-	255		243		240	
			196		-	
136						Lattice, $\delta(\text{NAgO})$, AgO_2 , 161 AgN_2
67						Lattice?

2.7. UV Spectroscopic Studies of Complex 1

Both the perchlorate and diamminesilver ions are colorless, thus no absorption is expected in the visible spectral region. Due to the closed d^{10} shell of the silver⁺ ion, there are no d-d transitions, but there is a weak Ag^{4d-5s} transition at 250 nm, and the wide band at 310 nm may be an O–Ag or N–Ag LMCT band (Figure S5) [22–24]. The perchlorate ion in itself is UV silent [10,24]).

2.8. Non-Isothermal Decomposition of Compound 1 in an Inert Atmosphere and Air

Due to the strong oxidation power of the perchlorate anion and the reducing ability of ammonia ligands, there are two possible thermal decomposition routes of diammine-silver(I) perchlorate: (1) a solid-phase redox reaction between the perchlorate anion and the ammonia ligand, as was found in the case of permanganate [23], and (2) the loss of the ammonia ligand without any redox reaction, as was observed in the case of sulfate [25].

The decomposition process in an inert atmosphere starts with a slow mass decrease from about 50 °C. The process turns more intensive around 125 °C and results in a saw-tooth like DTG curve, with at least four decomposition processes, distinguished by their peak maxima at peak temperatures of 132, 173, 207, 251, and 408 °C. The character of the decomposition process is similar in both atmospheres (inert and air, Figure 7), but the mass decrease in argon is significantly higher than in air. The DTG peak maxima in air are also shifted below 300 °C by about 10 °C, and at higher temperatures by about 20 °C

toward higher temperatures, compared to peak maxima in argon. This result suggests that oxygen in the air hinders the thermal decomposition of **1**. The small mass loss of 0.7% in argon and 0.5% in air atmosphere up to ~ 125 °C most probably corresponds to adsorbed water and ammonia loss. Above this temperature, the decomposition processes overlap. In argon, in the next two steps (DTG maxima at 173 and 207 °C), complex **1** loses 7.6% and 7.8% of its mass, which is somewhat more than a molecule of NH_3 per step (7.06%). It suggests that not only ammonia loss occurs up to ~ 270 °C. An additional mass loss of 6.7% with a DTG peak at 251 °C is also in accordance with this. In air, the decomposition processes overlap more with DTG peaks at 181 and 196 °C and corresponding mass losses of 11.8% and 4.5%. The overall mass decrease up to ~ 270 °C in air (16.8%) is significantly less than in argon (22.8%). In argon, based on the TG results, a part of the ammonia evaporates, and the other part reacts with the oxygens of the perchlorate even in the solid phase. As the result of the decomposition reaction, the intermediate ($\text{AgClO}_{2.5}$) is formally the mixture of AgClO_3 and AgClO_2 , in a ratio of approximately 1:1. The decomposition temperatures of supposed intermediates (AgClO_2 , AgClO_3 , AgClO_4) do not agree with the decomposition temperatures we found for the decomposition intermediates of compound **1**. First, AgClO_4 does not decompose at 220 °C into lower chlorine oxyacid salts because the thermal stability of AgClO_x ($x = 2-4$) salts increases with an increasing oxidation degree of chlorine in the oxyacid residue [26–28]. According to this result, the thermal decomposition temperatures of these chlorine oxyacid salts, AgClO_2 , AgClO_3 , and AgClO_4 are 156 °C, 270 °C, and 476 °C, respectively [26]. It means that the formal $\text{AgClO}_{2.5}$ composition, which decomposes at 441 °C, cannot be a mixture of AgClO_2 and AgClO_3 , but only a mixture of AgCl and AgClO_4 . On the other hand, in air, the mass of the intermediate at ~ 285 °C is 83.3%, which is in good agreement with the calculated value for AgClO_4 (85.89%). A possible explanation is that the aerial oxygen takes a role in ammonia oxidation instead of the ClO_4^- ion. The mass of the final residue in argon is very low, only 47.2%, somewhat higher than the metallic silver content of **1**, while in air, it is 60.3% at 500 °C. This is in good agreement with the theoretical value for AgCl (59.38%).

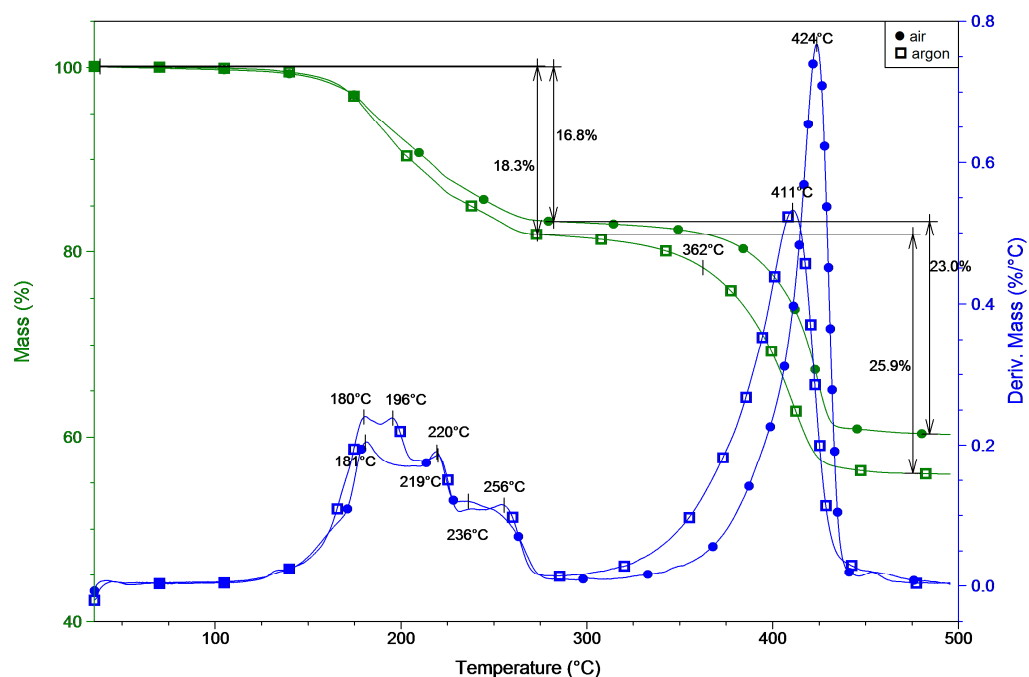


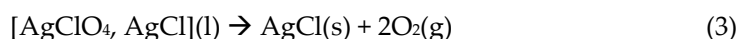
Figure 7. DTG curve of compound **1** in an inert (□) and air (●) atmosphere.

In an inert atmosphere, the first two decomposition processes are followed by endothermic heat effects (DSC minima at 174 and 211 °C), and the third process by an exothermic effect with a DSC maximum at 256 °C. In air, the tendency of the DSC curve is the same, but with higher peak temperatures and lower enthalpy changes (Figure S6). These heat effects support the theory of partial ammonia oxidation discussed above.

The last decomposition step with a DTG maximum at 408 °C is an exothermic process, with a mass loss of 30.1%.

2.9. The Nature of the Final Decomposition Product Formed in the Thermal Decomposition of Compound 1

To solve the problems listed above, we studied the final decomposition product (P-480, prepared by heating the sample at 480 °C for 2 h) by XRD. The powder X-ray diffractogram of the P-480 sample showed the absence of metallic silver and the presence of silver chloride (Figure S7). Thus, the last decomposition step that produces AgCl is a thermal decomposition reaction of silver(I) perchlorate into silver(I) chloride and oxygen.



This reaction can explain the unexpected mass loss in the last decomposition step, because silver chloride melts at 452 °C (Figure S6) and its vapor pressure above the liquid and the inert gas stream results in a partial evaporation of AgCl [29]. The decomposition peak temperature ($T_{\text{dec}} = 441$ °C) of AgClO₄ is lower than expected ($T_{\text{dec}} = 476$ °C) [26], because AgCl (Mp. = 452 °C) and AgClO₄ (Mp. = 476 °C) form a eutectic melt (Mp. 325 °C, 1:1 AgCl:AgClO₄), and the decomposition temperature of AgClO₄ in this melt is lower than the decomposition point of pure AgClO₄. The silver perchlorate decomposes in the liquid melt phase in an autocatalytic process at a lower temperature than in the solid phase [26]. The reaction heat was measured to be -85.68 kJ/mol.

2.10. Redox Interactions in the Thermal Decomposition Steps of Compound 1

The presence of AgCl in the decomposition intermediate formed at 300 °C—deeply below the thermal decomposition temperature of AgClO₄—confirms that AgCl was formed in the solid phase quasi-intramolecular redox reaction of [Ag(NH₃)₂]ClO₄ (compound 1). The TG-MS results (Figure 8, Figure S8) show that in an inert atmosphere, redox reactions take place between 180 and 300 °C, when N₂ ($m/z = 28$ (N₂⁺) and H₂O (18 (H₂O⁺)) as ammonia oxidation products are formed. Supposing that ammonia is only oxidized into N₂ and H₂O, the six hydrogens of two molecules of ammonia in compound 1 can convert the perchlorate ion into formal “AgClO” (AgClO, AgCl + AgClO₂ = 1:1, 2AgCl + AgClO₃ or 3AgCl + AgClO₄). This shows that 3/4 of the overall perchlorate ion content can only be reduced into AgCl. (Figure 9).

In an inert atmosphere, a small amount of nitric oxide ($m/z = 30$, NO), oxygen ($m/z = 16$, O⁺), and HCl ($m/z = 36$, HCl⁺) were also detected. Nitrous oxide ($m/z = 44$, N₂O⁺) was not produced. The peak at $m/z = 16$ cannot be an NH₂⁺ fragment of ammonia, because then, a much more intensive $m/z = 17$ (NH₃⁺) peak would be observed; instead of that, only a weak $m/z = 17$ (OH⁺) peak was found (Figure S8) [3].

In air, NO is the main oxidation product. HCl did not form, and a small amount of N₂O ($m/z = 44$, N₂O⁺) was also detected (Figure S9). The lack of HCl and the appearance of N₂O might be attributed to the concurrent oxidation reaction with aerial oxygen instead of the perchlorate ion. The solid-state reaction product at 300 °C in air contains AgCl, AgClO₄ (cubic and tetragonal), and AgClO₄·H₂O, together with a minor amount of metallic Ag (Figure 10). Since AgClO₄ is extremely hygroscopic, AgClO₄·H₂O forms during the preparation of the samples for XRD by taking up humidity from the air.

With the use of stoichiometric relations, the formal reaction equations for the main and by-reactions (redox and non-redox) can be summarized according to Figure 9.

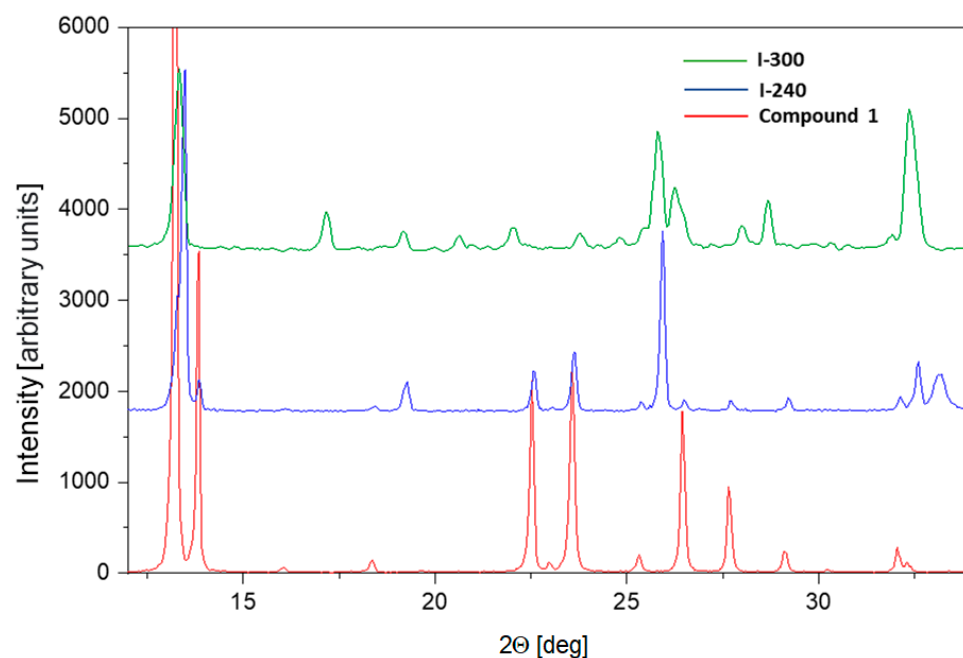


Figure 10. The powder X-ray diffractogram of compound 1 and decomposition intermediate 240 at 300 °C.

The composition of **I-240** changed during storage in air for 3 days. The XRD of **I-240** recorded as a function of time (Figure 11) showed that the peak intensities of silver perchlorate monohydrate increased, whereas the peak intensities belonging to the anhydrous polymorphs almost completely disappeared with increasing storage time. It means the changes in phase composition can be attributed to the hygroscopic nature of AgClO_4 ; the transformation of the primary decomposition product is fast and decomposes easily into AgClO_4 and AgCl even at 240 °C. However, there are some unassigned XRD peaks of the **I-240** intermediate, which belong to a compound that contains coordinated ammonia (see the IR spectrum in Figure S10).

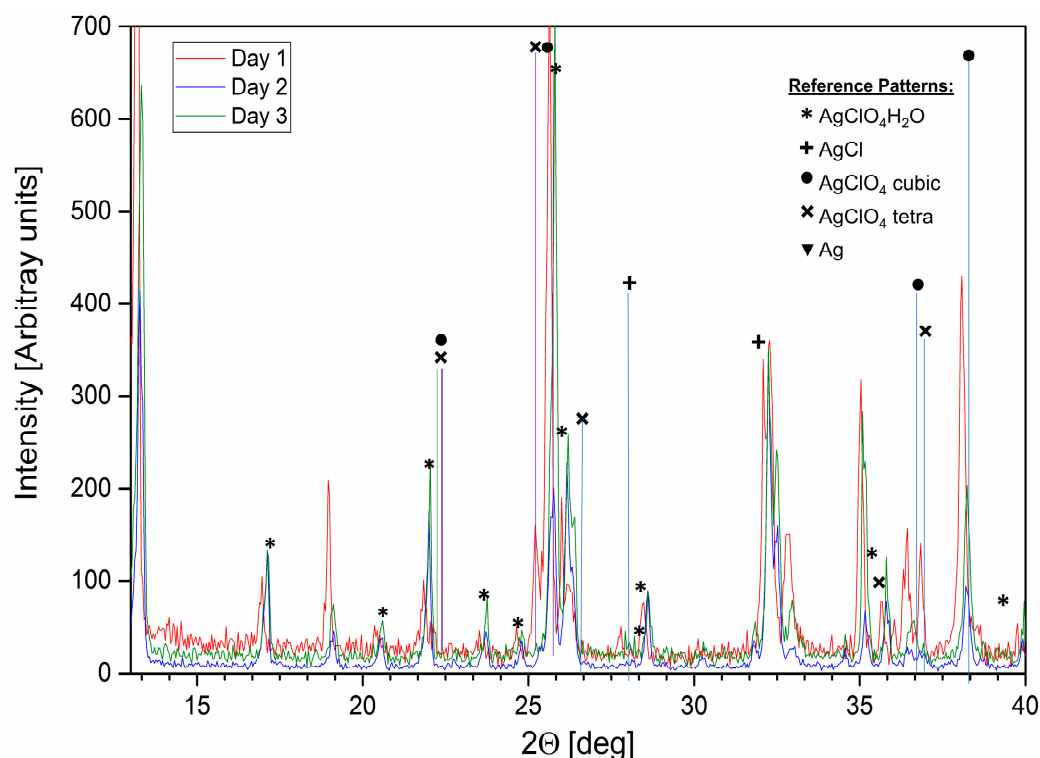


Figure 11. The powder X-ray diffractograms of the change of the composition of **I-240**.

There are important differences in the shape of the perchlorate $\nu_{\text{as}}(\text{ClO})$ bands, and band positions belong to the coordination-sensitive ammonia modes ($\delta(\text{NH})$ and $\rho(\text{NH})$) in the IR spectra of **I-240** and compound **1** (Figure 12). Aqueous extraction of **I-240** and the evaporation of the solution resulted in silver perchlorate monohydrate and an unknown phase. The IR (Figure 12) spectrum showed an important shift (60 cm^{-1}) of peak positions and the shape of $\nu_{\text{as}}(\text{Cl-O})$. The $\delta_{\text{s}}(\text{NH})$ and $\rho(\text{NH})$ bands can be assigned to a more strongly coordinated NH_3 ligand, as in compound **1** (shifting the $\delta(\text{NH})$ towards the higher wavenumber range [30]). The unknown phase may be a silver salt of a chlorine oxyacid, which contains fewer than 2 ammonia per silver (e.g., a polymeric mono-ammine). This intermediate can decompose into the same compounds (N_2 , H_2O , AgCl , AgClO_4) as compound **1**. We attributed the change in the shape of the antisymmetric Cl-O IR band to the presence of two different chlorine oxyacid compounds in the intermediate **I-240**. One is a perchlorate, the other is a chlorate compound. For the chlorate component, the $\nu_{\text{as}}(\text{Cl-O})$ bands (AgClO_3 ($\nu_{\text{as}}(\text{ClO}) = 987, 967$) (Figure 12)) are expected to appear in the lower wavenumber range of the wide perchlorate $\nu_{\text{as}}(\text{ClO})$ band. The presence of diamminesilver(I) chlorate cannot be expected because it melts and decomposes at $100\text{ }^\circ\text{C}$ with the formation of AgClO_3 [31].

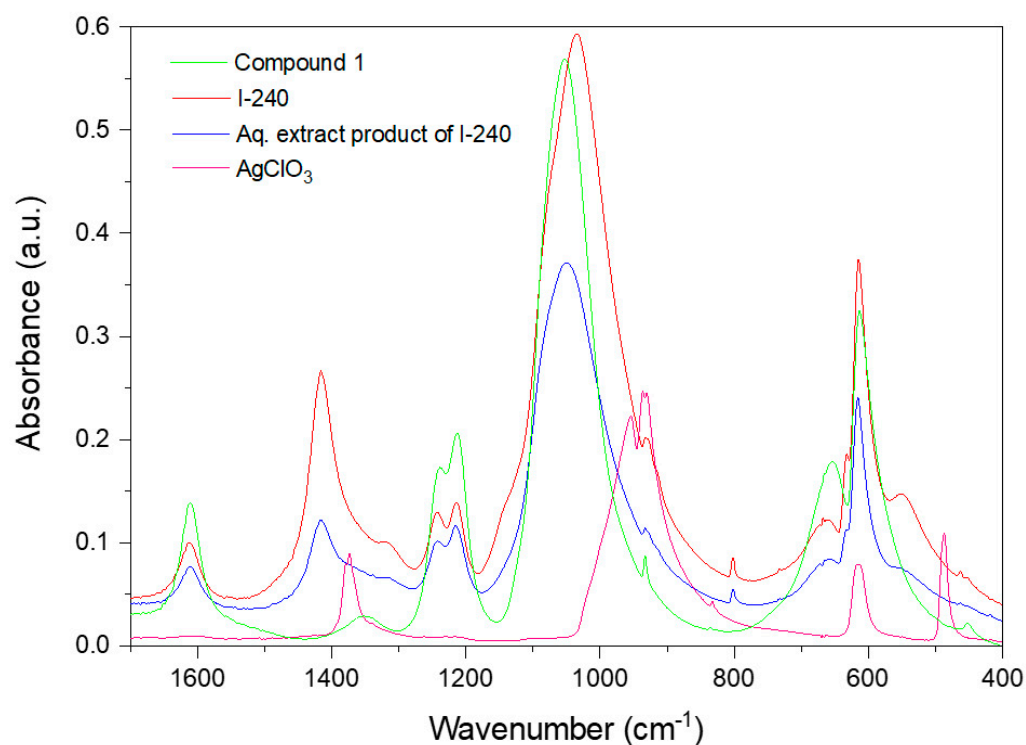


Figure 12. The IR spectroscopic study of the decomposition of the **I-240**.

The relative amount of perchlorate, chlorate, and chloride ions in the samples **I-240** and aq. extract of **I-240** were found to be 8.25:3.35:1 and 8.02:3.01:1, respectively. The chlorite intermediate contents of the two samples are close to each other, and ca. 1/4 of the compounds containing chlorine proved to be chlorate ions on both sample surfaces. The silver XPS signals (368.3 and 368.5 eV) are the same for all three expected component types (AgCl, AgClO₃, and AgClO₄ or their complexes with ammonia). The oxygen signals of chlorate and perchlorate ions could easily be distinguished (531.2–531.9 eV: chlorate, and 532.8 eV: perchlorate) (Figure 13).

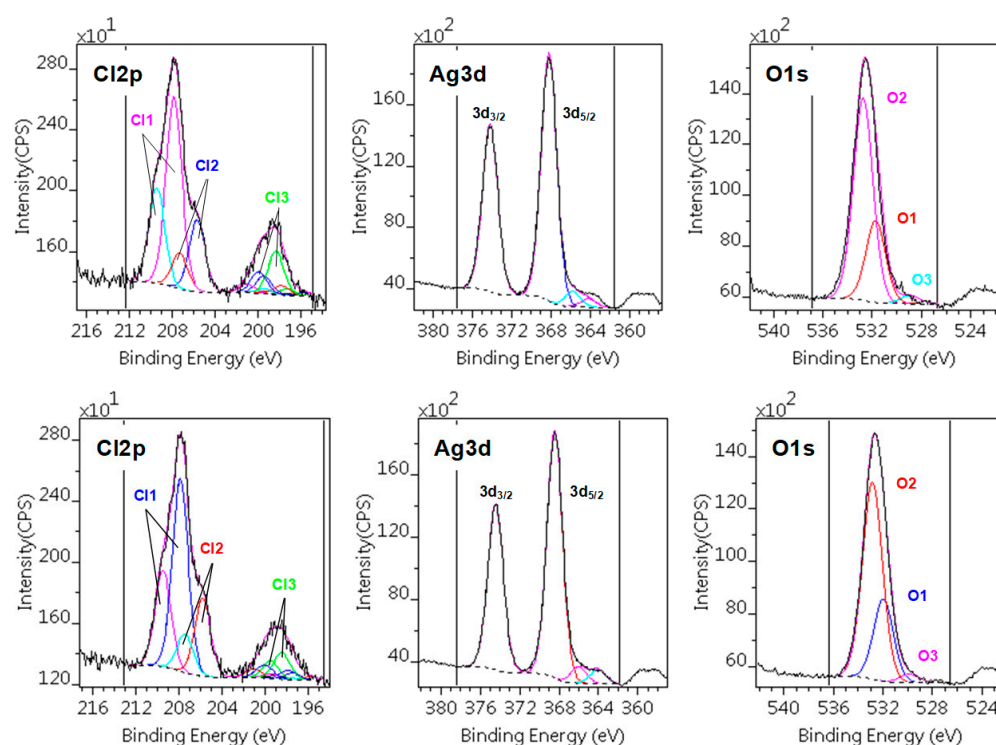


Figure 13. The XPS spectrum of the (a) **I-240** and (b) the solid isolated from the aq. extract of **I-240**.

Summarizing the IR, XRD, and XPS results, the reduction process is not a single-step process, but probably different parallel and consecutive reactions take place with intermediates like AgClO , AgClO_2 , and AgClO_3 , or their NH_3 complexes. Taking into consideration the gaseous and solid reaction products, the relative amounts of these, the IR and XPS results of the intermediates, and the decomposition temperatures of AgClO , AgClO_2 , AgClO_3 , AgClO_4 , and their known ammonia complexes, we suggest the following general decomposition scheme (Figure 14). The stable intermediate complex formed at the beginning of decomposition (180–240 °C) may only be a material containing chlorate.

The decomposition reaction temperature of compound **1** is >180 °C; thus, the reactions on the left side of the scheme (decomposition routes/disproportionation of AgClO and AgClO_2) may only be formed/decomposed in situ processes. TG-MS shows the presence of O_2 ($m/z = 32$, O_2^+) in the temperature range of the decomposition of compound **1**, thus the evolution of oxygen below 235 °C (the melting point of AgClO_3) probably belongs to the decomposition of $\text{AgClO}/\text{AgClO}_2$, while above this temperature, it can only be a result of the decomposition of AgClO_3 into AgCl and O_2 . The AgClO_3 detected by IR and XPS may form directly (Figure 14) by the reduction of perchlorate ion, or from AgClO_2 via disproportionation as well. The consecutive character of the first decomposition steps strongly suggests that the consecutive disproportionation steps (AgClO - AgClO_2 - AgClO_3) are the main decomposition steps, and the decomposition with oxygen evolution, the redox reaction with HCl and Ag formation, and the non-redox thermal deammoniation of compound **1** are only by-reactions.

DSC tests were performed with a Perkin Elmer DSC 7 instrument (Perkin-Elmer, Waltham, USA). Measurement details were the same as in our previous publications [25,34].

4. Conclusions

The reaction of ammoniacal AgNO_3 solution or $[\text{Ag}(\text{NH}_3)_2]\text{NO}_3$ with aq. NaClO_4 resulted in $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$. We did not find a monohydrated form described previously, and the hydration of the diamminesilver(I) cation cannot be expected based on thermochemical considerations.

The temperature and enthalpy of phase change were determined by DSC, and found to be 225.7 K and 103.04 kJ/mol, respectively. Detailed spectroscopic (correlation analysis, IR, Raman, and UV) analyses were performed on $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$, and all the vibrational modes in the IR and Raman spectra were assigned.

The thermal decomposition of $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$ involves a solid-phase quasi-intramolecular redox reaction between the perchlorate anion and the ammonia ligand, resulting in lower valence chlorine oxyacid (chlorite, chlorate) components.

The formal loss of ammonia and the formation of AgClO_4 during heating is not a favored reaction route. The detected chlorate/chlorite intermediates disproportionate both in solid phase or in aqueous solutions (after dissolution of these decomposition intermediates in water) into AgCl and silver perchlorate.

The solid-phase AgCl - AgClO_4 mixture eutectically melts, and the resulting AgClO_4 decomposes in this melt into AgCl and O_2 ; thus, the final product of decomposition is AgCl , N_2 , and H_2O .

The intermediate (chlorite, chlorate) phases were detected by IR, XPS, and titrimetric methods.

Supplementary Materials: The following are available online at www.mdpi.com/2304-6740/9/5/38/s1, Figure S1: Cryogenic DSC results of $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$; Figure S2: (a) and (b) The hindered rotations (R_{xyz})(F1) and translations (T_{xyz} (F2)) of perchlorate ion in the polymorphs of $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$; Figure S3(a) and (b): External vibrational modes (T_z (A1), R_z (A2) and T_{xy} (E) or R_{xy} (E)) for ammonia in the polymorphs of $[\text{Ag}(\text{NH}_3)_2]\text{ClO}_4$; Figure S4: The far-IR spectra of compound **1-O**; Figure S5: The UV spectra of compound **1**; Figure S6: TG-DTG-DSC results of compound **1** in air; Figure S7: The powder X-ray diffractogram of **P-480**; Figure S8: TG-MS of compound **1** in inert atmosphere, formation of $m/z = 14$, $m/z = 16$, $m/z = 17$, $m/z = 18$; Figure S9: TG-MS of compound **1** in air atmosphere, formation of $m/z = 44$; Figure S10: IR spectrum of **I-240** intermediate.

Author Contributions: Methodology, formal analysis, V.M.P.; investigation, L.A.F., L.B., M.M., B.B.-H., K.A.B., I.E.S.; resources, I.M.S.; data curation, V.M.P., L.A.F., K.A.B., L.K.; writing—original draft preparation, L.A.F., L.K.; writing—review and editing, K.A.B., F.P.F., V.M.P., B.B.-H., L.K.; visualization, L.A.F.; supervision, V.M.P., B.B.-H., L.K.; project administration, L.A.F.; funding acquisition, I.M.S. All authors have read and agreed to the published version of the manuscript.

Funding: The research within project No. VEKOP-2.3.2-16-2017-00013 was supported by the European Union and the State of Hungary, co-financed by the European Regional Development Fund. I.M.Sz. thanks to the Hungarian Academy of Sciences for a János Bolyai Research Fellowship and acknowledges the ÚNKP-18-4-BME-238 grant supported by the New National Excellence Program of the Ministry of Human Capacities, Hungary. The research reported in this paper was supported by the Higher Education Excellence Program of the Ministry of Human Capacities in the frame of the Nanotechnology and Materials Science research area of Budapest University of Technology (BME FIKP-NAT). B.B.H. acknowledges financial support (Grant No. 451-03-68/2021-14/200125) of the Ministry of Education, Science, and Technological Development of the Republic of Serbia.

Institutional Review Board Statement: Not applicable.

Informed Consent Statement: Not applicable.

Data Availability Statement: Data sharing is not applicable to this article. The data presented in this study are available on request from the corresponding author.

Conflicts of Interest: The authors declare no conflict of interest.

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